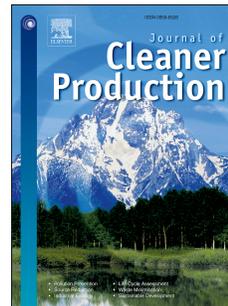


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Bio-lubricants production from fish oil residue by transesterification with trimethylolpropane

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Abstract

The fatty acid ethyl esters mixture, a fish oil residue obtained after the extraction of omega-3 polyunsaturated fatty esters, has been converted into mixtures of mono-, di-, and triesters of trimethylolpropane by transesterification at 100-140°C under vacuum with sodium ethoxide as catalyst. This method has shown to be more efficient than the enzymatic transesterification using commercially available lipases. The crude reaction mixture (84% conversion of ethyl esters), enriched in trimethylolpropane triesters (96% selectivity) was characterized and its properties compared with those of the starting ethyl esters mixture and the trimethylolpropane esters obtained from vegetal sources.

Keywords: bio-lubricants; waste fish oil; transesterification; trimethylolpropane

Abbreviations

ASTM	American Society for Testing and Materials
DE	Diesters
DHA	Docosahexaenoic acid
EPA	Eicosapentaenoic acid
EtONa	Sodium ethoxide
FAs	Fatty Acids
FAEEs	Fatty Acid Ethyl Esters
FAMEs	Fatty Acid Methyl Esters
ISO	International Standards Organization

ISO-VG	International Standards Organization Viscosity Grade
ME	Monoesters
MeONa	Sodium methoxide
MUFAs	Monounsaturated Fatty Acids
NMR	Nuclear Magnetic Resonance spectroscopy
PUFAs	Polyunsaturated Fatty Acids
SFAs	Saturated Fatty Acids
TE	Triesters
TMP	Trimethylolpropane
UNE	Una Norma Española (A Spanish Standard)

1. Introduction

Omega-3 fatty acids, abundant in fish oil, are engaged in various physiological actions as essential fatty acids and they play an important role in homeostatic maintenance in the body (Tokuyama and Nakamoto, 2014). Given this important role of omega-3 polyunsaturated fatty acids (PUFAs) in health (Ruxton et al., 2005), many nutrition products are enriched with omega-3 concentrates from natural sources to compensate their low consumption in the typical diets of many developed countries. In spite of the search for alternative sources of omega-3 fatty acids (Ward and Singh, 2005), fish remains as the main feedstock to produce omega-3 concentrates (Rubio-Rodríguez et al., 2010). Their manufacture implies the separation of the omega-3 PUFAs from the rest of fatty acids (FAs) present in the fish oil, which are then considered as a waste. Fish industry uses a billion tons of fish for edible purpose, which generates a very large amount of waste. Around 5.2 million tons/year in the European Union (Lopes et al., 2015) are produced of this non-edible waste. A considerable effort is being done to produce biofuels from those fish wastes, mainly by extraction of the lipids and transformation into biodiesel (Dave et al., 2014; Yuvaraj et al., 2016). Although specific installations have been designed, and the process through chemical transesterification with basic catalysts has been optimized, there is a low interest in the manufacturing of this bioproduct when the fish oil waste keeps high quality, in terms of characteristics as acid value, oxidation parameters, and presence of contaminants (such as dioxins, heavy metals, polychlorobiphenyls and polyaromatic hydrocarbons). In those cases, when the value of the waste is higher, searching of new valorization alternatives would be

desirable to obtain specific biobased products. This solution has been already proposed as an alternative for biodiesel manufacturers to extend the potential sales market (Kleinat  et al., 2014).

