



Superior Solubility and Dissolution of Zaltoprofen via Pharmaceutical Cocrystals

Farmasötik Cocrystal ile Zaltoprofen'in Üstün Çözünürlük ve Çözünmesi

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ABSTRACT

Objectives: Pharmaceutical cocrystals are a promising tool to enhance the solubility and dissolution of poorly soluble drugs. Zaltoprofen (ZFN) is nonsteroidal anti-inflammatory drug with a prevalent solubility problem. The present study was undertaken to enhance the solubility and dissolution of ZFN through pharmaceutical cocrystals by screening various coformers.

Materials and Methods: Cocrystals of ZFN were prepared in 1:1 and 1:2 ratio of drug:coformer by the dry grinding method. The melting point and solubility of the crystalline phase were determined. The potential cocrystals were characterized by differential scanning calorimetry (DSC), infrared spectroscopy, and powder X-ray diffraction (PXRD). Cocrystals were subjected to dissolution rate and stability study.

Results: ZFN-nicotinamide (NIC) cocrystals demonstrated deviation in melting point and solubility. The cocrystals were obtained in both 1:1 and 1:2 ratios with NIC. The infrared analysis noticeably indicated the shifting of characteristic bands of ZFN. The crystallinity of the cocrystals was evident from the XRPD pattern and notable difference in the 2θ values of intense peaks. The DSC spectra of the cocrystals exhibited altered endotherms analogous to melting point. The cocrystals showed a faster dissolution rate and a 55% increase in the extent of dissolution compared to pure drug. The cocrystals were stable at room temperature and accelerated conditions.

Conclusion: The prepared cocrystals exhibited greater solubility and dissolution compared to the pure drug and were stable at room temperature and accelerated conditions.

Key words: Pharmaceutical cocrystal, zaltoprofen, solubility, dissolution

ÖZ

Amaç: Farmasötik kokristal, zayıf çözünür ilaçların çözünürlüğünü ve çözünmesini arttırmak için umut veren bir araçtır. Zaltoprofen (ZFN) yaygın çözünürlüğe sahip nonsteroid antiinflamatuar ilaçtır. Bu çalışma, çeşitli koformerlerin taranması yoluyla farmasötik kokteyli aracılığıyla ZFN'nin çözünürlüğünü ve çözünmesini arttırmak için üstlenilmiştir.

Gereç ve Yöntemler: Kuru öğütme yöntemi ile 1:1 ve 1:2 oranında ilaç:koformer oranında ZFN kristalleri hazırlanmıştır. Erime noktası ve kristalin fazın çözünürlüğü belirlenmiştir. Potansiyel kristaller differansiyel tarama kalorimetrisi (DSC), kızılötesi spektroskopi ve toz X ışını kırınımı (PXRD) ile karakterize edilmiştir. Kokristaller çözünme hızına ve stabilize çalışmasına tabi tutulmuştur.

Bulgular: ZFN-nikotinamid (NIC) kokristal erime noktasında ve çözünürlükte sapma göstermiştir. Kristaller, NIC ile hem 1:1 hem de 1:2 oranında elde edilmiştir. Kızılötesi analizi, ZFN karakteristik bantlarının kaymasını belirgin bir şekilde göstermiştir. Kristallerin kristallenmesi XRPD paterninden belirgin olarak görülmüştür ve 2θ değerindeki yoğun zirvelerdeki kayda değer farklılıklar gözlenmiştir. Kristallerin DSC spektrumları, erime noktasına benzer değiştirilmiş endoterm sergilemiştir. Kristaller, daha hızlı çözünme oranı ve saf ilaçla karşılaştırıldığında çözünme derecesinde % 55 artış göstermiştir. Kristaller, oda sıcaklığında ve hızlandırılmış koşullarda kararlı bulunmuştur.

Sonuç: Hazırlanan kristaller, saf ilaca kıyasla daha fazla çözünürlük ve çözünme sergilemiş ve oda sıcaklığında ve hızlandırılmış koşullarda sabit bulunmuştur.

