

## Magnetic Properties of Magnesium Substituted Y-Type Hexaferrite

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**Abstract-** In the present work, we have synthesized Y-type hexaferrite  $Ba_2Mg_2Al_{x/2}Cr_{x/2}Fe_{12-x}O_{22}$  ( $x = 0.0, 0.5, 1.0$ ) using solution gel method. The structural and magnetic properties have been investigated using XRD, FTIR and VSM. X-ray study shows that grain size lies in nano range. Presence of two prominent peaks near  $550-600\text{ cm}^{-1}$  to  $500-600\text{ cm}^{-1}$  in FTIR indicates the formation of hexaferrite. From VSM analysis, It is found that with increasing the amount of co-dopants ( $Al - Cr$ ) the saturation magnetization, retentivity and increases for  $x = 0$  and  $x = 0.5$  it was observed that coercivity is zero which indicates that material is super-paramagnetic in nature.

### 1. INTRODUCTION

Development of ferrites materials is one of the major advances in the field of the nanotechnology. The Magnesium substituted hexaferrite has great importance for the modern society because they are component of a variety of electronic devices used in means of medical equipment, communication, distribution and generation of electric energy, automotive and magnetic recording [1]. Y-type hexagonal ferrite is a frontier material for these applications and has attractive for researchers for many years. This type of ferrite is composed of Y-type Magnesium hexagonal, maintaining a hexagonal structure, with the easy direction of magnetization along the c-axis. However,  $Co_2Y$ ,  $Mg Y$  (where Co and Mg are present as divalent cations) are known as ferroplana ferrites because they have an easy plane (basal plane) of magnetization perpendicular to the c-axis [2].

To synthesize Nano-ferrites there are lot of methods like mechanical grinding [3], glass crystallization [4], hydrothermal reaction sintering [5], micro- emulsion technique [6], route etc. In addition, many synthetic techniques such as combustion method [7], spray pyrolysis [8], ammonium nitrate melt [9], co-precipitation [10], Sol-gel [11], oxidation in nitric acid [12] have been exploited extensively. The sol-gel method is one of the most widely used methods for the formation of pure hexagonal ferrites. Using this method, it is possible to obtain to order of submicron particles for attaining useful properties for nano materials.

In the present work, we have investigated the effects of Mg substitution for  $Ba^{2+}$  in barium based Y-type hexagonal ferrites ( $Ba_2Mg_2Al_{x/2}Cr_{x/2}Fe_{12-x}O_{22}$ ) synthesized by sol-gel auto combustion method. The structural and magnetic properties have been investigated using X-ray diffraction (XRD) Fourier Transform Infrared Spectroscopy (FTIR) and Vibrating sample magnetometer (VSM).



## 2. CHARACTERIZATION TECHNIQUES

Y-type hexagonal ferrites ( $Ba_2Mg_2Al_{x/2}Cr_{x/2}Fe_{12-x}O_{22}$ ) powders have been synthesized via sol gel combustion method using AR grade Barium Nitrate ( $Ba(NO_3)_2$ ), Chromium Nitrate Nonahydrate ( $Cr(NO_3)_3 \cdot 9H_2O$ ), Aluminium Nitrate Nonahydrate ( $Al(NO_3)_3 \cdot 9H_2O$ ), Magnesium Nitrate Hexahydrate ( $Mg(NO_3)_2 \cdot 6H_2O$ ), Ferricnitrate Nonahydrate ( $Fe(NO_3)_3 \cdot 9H_2O$ ) and Monohydrate Citric Acid ( $C_6H_8O_7 \cdot H_2O$ ) citric acid. Aqueous solution of iron and metal salts are prepared separately in stoichiometric proportions by dissolving them in distilled water. Aqueous solution of citric acid is added to the salt solution with cations to citric acid molar ratio of 1:1.5. Ammonium hydroxide solution is further added drop by drop to attain the 7 pH value. Then the solution is heated at 80C–85C for 4–6 hours with continuous stirring using magnetic stirrer. After evaporation of water, the liquid gets converted into a homogenous brown colored gel. The viscous solution is dried over hot plate at 280–300 C for 3 hours to form the precursor material. Presintering is done at 500 C for 2 h. Then the precursor material is calcined at 700 C, 900 C, 1100 C and 1200 C for 5 h. Following devices were used to know the properties of synthesized Y- type hexaferrites.