This strategy has been very recently explored from non-edible Tilapia (*Oreochromis niloticus*) oil obtained from viscera (Valle et al., 2018). In this case, methyl esters were epoxidized and the epoxides opened with medium size alcohols, a method also used in the valorization of other waste oils, such as cooking oil (Borugadda and Goud, 2016). Trimethylolpropane (TMP) fatty esters are considered as biolubricants (Panchal et al., 2017; Syahir et al., 2017; Zainal et al., 2018) with improved stability properties, in comparison with the analogous triglycerides (Sripada et al., 2013). TMP triesters have been prepared by esterification of TMP with free fatty acids or by transesterification of fatty acid methyl esters (FAMES). In the first case, the esterification has been carried out using either acid catalysis ( kerman et al., 2011a; Kamalakar et al., 2013) or immobilized lipases ( kerman et al., 2011b). On the other hand, transesterification has been described using either basic catalysis, usually sodium (Yunus et al., 2005) or calcium (Gryglewicz et al., 2003) methoxide, or immobilized lipases (Cavalcanti da Silva et al., 2015; Uosukainen et al., 1998). However, this strategy had not been applied to waste or fish oils.

In this paper, we compare the chemical and enzymatic methods of transesterification of TMP with methyl oleate, as a model, and the subsequent application of the best method to the fatty acid ethyl esters waste mixture from fish oil after PUFA extraction. Some of the key properties to be usable as biolubricant have been also measured.

2. Materials and methods

Nuclear Magnetic Resonance (NMR) spectroscopy.

Proton (^1H) spectra were recorded in a Bruker Avance 400 at a frequency of 400 MHz and the chemical shifts are expressed in ppm. Deuterated chloroform (CDCl_3) was used as solvent and reference (7.26 ppm). All the reactions were monitored (conversion and selectivity) by ^1H -NMR.

Enzymatic transesterification reactions.

A mixture of methyl oleate (2 g, 6.75 mmol), Novozym 435 (89 mg, 4% w/w), TMP (232 mg, 1.73 mmol) and in some cases a solvent (acetone or *tert*-butanol, 2 mL) was

stirred at 150 r.p.m. and 45°C (unless otherwise stated in table 1) for 4-24 h. Some of the reactions were also carried out under vacuum (1 mbar). After the corresponding reaction time, the immobilized enzyme was filtered off and washed with acetone (15 mL). Finally, the excess of alcohol and the organic solvent were eliminated under reduced pressure.

Base-catalyzed transesterification reactions.

A mixture of methyl oleate (2 g, 6.75 mmol), TMP (226.4 mg, 1.69 mmol) and EtONa (20 mg, 1% w/w, 1.47 mmol) was stirred and heated at 140°C under vacuum (1 mbar) for 6 h. The reaction was quenched by adding 10 mL of water and then HCl (35% w/w) until neutral pH. The mixture was extracted with CH₂Cl₂ (4 × 10 mL). The combined organic phases were dried with anhydrous MgSO₄ and filtered, and the solvent was evaporated under reduced pressure.

Optimized method to obtain triesters from the fish oil fatty esters mixture: 8.9 g of ethyl esters mixture (around 30 mmol), 1.00 g of TMP (7.15 mmol) and 89 mg of EtONa (1.46 mmol) were heated at 100°C under vacuum (1 mbar) for 3 h. The treatment was carried out in the same way as described above with 15 mL of water and 4 × 15 mL CH₂Cl₂ extractions.

Characterization of the properties of the TMP fatty esters.

The density at 20°C was determined according to the UNE 84156:2000 method. The kinematic viscosities were determined at 40°C and 100°C based on the ASTM D445 method. The viscosity index was calculated from kinematic viscosities according to ASTM D-2270-10 method. The pour point was determined according to ASTM D5950-14. The flash point was determined according to ASTM D92-16b.

3. Results and discussion

Determination of the results by NMR

Although the results can be determined by gas chromatography, as reported in the literature (Yunus et al., 2005; Sripada et al., 2013), it requires an additional functionalization (silanization) step, and the identification and integration of all the peaks of the real sample, which is a complex mixture of fatty esters (see below). Thus, we envisaged the use of ¹H-NMR, a method that has shown its usefulness in the