Anahtar kelimeler: Farmasötik kokristal, zaltoprofen, çözünürlük, çözünme

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INTRODUCTION

After oral administration the solubility and dissolution rate of a drug are crucial factors for its sufficient bioavailability. These factors are the main challenge to the formulation scientist for the development and formulation of effective drugs. More than 40% of drugs in development suffer from bioavailability problems owing to poor solubility. Alternative strategies have been introduced to enhance solubility, the dissolution rate, and bioavailability. These involve salt formation, solid dispersion, cyclodextrin complexation, microemulsification, solubilization, micronization, etc.¹⁻⁴

Recently pharmaceutical cocrystals have attracted considerable attention from formulation experts busy in formulation development. Due to the inherent thermodynamic stability of crystalline active pharmaceutical ingredients (APIs), these are preferred in the pharmaceutical industry. Pharmaceutical cocrystals have emerged as an effective tool to tailor the physical properties of APIs like solubility and dissolution along with stability. The principal advantage of this technique is that the pharmacological effect of the drug remains unchanged.⁵⁻⁷ Cocrystals are defined as stoichiometric multicomponent systems united by noncovalent interactions in which two diverse components are solid under ambient conditions. The documented advantages of cocrystals are improved stability against humidity, chemical stability, improved dissolution and bioavailability, and tabletability. Various methods were studied to enhance solubility like hydrotropy and solid dispersion. To the best of our knowledge, pharmaceutical cocrystals of zaltoprofen (ZFN) have not been reported to date.⁸⁻¹⁴

ZFN is a nonsteroidal anti-inflammatory propionic acid class drug. It is used in the treatment of acute and chronic inflammation and rheumatoid arthritis. It is practically insoluble in water and associated with side effects like ulcerogenicity, bellyache, and indigestion. Moreover, ZFN is weakly ionizable and so salt formation cannot enhance the solubility of the drug. Rapid onset and improved bioavailability are desired for analgesics. Hence there is a strong scientific and clinical need to prepare novel forms of ZFN possessing modified solubility and dissolution rates that can be formulated for oral administration. Accordingly, the aim of the present study was to prepare novel pharmaceutical cocrystals of ZFN with improved solubility and dissolution.^{15,16}

MATERIALS AND METHODS

Materials

ZFN was received as a gift sample from ICPA Laboratory Ltd. (Mumbai, India). All other chemicals were purchased from the SD Fine Chemicals (Mumbai, India). Double distilled water was used throughout the research.

Preparation of cocrystals

The dry grinding method was adopted for the preparation of ZFN cocrystals. The drug and coformers were mixed in different molar ratios (1:1 and 1:2) in a mortar and pestle for 45 min to form cocrystals. They were dried overnight at ambient

temperature and stored in tight containers.¹⁷ Twenty-five coformers were screened for the preparation of cocrystals, i.e. salicylic acid, nicotinamide (NIC), glutaric acid, malonic acid, benzoic acid, tartaric acid, oxalic acid, citric acid, urea, succinic acid, saccharine sodium, Pluronic 68 AR, magnesium stearate, crotonic acid, P-hydroxy benzoic acid, caffeine, 3,5 dihydroxy benzoic acid, piperazine citrate, cinnamic acid, adipic acid, hydroquinone, isonicotinic acid, acetamide, maleic acid, and ascorbic acid.

Evaluation of cocrystals

Drug content

Cocrystal powder equivalent to 10 mg of drug was accurately weighed and dissolved in a 10 mL volumetric flask and the volume was adjusted with phosphate buffer pH 6.8. The resulting solution was filtered, suitably diluted, and the absorbance of the solution was measured at 243 nm (Shimadzu UV 1800).¹⁸

Determination of melting points

Melting points of the compounds were determined using a digital melting point apparatus (Labtronics Ltd).