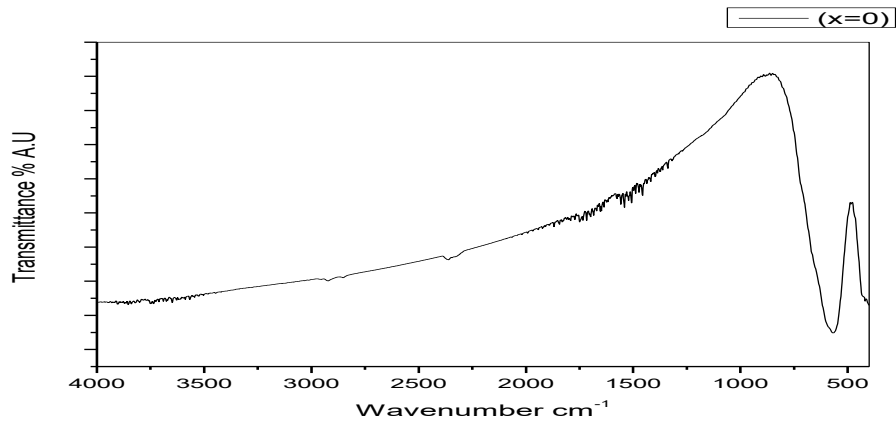
## 3. RESULTS AND DISCUSSION

### 3.1 Structural properties

#### 3.1.1 FTIR Analysis

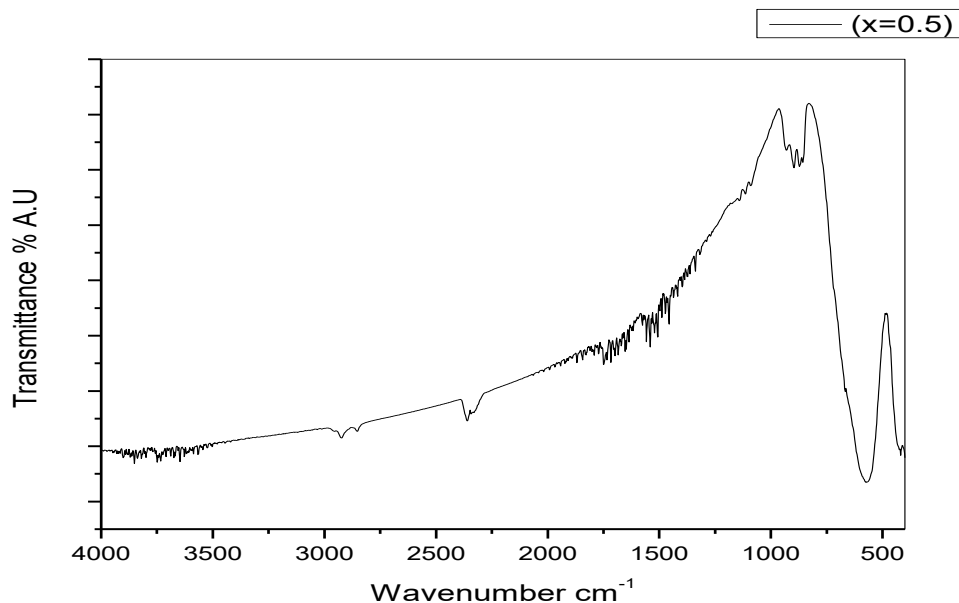
Fourier Transform Infrared Spectroscopy is a technique that helps to know the chemical, structural and phase changes present in the sample during combustion and sintered process. Nearly 2 mg of ferrite powder is mixed with powdered  $KBr$  in the ratio 1:100 by weighing to ensure uniform dispersion. The infrared spectra of prepared sample  $Ba_2Mg_2Al_{x/2}Cr_{x/2}Fe_{12-x}O_{22}$  was recorded in the range of  $4000\text{ cm}^{-1}$  to  $400\text{ cm}^{-1}$ .