analysis of vegetable oils (Popescu et al., 2015). The key point was to distinguish the protons of TMP after one, two or three esterifications, which would allow directly quantifying the molar ratio of mono- (ME), di- (DE) and triesters (TE) of TMP. The identification of the signals was carried out using reaction mixtures with different conversions and ratios of products, as can be seen in Figure 1. The CH₂OH groups shift from 3.73 ppm in TMP (6H) to 3.58-3.51 ppm in ME (two doublets corresponding to diastereotopic protons, 4H) and to 3.39 ppm in DE (2H). The esterified groups (CH₂OOC-) appear at higher chemical shifts, 4.17 ppm in ME (2H), 4.01 ppm in DE (4H) and 4.00 ppm in TE (6H). These two signals are slightly overlapped, but they can be easily deconvoluted. All those ester signals are different from those of the starting fatty esters, 4.10 ppm (q, 2H) of ethyl esters and 3.66 ppm (s, 3H) of methyl ester (not shown in Figure 1). The -CH₂-COO signal at 2.37-2.25 ppm was used as reference (2H) to calculate the conversion and to check the absence of a significant amount of hydrolysis to fatty acids.

Enzymatic transesterification of methyl oleate with TMP

Novozym 435 (*Candida antarctica*) lipase was chosen as a typical successful enzymatic catalyst for transesterification reactions (Linko et al., 1998; Åkerman et al., 2011b; Gryglewicz et al., 2013; Cavalcanti da Silva et al., 2015). In a first set of experiments, the reaction was carried out with the stoichiometric amount of reagents or a slight excess of methyl oleate (oleate/TMP molar ratio = 3.0-3.9) under atmospheric pressure, with 4% w/w of enzyme and 15% w/w of water. Moderate conversion of methyl oleate was obtained (24- 49%), even after 24 h, but only hydrolysis to oleic acid was observed. Only in the case of carrying out the reaction under vacuum (1 mbar), to shift the equilibrium by the elimination of methanol from the reaction medium, a low conversion to TMP esters (16%) was obtained. Thus, the rest of experiments were carried out in the absence of water, either with a solvent to assist the contact of the enzyme with the reagents or in the absence of solvent. Some reactions were carried out under vacuum. The reaction conditions and the results are collected in Table 1.

At atmospheric pressure (entry 1), the conversion is very low, 10% to TMP esters and 7% to oleic acid (hydrolysis). One possible explanation is the poor contact between the hydrophilic environment of the enzyme and the hydrophobic reagents. Acetone was used in an attempt to improve the contact, but no significant improvements were obtained (entries 2 and 3). The application of dynamic vacuum does not improve either

the results, unless the temperature is raised to 100°C, although the result is highly dependent on the amount of enzyme. The best result (47% conversion, with only 14% of hydrolysis, entry 6) was obtained with only 40 mg of enzyme, at 100°C and 1 mbar, leading to a mixture of TMP diester (DE) and triester (TE) in 82/18 ratio. The use of acetone as a co-solvent at lower temperature (60°C) reduces the hydrolysis to only 7%, together with a change in selectivity, as a mixture of ME and DE (55/44) is obtained (entry 8). The dynamic vacuum eliminates in fact the acetone from the reaction mixture, and then a reaction was carried out under atmospheric pressure for 24 h, to ensure the mixing of enzyme and reagents, and then 12 h under vacuum (entry 9), with a oleate/TMP molar ratio of 1, to favor the formation of ME. However, this strategy led to lower oleate conversion. The best result was obtained at atmospheric pressure, by using acetone for 6 h, and then evaporation of the solvent before continuing the reaction for an additional 18 h period (entry 10). In this way, in spite of the larger amount of enzyme, the conversion was only slightly better, but the hydrolysis was nearly suppressed.

Base-catalyzed transesterification of methyl oleate with TMP

Given the disappointing results obtained with Novozym 435, the base catalyzed transesterification was tried, using sodium alkoxides (Table 2). The reactions were carried out without solvent at 140°C under vacuum, given that the conversion at atmospheric pressure was very low (entry 1). The first advantage of the base catalyzed transesterification is the lack of hydrolysis, as oleic acid was not detected in any of the reactions. With only 1% sodium ethoxide (EtONa), it is possible to get a nearly complete conversion of methyl oleate by using an excess of TMP (entries 2 and 3), leading to a mixture of ME (as the main component) and DE. The TE can be obtained under similar conditions, by only changing the oleate/TMP ratio to a slight excess of oleate, with also nearly total conversion (entries 4 and 5), given that the theoretical maximum is 75%.