Saturation solubility

An excess amount of pure drug and cocrystals were dissolved in 10 mL vials containing the drug to estimate solubility. The vials were agitated on rotary shaker and allowed to stand for equilibration for 24 h. The samples were filtered after 24 h, suitably diluted with distilled water, and analyzed by UV spectrophotometer at 243 nm.

Infrared spectroscopy

Infrared (IR) spectroscopy was employed to determine the possible interaction between the drug and coformers. Samples were mixed with potassium bromide and compressed into discs before scanning between 4000 and 400 cm⁻¹ with resolution of 4 cm⁻¹ by Shimadzu IR spectrophotometer.

Differential scanning calorimetry

The thermal behavior of the drug alone and cocrystals was determined on a Mettler Toledo DSC 822e Module. Weighed samples were loaded into an aluminum pan before crimping and heated at a rate of 5°C/min, covering the 0 to 300°C temperature range, under a nitrogen stream. The instrument was calibrated using indium and an empty aluminum pan was used as a reference.

Powder X-ray diffraction

Silicon sample holders were utilized to get diffraction patterns for pure ZFN and cocrystals (Bruker D8 Advance diffractometer). The instrument was equipped with a fine focus X-ray tube and each sample was placed onto a goniometer head that was motorized to permit spinning of the sample during data acquisition.

In vitro dissolution study

Pure ZFN and its cocrystals were subjected to dissolution study by USP type II apparatus (Electrolab, Mumbai, India). The dissolution study was performed in 900 mL of pH 6.8

phosphate buffer at $37 \pm 0.5^\circ\text{C}$ and 50 rpm for 60 min. The pure drug and cocrystals equivalent to 80 mg of drug were used for the study. Then 5 mL samples were withdrawn after specified time intervals and analyzed by UV spectrophotometer at 243 nm.¹⁹

Stability study

The selected cocrystals were subjected to a stability study at room temperature and $40 \pm 2^\circ\text{C}$ with $75 \pm 5\%$ RH for 3 months. A sample of 1 g was placed in an eppendorf tube in a stability chamber throughout the stability duration and analyzed after 30 days, 60 days, and 90 days. Different attributes were studied to assess the stability, i.e. drug content, melting point, solubility, *in vitro* drug release, etc.

RESULTS AND DISCUSSION

The 25 coformers were screened to prepare cocrystals with ZFN by the dry grinding method. The coformers were selected based on a literature survey and to increase the chances of formation of new cocrystals. Among the various coformers studied, NIC successfully interacted with ZFN, giving novel cocrystal forms. The obtained ZFN cocrystals were subjected to evaluation and stability studies.

Drug content

The drug content of ZFN-NIC 1:1 and 1:2 cocrystals was determined in phosphate buffer pH 6.8 as $95.87 \pm 0.98\%$ and $95.88 \pm 1.10\%$, respectively.

Melting points and saturation solubility

The melting points of pure drug, coformers, and cocrystals were estimated and are reported in Table 1. In addition, the saturation solubility of pure drug and cocrystals was also determined and is reported in Table 1. These parameters were used for preliminary screening of the cocrystals. The melting points of the ZFN-NIC cocrystals were lower than that of the pure drug. This may be attributed to the multicomponent system and the probable formation of cocrystals. The altered melting points might be due to an interaction between ZFN and NIC, modified crystallinity of molecules, or distinct packing arrangement. This interaction results in an altered molecular arrangement, which leads to novel crystal forms with distinct physical properties.^{20,21}

The solubility of a few cocrystals was improved but ZFN-NIC cocrystals exhibited a remarkable increase in solubility, indicating successful interaction of drug and coformer. However, greater solubility was obtained with ZFN-NIC 1:2 cocrystals (1.516 ± 0.467 mg/mL) than with 1:1 (0.926 ± 0.134 mg/mL). The ZFN-NIC 1:1 and 1:2 cocrystals showed 42-fold and 66-fold increases in solubility in comparison to the pure drug. The results were compared using Dunnet's test and statistically significant differences were found in solubility ($p < 0.05$) between the pure drug and cocrystals. This indicates an interaction between ZFN and NIC leading to cocrystal formation. The interaction between the oxygen atom of the drug and the primary amide hydrogen of the NIC might have formed the cocrystal. Similar studies were reported on cocrystals of meloxicam, lornoxicam, aceclofenac, etc.^{22,23}

On the basis of the results, ZFN-NIC 1:1 and 1:2 cocrystals were further characterized and confirmed.

