

Notes

Synthesis, Spectroscopic and Ligand Field Properties of *trans*-Difluorobis(2,2-dimethylpropane-1,3-diamine)chromium(III) Perchlorate

Jong-Ha Choi,* Keon Sang Ryoo, and Sang Hak Lee†

Department of Chemistry, Andong National University, Andong 760-749, Korea. *E-mail: jhchoi@andong.ac.kr

†Department of Chemistry, Kyungpook National University, Daegu 702-701, Korea

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The preparation and characterization of novel chromium(III) complexes have received much attention because the chromium(III) materials are promising materials for tunable solid state laser in the spectral region between 600 and 1100 nm.¹ As a prerequisite for these applications, a detailed study of the spectroscopic and ligand field properties is needed. Furthermore, the geometrical and conformational isomerism in transition metal complexes has been an area of intense activity and has provided much basic structural information.² The 2,2-dimethylpropane-1,3-diamine (Me₂tn) as a bidentate ligand can coordinate to a central metal ion forming a six-membered ring, and the spectroscopic properties of the resulting metal complexes can be compared with those of the corresponding complexes containing the unsubstituted tn diamine system.³

The preparation, spectroscopic properties and crystal structures of the series diacidochromium(III) complex containing Me₂tn as a bidentate ligand have been reported extensively.⁴⁻⁸ The [Cr(Me₂tn)₂L₂]⁺ (L = monodenate) cation can exist as *trans* and *cis* geometric isomers. In addition, there are two possible conformations with respect to the six-membered rings in the *trans* isomer. The carbon atoms of the two chelate rings of the Me₂tn ligands can be located on the same side (*syn* conformer) or on opposite side (*anti* conformer) of the equatorial plane. In the case of *trans*-[Cr(Me₂tn)₂Br₂]₂Br₂HClO₄·H₂O and *trans*-[Cr(Me₂tn)₂Cl₂]₂Cl, independent *syn* and *anti* conformational isomers were found within the crystal lattice^{4,5} while structural analyses of *trans*-[Cr(Me₂tn)₂(NCS)₂]₂NCS·1/2H₂O, *trans*-[Cr(Me₂tn)₂Cl₂]₂ClO₄ and *trans*-[Cr(Me₂tn)₂Cl₂]₂ZnCl₄ indicated that two chelate rings of Me₂tn ligands are only in the *anti* chair-chair conformation.⁶⁻⁹ The different arrangements of the two six-membered chelate rings of Me₂tn ligands may be dependent on the packing forces and counter anions in the crystal structure. The infrared, electronic absorption and emission spectroscopic properties are useful in distinguishing the geometric isomers of chromium(III) complexes. In particular, sharp-line electronic spectroscopy can be used to probe the inequivalent sites of chromium(III) complexes because the zero phonon line and its splitting are very sensitive to the exact bond angles around the metal ion.¹⁰

By refluxing *trans*-[Cr(py)₄F₂]₂ClO₄ with the 2,3,2-tet and 3,2,3-tet in 2-methoxyethanol, the stereochemistry is retentive.

However, in the case of 2,2,2-tet, a *cis* arrangement of the fluoro ligand reported to occur. The last fact suggests to us a *trans* to *cis* isomerization of starting materials.¹¹

In this paper, we report on the synthesis, spectroscopic and ligand field properties of newly prepared *trans*-[Cr(Me₂tn)₂F₂]₂ClO₄. The geometric configuration of the complex is proposed on the basis of its electronic and infrared spectra.

Infrared spectroscopy is useful in assigning the configuration of *cis* and *trans* isomers of octahedral chromium(III) complexes.¹²⁻¹⁴ The FT-infrared spectrum is presented in Figure 1.

The absorption near 3438 cm⁻¹ and broadness of the band indicate that there are hydrogen bonds in this complex. The infrared spectrum shows strong bands in the region 3250 - 3100 cm⁻¹ and 2960 - 2850 cm⁻¹ due to the symmetric and anti-symmetric N-H and C-H stretching modes, respectively. The four main absorptions in the N-H and in the C-H stretching regions are indicative of the *trans* configuration. The infrared absorption spectrum of *trans* isomer reveals a simpler pattern than that of the *cis* isomer. This pattern may be rationalized on the basis of the higher symmetry of the *trans* isomer. Particularly, the NH₂ bending at approximately 1600 cm⁻¹ and the CH₂ rocking bands at 900 - 800 cm⁻¹ can discriminate between two geometric isomers of CrN₄X₂ⁿ⁺ typed complexes.¹²⁻¹⁴ The absorption band in the 1500 - 1600 cm⁻¹ region due to the asymmetric deformation mode of the NH₂ group is split in the case