The use of EtONa leads also to a small amount of ethyl oleate, and hence sodium methoxide (MeONa) was also tested to prevent this effect. The results were not so favorable, with lower conversion under the same conditions (entry 7 vs entry 4). The equimolecular mixture of oleate and TMP gives a nearly statistical distribution of ME, DE and TE (entry 6).

Table 1. Results of the transesterification of methyl oleate with TMP using Novozym 435 lipase.^a

Entry	Oleate/TMP molar ratio	Novozym mass (mg)	Solvent	P	T (°C)	time (h)	Conv. (%) ^b	Hydrol. (%) ^b	Selectivity (%)		
									ME	DE	TE
1	3.9:1	89	-	atm	45	24	17	7	58	42	0
2	3.9:1	40	Acetone ^c	atm	45	24	10	6	77	23	0
3	3:1	40	Acetone	atm	60	48	12	8	100	0	0
4	3.9:1	89	-	1 mbar	45	6	13	9	63	32	5
5	3.9:1	89	t-BuOH	1 mbar	45	6	8	3	86	11	3
6	3:1	40	-	1 mbar	100	6	47	14	0	82	18
7	3:1	150	-	1 mbar	100	6	16	10	67	33	0
8	3:1	40	Acetone	1 mbar	60	36	47	7	55	44	1
9	1:1	40	Acetone	atm-1 mbar ^d	60	36 ^d	27	8	66	34	0
10	1:1	100	Acetone ^e	atm	60	24 ^e	53	1	61	39	0

^a Reaction conditions: methyl oleate (2g, 6.75 mmol), Novozym 435 (mass indicated in table), TMP (molar ratio indicated in table), solvent (2 mL). ^b Total conversion of methyl oleate and conversion to oleic acid (hydrolysis). ^c Only 0.43 mL of acetone. ^d Reaction carried out for 24 h at atmospheric pressure and 12 h under vacuum. ^e Acetone was evaporated after 6 h and the reaction was continued without solvent.

Table 2. Results of the transesterification of methyl oleate with TMP using basic catalysts.^a

Entry	Oleate/TMP molar ratio	Catalyst (%) ^b	P	T (°C)	time (h)	Conv. (%)	Selectivity (%)		
							ME	DE	TE
1	1:4	EtONa (1%)	atm	140	24	22	90	10	0
2	1:4	EtONa (1%)	1 mbar	140	6	94	61	37	2
3	1:3	EtONa (1%)	1 mbar	140	6	95	67	31	2
4	4:1	EtONa (1%)	1 mbar	140	6	71	0	12	88
5	4:1	EtONa (5%)	1 mbar	140	6	78	0	6	94
6	1:1	MeONa (1%)	1 mbar	140	6	64	16	49	35
7	4:1	MeONa (1%)	1 mbar	140	6	60	5	7	88
8	3.2:1	MeONa (1%)	1 mbar	120	6	30	41	38	21
9	3.2:1	MeONa (1%)	50 mbar	120	6	15	100	0	0
10	3.9:1	MeONa (0.7%)	60 mbar	130	6	7	87	13	0

^a Scale: 2 g of methyl oleate (6.75 mmol). ^b % w/w with respect to methyl oleate.

Lowering the reaction temperature (entry 8) and increasing the pressure (entries 9 and 10) produce a significant drop in activity, as demonstrated by the lower conversions, with an improved selectivity to ME. Thus, the use of EtONa was chosen as the best transesterification method to perform the reaction from the fish oil fatty ethyl esters mixture.

Composition of the fatty esters mixture from fish oil

The starting material for Solutex is the fish oil, in the form of ethyl esters (FAEEs), and the PUFAs are extracted and concentrated following sustainable methodologies. After this process, Solutex generates 500-600 ton/year of a waste mixture of FAEEs that is currently valorized as biodiesel. The composition of this waste mixture is determined by gas chromatography (Table 3, Figure S1). As can be seen, it contains around 50 wt% of saturated FAs (SFAs), mainly palmitic acid but also myristic and stearic ones. Monounsaturated FAs (MUFAs), mainly palmitoleic and oleic, account for 25 wt%, whereas the content of polyunsaturated acids is also around 25 wt%, with eicosapentaenoic acid (EPA) as the most abundant one. In the $^1\text{H-NMR}$ spectrum (Figure S2), the total count of protons is around 34-36, with an average of 1-1.5 insaturations per fatty acid chain. Thus, the averaged molecular weight of the sample, used for calculation of molar ratios, was taken as that of methyl oleate ($\text{C}_{19}\text{H}_{36}\text{O}_2$).