IR spectroscopy

The IR spectra for the pure drug, coformer, and ZFN cocrystals were recorded and are shown in Figures 1 and 2. The principle bands were identified and related changes were recorded. The IR spectrum of pure ZFN shows the presence of the characteristic peaks, which were recorded at 1699 cm^{-1} and 1668 cm^{-1} for stretching of the carboxylic group, $-\text{C}-\text{S}-\text{C}-$ aromatic stretching peaks observed at 939.39 cm^{-1} , OH stretching in the carboxylic group at 2950 cm^{-1} , and CH_3 stretching at 1330 cm^{-1} . The IR spectrum of NIC revealed an absorption band at 3145 cm^{-1} for NH_2 stretching of primary amide and 3342 cm^{-1} for the pyridine ring region, NH bending is observed at 1593 cm^{-1} , and aromatic $\text{C}=\text{C}$ peaks are observed at 1614 cm^{-1} . These spectra are in good agreement with the published data. The IR bands were significantly changed in the cocrystals in comparison to the pure drug and coformer, indicating an interaction between drug and coformer.²⁴

In the case of the 1:1 cocrystal changes were observed in the peaks corresponding to carboxylic group stretching, which was observed at 1634 cm^{-1} , OH stretching at 3000 cm^{-1} in comparison to the drug, and NH_2 stretching and NH bending at 3450 and 1583 cm^{-1} as compared to NIC and 1654 cm^{-1} , 3000 cm^{-1} , and 1583 cm^{-1} for the 1:2 cocrystal, respectively.

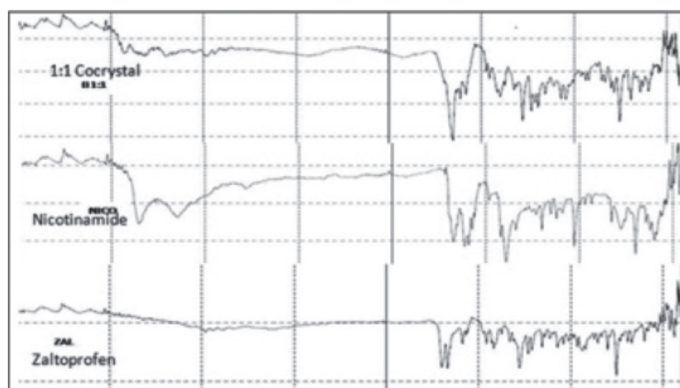


Figure 1. Overlay IR spectra of 1:1 cocrystal

IR: Infrared spectroscopy

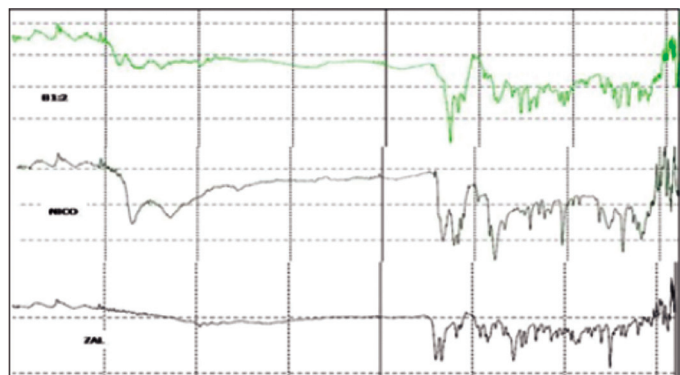


Figure 2. Overlay IR spectra of 1:2 cocrystal

IR: Infrared spectroscopy

A new peak at 3450 cm^{-1} and one at 3400 cm^{-1} were observed, indicating the formation of a hydrogen bond between the drug and coformer in the ZFN-NIC 1:1 and 1:2 cocrystals, respectively, prepared by the neat grinding method.²⁵