For the sample  $Ba_2Mg_2Fe_{12}O_{22}$  spectra is shown in the Fig 1.1. The band is in the range of  $4000\text{--}400\text{ cm}^{-1}$ . They have dominant absorption bands at range of  $400\text{--}600\text{ cm}^{-1}$ ,  $1250\text{--}1600\text{ cm}^{-1}$  and  $3600\text{--}4000\text{ cm}^{-1}$ . The bands in the range of  $400\text{--}600\text{ cm}^{-1}$  assigned to be for broadening of metal oxygen bonds which is the characteristics of hexaferrites. Further, the band in the range of  $1250\text{--}2000\text{ cm}^{-1}$  indicates the presence of carboxylic acid (citric acid) and nitrile. The downward peak at  $2400\text{ cm}^{-1}$  is due to presence of  $OH^-$ . The band from range  $3400\text{--}4000\text{ cm}^{-1}$  exhibit the stretching of hydroxyl group which indicates the presence of free or absorbed moisture in the sample.



**Fig 1.1: FT-IR Spectra of  $Ba_2Mg_2Fe_{12}O_{22}$  at pH=7**

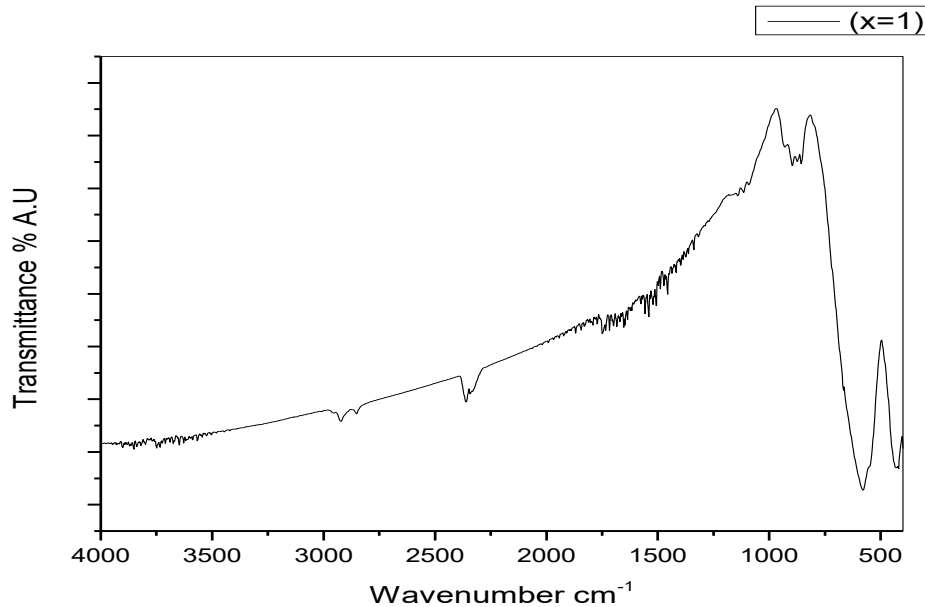
Similarly Spectra of sample  $Ba_2Mg_2Al_{0.25}Cr_{0.25}Fe_{11.5}O_{22}$  is shown in Fig 1.2. The band is in the range of  $4000-400\text{ cm}^{-1}$ . It has dominant absorption bands at ranges of  $400-600\text{ cm}^{-1}$ ,  $1250-1600\text{ cm}^{-1}$ , and  $3500-3900\text{ cm}^{-1}$ . The bands in the range of  $400-600\text{ cm}^{-1}$  are assigned to be for broadening of metal oxygen bonds, which is characteristic of hexaferrites. Further, the band in the range of  $1250-1700\text{ cm}^{-1}$  indicates the presence of carboxylic acid (citric acid) and nitrile. The downward peak at  $2400\text{ cm}^{-1}$  and  $2800\text{ cm}^{-1}$  is due to the presence of  $OH^-$ . The band from range  $3400-4000\text{ cm}^{-1}$  exhibits the stretching of hydroxyl group, which indicates the presence of free or absorbed moisture in the sample.