Table 3. Fatty acid ethyl esters composition of fish oil waste.

Type	Fatty acid	% w/w
SFAs	Myristic (C14:0)	13.5
	Palmitic (C16:0)	29.6
	Stearic (C18:0)	3.8
MUFAs	Palmitoleic (C16:1 n-7)	13.7
	Oleic (C18:1 n-9)	9.1
	C18:1 n-7	3.6
PUFAs	C16:4 n-1	4.1
	Linoleic (C18:2 n-6)	1.4
	Stearidonic (C18:4 n-3)	3.6
	EPA (C20:5 n-3)	11.0
	DHA (C22:6 n-3)	1.8

Others	4.8
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Transesterification of the fatty esters mixture from fish oil

The results were determined by ^1H NMR as explained above. One full spectrum is shown in Figure S3.

In a first set of experiments, the reactions were carried out with 5 g of the mixture of fatty ethyl esters at 140°C under vacuum (1 mbar) for 6 h. The effect of the fatty esters/TMP molar ratio was studied at a fixed amount of EtONa as catalysts, 1% w/w with respect to the mass of esters. The results are summarized in Figure 2. As can be seen, the conversion of FAEEs is nearly quantitative ($>95\%$) up to a FAEEs/TMP molar ratio of 2.7:1, slightly lower than the stoichiometric amount. As expected, the major products at low ratio are ME (66% selectivity), whereas the amount of DE and TE increases at higher ratios. The use of the stoichiometric amount (3:1) leads to a slight decrease in the conversion, only 86% that can reach 93% at longer reaction time (8 h), with the highest selectivity to TE (79%).

Those optimal conditions of time and molar ratio were used to check the effect of the amount of catalyst (Figure 3). As expected, a lower amount (0.7% w/w) leads to a slower reaction, with both lower conversion and lower selectivity to TE. However, rather unexpectedly, the use of a higher amount of sodium ethoxide is also detrimental for the final result. This effect had been already observed in other transesterification reactions and it was attributed to the presence of free fatty acids traces (Anastopoulos et al., 2009).

Some of those reactions were scaled up (Table 4). In a first step, they were conducted at 50 g scale with different FAEEs/TMP molar ratios, trying to obtain mixtures enriched in ME, ME+DE and TE. Although TE has commonly been the target molecule to be used as lubricant, the presence of partially esterified molecules ME and DE has been described as positive for the lubrication properties (Abu Bakar et al., 2007), and hence the optimization of their synthesis would also be of interest. The conversions were quantitative or almost quantitative, slightly better than those obtained at lower scale. The selectivities were very similar, leading to a mixture highly enriched in TE (82%, entry 1), in ME (83%, entry 3), or in ME+DE (43%+41%, entry 2). The results were nearly identical in a further step in 100 g scale (entries 4 and 5). However, in 500 g scale the results were slightly worse, with only 90% conversion and 60% TE selectivity

after 12 h. The increase in viscosity of the reaction mixture and the problems in mixing with a magnetic stirrer may account for this behavior, indicating that a different type of reactor would be necessary for such scale.

Finally, the objective was to obtain a mixture with a selectivity to TE higher than 90% in order to check its suitability as lubricant. The use of a higher excess of fatty esters (molar ratio 4:1) under the same conditions (140°C, 1 mbar, 1% w/w EtONa) did not improve the selectivity to TE. However, when the reaction was carried out at 100°C, 96% TE selectivity was obtained with 84% conversion after only 3 h. A possible explanation for this result might be the partial elimination of some of the reagents (mainly ethyl myristate) under vacuum at higher temperature, which also would explain the conversion higher than the theoretical maximum (75%), although this might also be due to an incorrect estimation of the averaged molecular weight of the fatty esters mixture.