Similar changes in the IR spectra of other drugs like piroxicam and hydrochlorothiazide were reported and considered as a sign of cocrystal formation.^{26,27} Hence the changes recorded in the present study can be regarded as an indicator of cocrystal formation between the drug and coformer.

Differential scanning calorimetry

ZFN, NIC, and ZFN-NIC cocrystals were characterized by DSC. The pure drug and NIC showed characteristic endothermic peaks at 137.69°C and 129.67°C , respectively, corresponding to their melting points. Similar thermal behavior was reported for the drug and coformer.²⁸

ZFN-NIC (1:1 and 1:2) cocrystals exhibited melting points at 109.20°C and 123.50°C , respectively, which are significantly different from that of the pure drug. Moreover, the peak onset for the pure drug was obtained at 131.52°C and at 102.40°C and 120.02°C for 1:1 and 1:2 cocrystals, respectively.

The changes in the thermal properties were reported as evidence for the formation of cocrystals.²⁹ Hence the present investigation indicates the formation of cocrystals (Figure 3).

Powder X-ray diffraction

The powder X-ray diffraction (PXRD) patterns for ZFN, NIC, and ZFN-NIC cocrystals are shown in Figures 4 and 5. The materials in the powder state give different peaks of varying intensities at certain positions. The diffractogram of the ZFN showed characteristic numerous sharp, intense diffraction peaks at

Table 1. Melting point and solubility of cocrystals

Drug/potential cocrystal	Melting point of coformer ($^\circ\text{C}$)	Cocrystal melting point (1:1) ($^\circ\text{C}$)	*Cocrystal solubility (mg/mL) (1:1)	Cocrystal melting point (1:2) ($^\circ\text{C}$)	*Cocrystal solubility (mg/mL) (1:2)
ZFN	133-135		0.022 \pm 0.005		
ZFN-Salicylic acid	158-159	133.5	0.452 \pm 0.078	135	0.445 \pm 0.095
ZFN-Nicotinamide	122-124	128	0.926 \pm 0.134	121	1.516 \pm 0.467
ZFN-Glutaric acid	96	119.5	0.0136 \pm 0.0089	120.5	0.0819 \pm 0.023
ZFN-Malonic acid	130	113.5	0.0083 \pm 0.093	116.5	0.0077 \pm 0.001
ZFN-Benzoic acid	122	103	0.416 \pm 0.098	100	0.721 \pm 0.278
ZFN-Tartaric acid	164-167	146.5	0.0218 \pm 0.013	146.5	0.0103 \pm 0.008
ZFN-Oxalic acid	99	117	0.0147 \pm 0.017	115.5	0.0261 \pm 0.009
ZFN-Citric acid	148-150	135.5	0.0098 \pm 0.00097	134.5	0.0080 \pm 0.002
ZFN-Urea	131	129.5	0.367 \pm 0.067	128.5	0.228 \pm 0.090
ZFN-Succinic acid	184	152	0.0287 \pm 0.008	153	0.0210 \pm 0.007
ZFN-Sodium saccharine	226-230	126.5	0.154 \pm 0.069	181	0.207 \pm 0.067
ZFN-Pluronic 68 AR	53-54	65.5	0.152 \pm 0.089	62.5	0.139 \pm 0.083
ZFN-Magnesium stearate	88.5	91.5	0.761 \pm 0.284	126.5	0.536 \pm 0.132
ZFN-Crotonic acid	74-75	124	0.064 \pm 0.016	95	0.0714 \pm 0.021
ZFN-Phydroxy benzoic acid	208	160	0.820 \pm 0.349	164.5	0.980 \pm 0.230
ZFN-Caffeine	238	175.5	0.354 \pm 0.078	152	0.435 \pm 0.098
ZFN-3,5 dihydroxy benzoic acid	236-238	184.5	0.256 \pm 0.086	187.5	0.372 \pm 0.068
ZFN-Piperazine citrate	183-187	197	0.160 \pm 0.067	199.5	0.179 \pm 0.043
ZFN-Cinnamic acid	132-134	113	0.339 \pm 0.129	117.5	0.440 \pm 0.065
ZFN-Adipic acid	151-154	141	0.0282 \pm 0.009	143	0.0202 \pm 0.009
ZFN-Hydroquinone	172.3	123.5	0.109 \pm 0.008	124.5	0.0597 \pm 0.006
ZFN-Isonicotinic acid	310	284	0.0950 \pm 0.021	292	0.103 \pm 0.089
ZFN-Acetamide	79-81	98.5	0.0246 \pm 0.026	92	0.0232 \pm 0.007
ZFN-Maleic acid	135	124.5	0.0700 \pm 0.039	130.5	0.0785 \pm 0.013
ZFN-Ascorbic acid	190	158	0.0593 \pm 0.009	160	0.0525 \pm 0.013