**Fig 1.2: FT-IR Spectra of  $Ba_2Mg_2Al_{0.25}Cr_{0.25}Fe_{11.5}O_{22}$  at pH=7**

Similarly, Spectra of sample  $Ba_2Mg_2Al_{0.5}Cr_{0.5}Fe_{11}O_{22}$  is shown in Fig 1.3. The band is in the range of  $4000-400\text{ cm}^{-1}$ . It has dominant absorption bands at ranges of  $400-600\text{ cm}^{-1}$ ,  $1250-1600\text{ cm}^{-1}$ , and  $3500-3900\text{ cm}^{-1}$ . The bands in the range of  $400-600$

$\text{cm}^{-1}$  is assigned to be for broadening of metal oxygen bonds which is the characteristics of hexaferrites. Further, the band in the range of  $1250\text{--}1800\text{ cm}^{-1}$  indicates the presence of carboxylic acid (citric acid) and nitrile. The downward peak at  $2400\text{ cm}^{-1}$  and  $2900\text{ cm}^{-1}$  is due to presence of  $\text{OH}^-$ . The band from range  $3600\text{--}4000\text{ cm}^{-1}$  exhibit the stretching of hydroxyl group which indicates the presence of free or absorbed moisture in the sample.



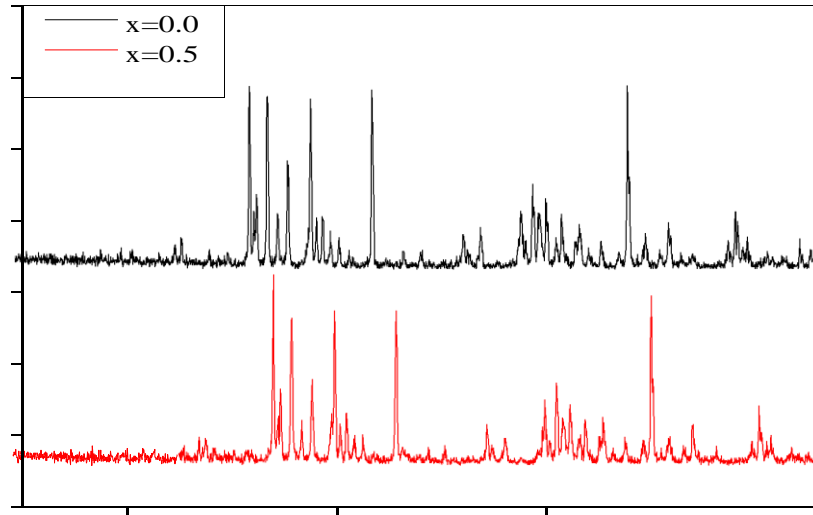
**Fig 1.3: FT-IR Spectra of  $\text{Ba}_2\text{Mg}_2\text{Al}_{0.5}\text{Cr}_{0.5}\text{Fe}_{11}\text{O}_{22}$  at pH=7**

### 3.1.2 XRD Analysis

Fig 1.4 shows the X-ray diffraction pattern of samples  $\text{Ba}_2\text{Mg}_2\text{Al}_{x/2}\text{Cr}_{x/2}\text{Fe}_{12-x}\text{O}_{22}$  ( $x = 0, 0.5, 1.0$ ) nanoparticles. Pure hexaferrites phase was the formation at a temperature of  $1000^\circ\text{C}$  for 6 hrs. The average crystallite size of the Nano-crystalline hexaferrite synthesized at different molar ratios of substitution for the peaks was calculated from the XRD data using the Scherrer formula.

$$D = \frac{k\lambda}{\beta \cos \theta}$$

Where  $D$  is the crystallite size. The shape factor of  $K$  can be  $0.62\text{--}2.08$  and is usually taken as about  $0.89$  for hexagonal systems.  $\lambda$  is the X-ray wavelength ( $1.54\text{\AA}$ ). It was observed and calculated and was found out to be  $44.9434\text{ nm}$  for  $x = 0$ ,  $57.9684\text{ nm}$  for  $x = 0.5$  and  $38.379\text{ nm}$  for  $x = 1$ .



**Fig 1.4: XRD Patterns of hexaferrite  $Ba_2Mg_2Al_{x/2}Cr_{x/2}Fe_{12-x}O_{22}$  at different composition**

The lattice parameter ( $a$  and  $c$ ) and cell volume ( $V$ ) was also calculated using formula given below.

$$\frac{1}{d_{hkl}^2} = \frac{4}{3} \left[ \frac{h^2 + hk + k^2}{a^2} \right] + \frac{l^2}{c^2}$$

$$V = 0.8666a^2c$$

With usual parameter. The lattice parameters of all the samples synthesized at different concentrations were calculated and is shown in the Table 1.1. It was observed that lattice parameter ' $c$ ' decreases and lattice parameter ' $a$ ' increases with increase in dopants.