Table 4. Scaling up results of the transesterification of the mixture of FAEEs with TMP.^a

Entry	Scale	FAEEs/TMP molar ratio	time (h)	Conv. (%)	Selectivity (%)		
					ME	DE	TE
1	50 g	3:1	8	96	1	17	82
2	50 g	1.5:1	6	100	43	41	16
3	50 g	0.5:1	1	100	83	16	1
4	100 g	3:1	8	98	2	19	79
5	100 g	1.5:1	6	100	46	41	13
6	500 g	3:1	12	90	8	32	60

^a With 1% w/w EtONa, at 140°C and 1 mbar.

Properties of the TMP fatty esters from fish oil

The possible use as lubricant is determined by the physical properties of the mixtures. Among them, viscosity and its change with temperature (viscosity index) are considered the most important ones for lubrication behavior, together with others, such as pour point and flash point, that mainly control the temperature limits of usage of the biolubricants (Mobarak et al., 2014; Zainal et al., 2018). In fact, the lubricants are classified according to their kinematic viscosities at 40°C in the range of 2-3200 cSt,

that covers all the possible applications. All of these properties of the TMP esters were measured, as well as those of the starting mixture of ethyl esters for the sake of comparison. The results are collected in Table 5.

Table 5. Properties of the starting ethyl esters mixture and the TMP esters.

Property	Mixture of esters	
	Ethyl	TMP
Density at 20°C (g/cm ³)	0.862	0.919
Kinematic viscosity at 40°C (cSt)	3.75	23.16
Viscosity ISO-VG	3	22
Kinematic viscosity at 100°C (cSt)	1.60	5.81
Viscosity index	-	211
Flash point (°C)	176	204
Pour point (°C)	-11	-6

Kinematic viscosity was determined at 40°C and 100°C. Values were significantly higher in the case of TMP esters, compared with those of the ethyl esters mixture, from 3.75 to 23.16 cSt at 40°C and from 1.6 to 5.8 cSt at 100°C. The ISO 3448 viscosity classification for industrial oils is based on the kinematic viscosity at 40°C. The starting mixture of fatty ethyl esters shows an ISO VG 3, whereas TMP esters have an ISO VG 22.

These values are compared with those reported in the literature for mixtures of TMP esters from vegetal oils (Table 6). All the kinematic viscosity values are higher than the result obtained with waste fish oil. The range of kinematic viscosities at 40°C lays from 30 to 50, depending on the starting material. Most of them are between 30 and 40 cSt, Jatropha oil (Nurdin et al., 2015), rapeseed oil (Uosukainen et al., 1998; Gryglewicz et al., 2003; Gryglewicz et al., 2013), olive oil (Gryglewicz et al., 2003), for rubber seed oil (Kamalakar et al., 2013), and even lard (Gryglewicz et al., 2003). Only canola biodiesel leads to a mixture of TMP esters with 40.5 cSt (Sripada et al., 2013) and the maximum value (50.6 cSt) is obtained with an oleic enriched fraction of sal fat oil (Kamalakar et al., 2015). Those values correspond to ISO VG 32 up to ISO VG 46. Something similar happens with kinematic viscosities at 100°C, showing values between 7.8 and 10, with the only exception of Jatropha, whose kinematic viscosity is lower than that of waste fish oil. One possible reason for this behavior is the larger

fraction of light fatty acids (60% with ≤ 16 C) in the fish oil mixture, in comparison with only 5-18% of C16 fatty acids in vegetal oils.

4. Conclusions

Fatty acid ethyl esters (FAEEs) obtained from fish oil after PUFAs extraction can be efficiently transformed in TMP esters by transesterification, using sodium ethoxide (EtONa) as catalyst. The selectivity to ME, DE and TE can be modulated by changing the transesterification conditions. The physical properties of a high TE sample, with kinematic viscosity in the ISO VG 22 range, indicate that this waste mixture might be useful in the preparation of bio-lubricants, with properties complementary to those of the TMP esters obtained from vegetal oils, in the ISO VG 32 up to ISO VG 46 range, that may open the way to its use as hydraulic oil.