ZFN: Zaltoprofen, *Average of three determinations mean \pm standard deviation

different 2θ values (15, 17.5, 19, 31, 32.5, and 42), indicating a crystalline nature. In addition, the diffraction peaks obtained for NIC were 25, 30, 34.5, 37, 47.5, and 50.5 2θ values. Similar diffraction patterns were reported in previous investigations. The PXRD pattern of the cocrystal was distinguishable from that of its components and some additional diffraction peaks

appeared that did not exist in the pure drug or coformer. The additional diffraction peaks for 1:1 and 1:2 cocrystals were obtained at 2θ values of 16, 17, 18, 19, 20, 30.5, and 37.5 and 17.5, 18.5, 26.5, 34.5, 37.5, 40.5, and 50.5, respectively. The appearance of new diffraction peaks in the diffractogram of cocrystals shows the formation of a new crystalline phase

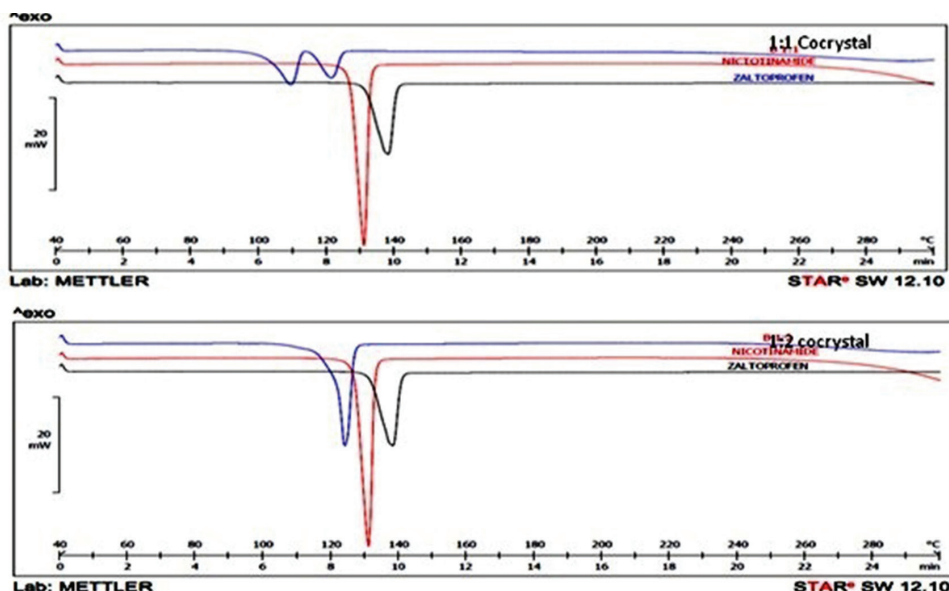


Figure 3. Overlay DSC thermogram of 1:1 and 1:2 cocrystals

DSC: Differential scanning calorimetry

Table 2. Stability study of cocrystals

Parameters	Sampling	Zaltoprofen (accelerated)	Zaltoprofen (room temperature)	Cocrystal neat grinding (accelerated)		Cocrystal neat grinding (room temperature)	
				1:1	1:2	1:1	1:2
Melting point (°C)	Initial	133-134	133-134	111-114	120-121	111-114	120-121
	1 month	133-134	133-135	114-115	115-117	116-117	124-126