**Table 1.1: Value of different parameter calculated for hexaferrite**

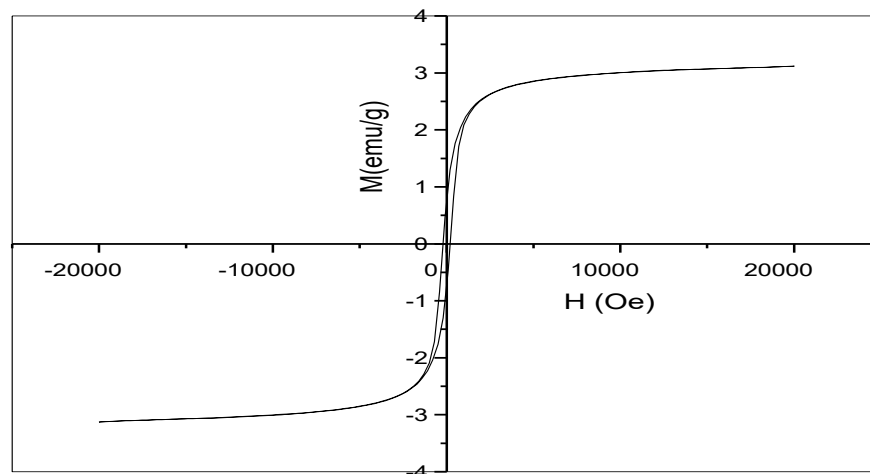
S.No.	Sample	Lattice constant $a$ (Å)	Lattice constant $c$ (Å)	Cell Volume (Å <sup>3</sup> )
1	$x = 0$	4.40	47.86	802.956
2	$x = 0.5$	5.85	44.432	1317.12
3	$x = 1$	5.87	39.76	1187.24

### 3.1.3 VSM Analysis

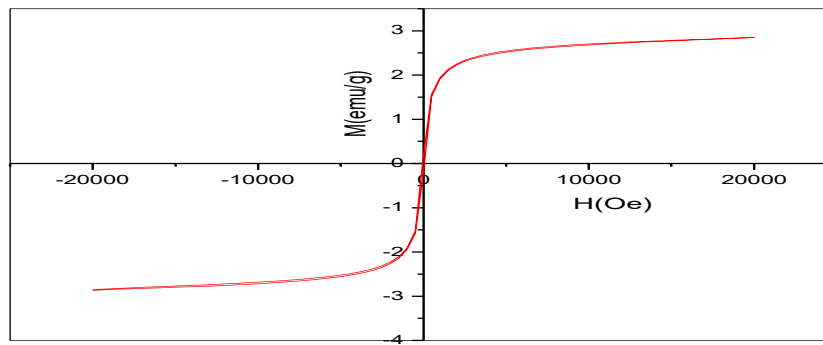
Figure 1.5 to 1.7 shows the room temperature hysteresis loop plots for all of the Nano size samples with the chemical formula of  $Ba_2Mg_2Al_{x/2}Cr_{x/2}Fe_{12-x}O_{22}$  sintered at 1000°C for 6hrs. The values of saturation magnetization ( $M_s$ ) and coercivity ( $H_c$ ) for all samples were accomplished from the loops, and variation of these parameters as a function of aluminum and chromium contents. According to Figure it is clear that by an increase in substitution content, the saturation magnetization was gradually increased. For  $x=0$  and  $x=0.5$  the value of coercivity is zero, which says nanomaterial is super-paramagnetic in nature whereas for  $x=1$  it is 838.7025. The retentivity also gets increasing with increase in substitution contents. The value of Retentivity, Coercivity and Saturation Magnetization is shown in Table 1.2 for each sample.