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Table 6. Comparison of the properties of different mixtures of TMP esters

Starting material	Kinematic viscosity (cSt)		Viscosity index	Pour point (°C)	Flash point (°C)	Reference
	40°C	100°C				
Waste fish oil	23.16	5.81	211	-6	204	This work
Jatropha biodiesel	32.29	5.14	81	-5	n.r. ^a	Nurdin et al., 2015
Rapeseed oil	32.9	n.r. ^a	220	-41	n.r. ^a	Uosukainen et al., 1998
Rapeseed oil	35.34	7.99	209.2	-15.5	n.r. ^a	Gryglewicz et al., 2003
Olive oil	36.00	8.32	218.3	-13	n.r. ^a	Gryglewicz et al., 2003
Lard	37.65	9.06	234.7	-10.5	n.r. ^a	Gryglewicz et al., 2003
Rapeseed oil	38.2	8.4	205	-18	n.r. ^a	Gryglewicz et al., 2013
Rubber seed oil	38.4	8.6	212	-6	299	Kamalakar et al., 2013
Canola biodiesel	40.5	7.8	204	-66	n.r. ^a	Sripada et al., 2013
Oleic rich sal fat	50.6	9.98	190	-3	258	Kamalakar et al., 2015

^a n.r. = not reported.

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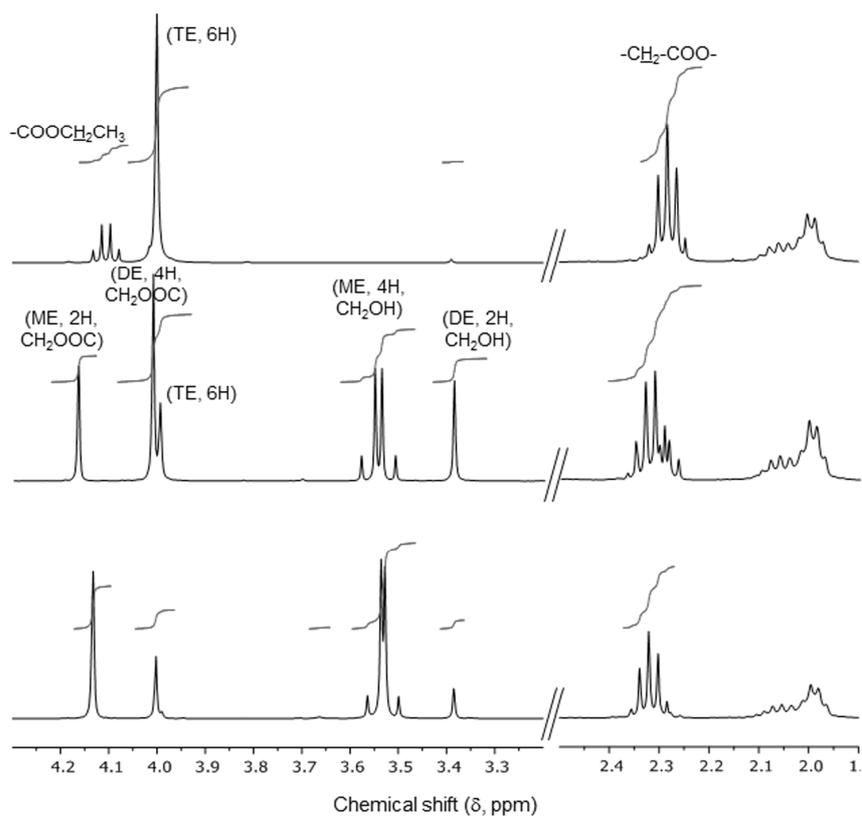