	2 month	132-133	131-133	115-118	117	116-118	115-117
	3 month	132-134	131-134	114-116	116	115-116	116-118
Solubility (mg/mL)	Initial	0.01513	0.0151	0.926	1.516	0.9261	1.516
	1 month	0.01518	0.0131	1.016	1.202	1.077	1.159
	2 month	0.01464	0.0136	1.126	1.268	1.126	1.308
	3 month	0.01445	0.0131	0.913	1.070	1.020	1.149
<i>In vitro</i> dissolution (%)	Initial	52.01	52.01	98.32	98.89	98.32	98.89
	1 month	64.76	65.49	98.78	94.51	99.26	99.45
	2 month	64.81	64.81	98.74	98.39	98.74	98.39
	3 month	64.06	63.63	99.53	99.74	98.69	98.81
Drug content	Initial	-	-	95.87	95.88	95.87	95.88
	1 month	-	-	95.70	95.3	95.74	95.57
	2 month	-	-	95.2	94.96	95.47	95.90
	3 month	-	-	95.00	95.05	95.08	95.59

(cocrystals). The formation of cocrystals based on the PXRD pattern was reported and showed new peaks that differ from the peaks corresponding to its input components.^{30,31}

In vitro dissolution study

The dissolution rate plays a crucial role in the bioavailability of drugs with poor solubility. The dissolution experiment was

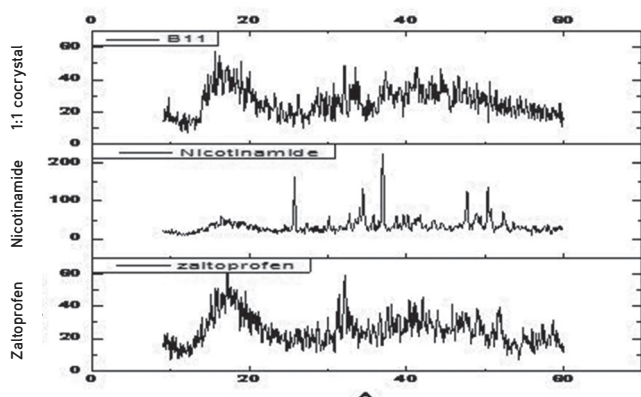


Figure 4. Overlay PXRD pattern for 1:1 cocrystal (1:1 cocrystal)

PXRD: Powder X-ray diffraction

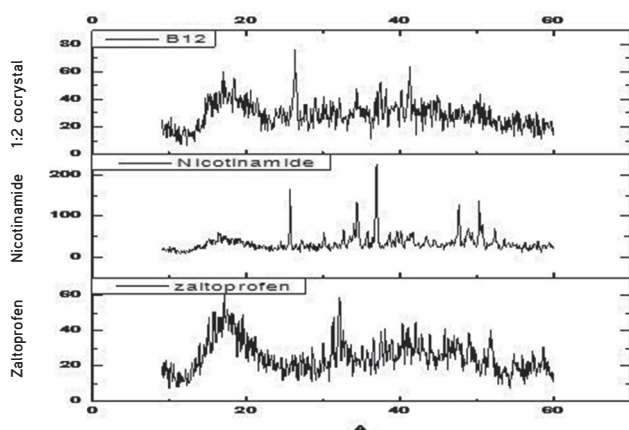


Figure 5. Overlay PXRD pattern for 1:2 cocrystal (1:2 cocrystal)