**Table 1.2:** The value of Retentivity, Coercivity, and saturation magnetization for ( $x=0, 0.5$ , and 1)

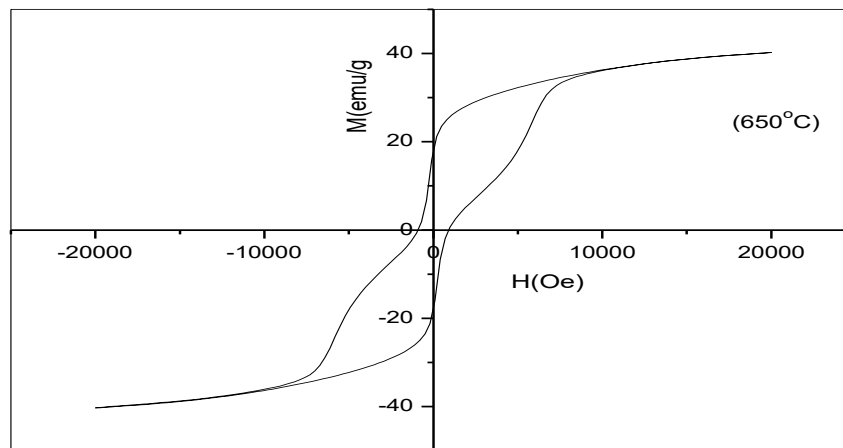
S. No	Samples ( $x$ )	Retentivity ( $M_r$ ) <i>emu / g</i>	Coercivity ( $H_c$ ) ( <i>Oe</i> )	Saturation Magnetization ( $M_s$ ) ( <i>emu / g</i> )
1	$x = 0$	4609.39	0	19441.4095
2	$x = 0.5$	5183.01011	0	19796.5037
3	$x = 1$	9628.5168	838.7025	21510.5163



**Fig.1.5:** Magnetic-Hysteresis loop for  $x=0$



**Fig 1.6: Magnetic-Hysteresis loop for  $x=0.5$**



**Fig 1.7: Magnetic-Hysteresis loop for  $x=1$**

## 5. CONCLUSIONS

The Y-type Nano-hexaferrites have been synthesized using sol-gel technique and was characterized using FT-IR, XRD and VSM. X-ray study shows that grain size lies in Nano range. Presence of two prominent peaks near  $550\text{-}600\text{ cm}^{-1}$  to  $400\text{-}600\text{ cm}^{-1}$  in FTIR indicates the formation of hexaferrite. From VSM analysis, It is found that with increasing the amount of co-dopants ( $Al-cr$ ) the saturation magnetization, retentivity and increases for and it was observed that coercivity is 0 for  $x=0$  and  $x=0.5$  which indicates that material is super-paramagnetic in nature.

## REFERENCES

1. R. B. Jotania, H.S. Virk, *Solid State Phenomena*. 189 209(2012).
2. Z. Haijun, Yao Xi, Zhang Liangying, *J. Eur. Ceram. Soc.* 22 835 (2002).

3. Z. Jin, W. Tang, J. Zhang, H. Lin, Y. Du, *Journal of Magnetism and Magnetic Materials*.182 231, **(1998)**.
4. L. Rezlescu, E. Rezlescu, P.D. Popar, N. Rezlescu, *Journal of Magnetism and Magnetic Materials*. 193 288 **(1993)**.
5. X. Liu, J. Wang, L.M. Gan, S.C. Ng, *Journal of Magnetism and Magnetic Materials*.195 452 **(1999)**.
6. X. Liu, J. Wang, L.M. Gan, S.C. Ng, J. Ding, *Journal of Magnetism and Magnetic Materials*,184 344 **(1998)**.
7. L. Junliang, Z. Wei, G. Cuijing, Z. Yanwei, *Journal of Alloys and Compounds*. 479 863**(2009)**.
8. P. Ren, J. Guan, X. Cheng, *Materials Chemistry and Physics*, 98 90 **(2006)**.
9. H. Sözeri, I. Kucuk, H. Ozcan, *Journal of Magnetism and Magnetic Materials*. 323 1799 **(2011)**.
10. M. Radwan, M.M. Rashad, M.M. Hessien, *Journal of Materials Processing Technology*, 181 106**(2007)**.
11. T. Kaur, B. Kaur, B. H. Bhat, S. Kumar, A.K. Srivastva,*Physica B*, 456 206 **(2015)**.
12. H. Sözeri, *Journal of Magnetism and Magnetic Materials*, 321 2717 **(2009)**.