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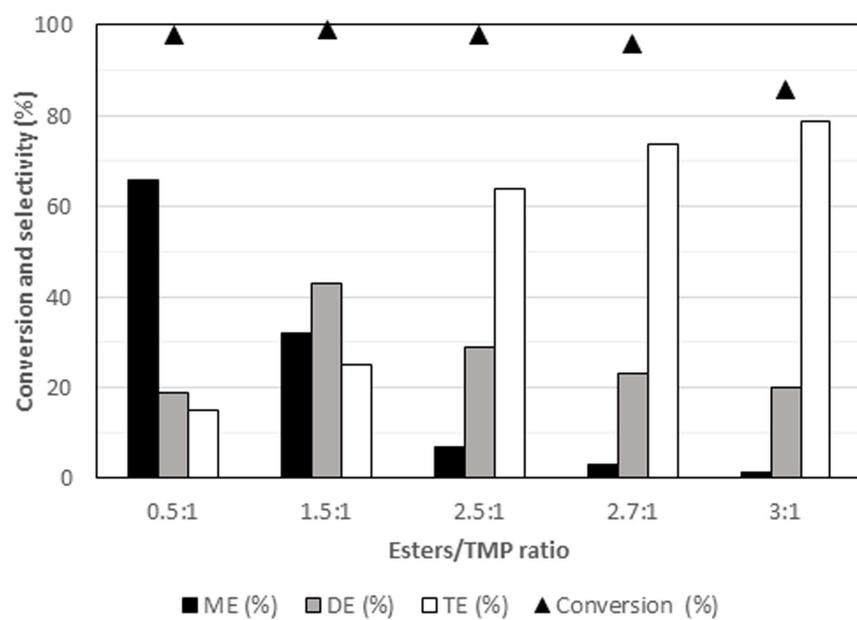
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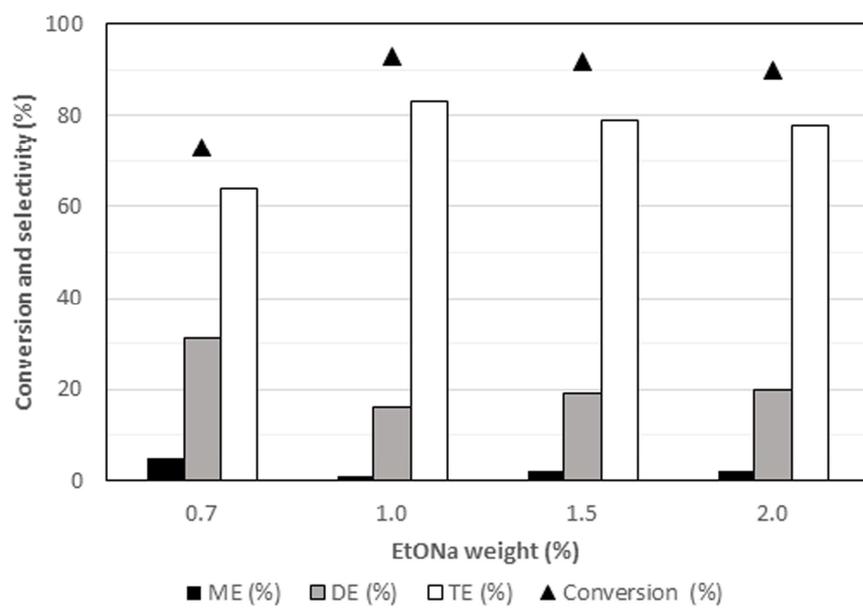
Figure 1. $^1\text{H-NMR}$ spectra of different reaction mixtures of waste fish oil FAEEs with TMP. Top: incomplete conversion of ethyl esters and nearly only TE formed. Medium: mixture of ME and DE (with a small amount of TE). Bottom: mainly ME (with a small amount of DE).

Figure 2. Effect of FE/TMP molar ratio. Scale: 5 g of FAEEs. Conditions: 1% EtONa (w/w FAEEs), 140°C, 6 h, 1 mbar.

Figure 3. Effect of the amount of EtONa (w/w FAEEs). Scale: 5 g of FAEEs. Conditions: FAEEs/TMP molar ratio 3:1, 140°C, 8 h, 1 mbar.







Highlights

- Waste fish oil fatty esters of trimethylolpropane have been prepared.
- Common lipases are not efficient for this transesterification.
- Sodium ethoxide allows obtaining trimethylolpropane esters.
- Selectivity can be tuned with the reaction conditions.
- Trimethylolpropane triesters show interesting physical properties as bio-lubricants.