PXRD: Powder X-ray diffraction

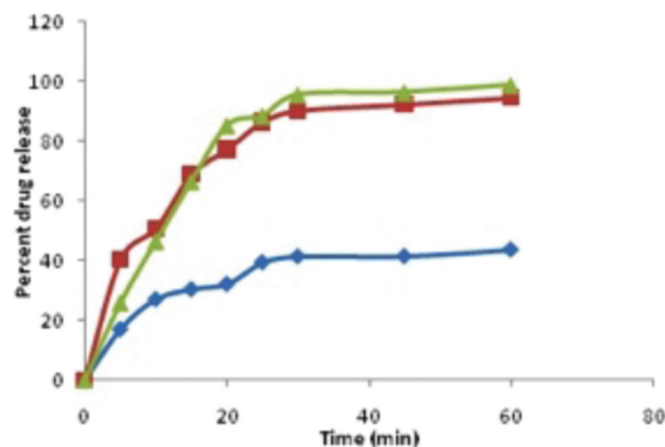


Figure 6. *In vitro* drug release

conducted on the pure drug and cocrystals. The dissolution profile of the pure drug and the prepared cocrystals is shown in Figure 6. The dissolution profile of the pure drug indicates a slow dissolution rate with only $27.17 \pm 0.89\%$ of the drug being dissolved in the first 10 min. The total amount of drug dissolved in 60 min was $43.82 \pm 1.06\%$ and the calculated dissolution efficiency was only 27.4%. However, cocrystals of the ZFN resulted in a substantial increase in the dissolution rate. The amount of drug dissolved in first 10 min was $50.66 \pm 0.32\%$ and $46.67 \pm 0.65\%$ for the 1:1 and 1:2 cocrystals, respectively. The maximum amount of drug dissolved was $98.89 \pm 0.48\%$ for the 1:2 cocrystal with dissolution efficiency of 86.71%, whereas it was $94.14 \pm 0.91\%$ for the 1:1 cocrystal, having a dissolution efficiency of 81.78%. This can indicate a weaker crystalline structure of the formed cocrystal as evident from the higher dissolution rate. Moreover, greater dissolution of ZFN from the cocrystal can be attributed to enhanced solubility of the cocrystal in the dissolution media. Cocrystallization had been well documented as a suitable technique for dissolution enhancement.³² The similarity factor test denoting the dissolution of pure drug showed dissimilarity to the prepared cocrystals (F2 value 20% and 22% for 1:1 and 1:2 cocrystals).

Stability study

The drug and cocrystals were subjected to a stability study at room temperature and accelerated conditions for 3 months to assess the stability of cocrystals. All the cocrystals were stable at both storage conditions and no substantial change in the estimated parameters like melting point, solubility, *in vitro* drug release, and drug content was obtained except ZFN:NIC 1:2 cocrystal solubility at accelerated conditions. However, the pure drug exhibited changes in solubility and percent dissolution during the stability period, indicating instability. Hence cocrystal stability was enhanced in comparison to the pure drug. This demonstrates the potential of cocrystals to improve drug stability. Similar results have been reported for theophylline.³³ The results are given in Table 2.

CONCLUSIONS

Dry grinding of ZFN with NIC resulted in cocrystal formation. This was ascertained by melting point transformations, DSC changes, shifts in infrared bands, and changes in 2θ values in XRPD that mutually supported each other. The newly prepared cocrystals exhibited greater solubility and dissolution as compared to the pure drug and were stable at room temperature and accelerated conditions. The study endorsed the high potential of the technique for future applications with other drugs.

Conflict of interest: No conflict of interest declared by authors.

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