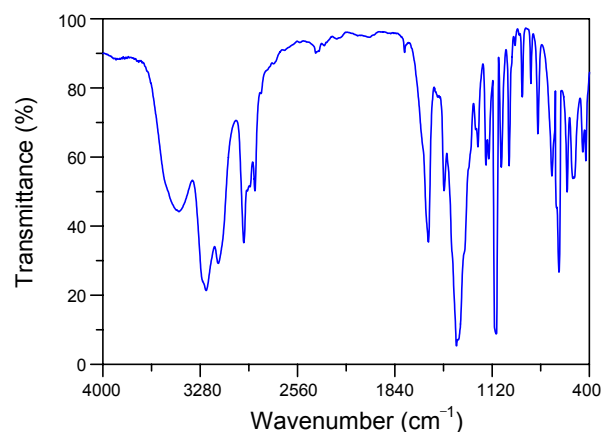


Figure 1. The mid-infrared spectrum of *trans*-[Cr(Me₂tn)₂F₂]₂ClO₄.

of the *cis* isomer, while the *trans* isomer exhibits only a single band in this region. The one strong absorption band at 1591 cm^{-1} and the one medium band at 1475 cm^{-1} can be assigned to NH_2 and CH_2 bending modes. The C-N stretching frequency occurs as a strong band at 1383 cm^{-1} . The two absorptions at 1166 and 1144 cm^{-1} , and two strong absorptions at 1051 and 994 cm^{-1} are due to the NH_2 and CH_2 twisting modes, respectively. The very strong absorptions at 1095 cm^{-1} and 626 cm^{-1} are due to the $\nu_a(\text{Cl-O})$ stretching and $\delta_d(\text{OCIO})$ bending modes of the perchlorate anion. The broadening or split of the perchlorate peaks results from the reduction in the symmetry of the ClO_4^- ion to C_{3v} or C_{2v} due to its interaction with the amine hydrogen atoms and the formation of partial H-bonding. A weak absorption at 940 cm^{-1} is also due to $\nu_s(\text{Cl-O})$ stretching mode. The bands $\nu_{as}(\text{ClO}_4^-)$ and $\delta_d(\text{ClO}_4^-)$ of the perchlorate anion are not significantly affected by differing metal ions.¹⁴

It has been suggested that the *cis* isomer exhibits at least three bands in the $900 - 820\text{ cm}^{-1}$ region due to the NH_2 rocking mode while the methylene vibration splits into two peaks in the $760 - 820\text{ cm}^{-1}$ region. However, the *trans* isomer shows two groups of bands, one band near 890 cm^{-1} arising from the amine vibration and a doublet near 800 cm^{-1} due mainly to the methylene vibration. The title complex exhibits two bands at 896 and 833 cm^{-1} in the NH_2 rocking frequency region. The CH_2 rocking bands at 781 cm^{-1} are observed.

Metal-ligand stretching bands occur in the far infrared range. There are also consistent differences between the infrared spectra of the *cis* and *trans* isomers of diacidobis(diamine) chromium(III) complexes in the region $560 - 400\text{ cm}^{-1}$. All the *trans* complexes show three bands in this region, whereas all the *cis* complexes show four strong bands. In the far-IR spectral range of *trans*- $[\text{CrN}_4\text{F}_2]^+$ complex with D_{4h} symmetry, group theory analysis predicts two Cr-N and one Cr-F infrared-active stretching vibrations, respectively. The one strong peak at 567 cm^{-1} can be assigned to the Cr-F stretching mode by comparing the IR spectra of the corresponding dichloro, dibromo and diisothiocyanatobis(2,2-dimethylpropane-1,3-diamine)chromium(III) complexes which the bands are absent.⁵⁻⁷ The title complex shows a pattern of three bands: one strong band at 514 cm^{-1} and two other strong bands at 448 and 426 cm^{-1} due to the Cr-N stretching modes.¹⁵ The infrared spectral properties

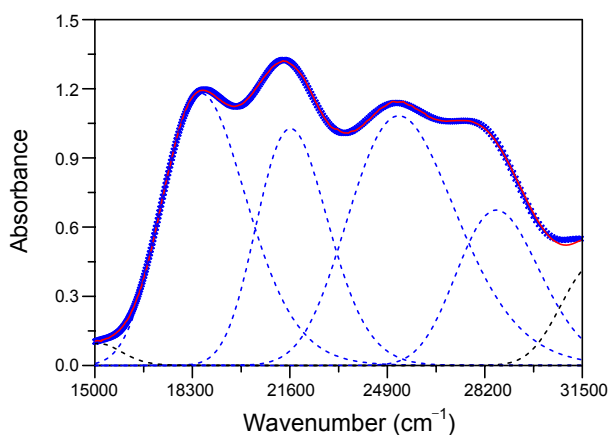


Figure 2. The visible absorption spectrum and its resolved overlapping peaks (dotted curves) of *trans*- $[\text{Cr}(\text{Me}_2\text{tn})_2\text{F}_2]^+$ in aqueous solution.

of the novel complex would strongly support the assignment of *trans* configurational geometry.

The assignment of geometrical configuration is also suggested by inspection of the d-d absorption spectra.^{16,17} The position of the spin-allowed transitions in the electronic spectra, the number of bands, and their coefficients are usually reliable indicators for distinguishing the *cis* and *trans* geometrical isomers. The more symmetrical *trans* chromium(III) complexes with two fluoro and Me_2tn ligands have four symmetric bands in the visible region, and these bands are located at higher wavelengths and have lower extinction coefficients than those of less symmetrical *cis*-isomers. The visible absorption spectrum (solid line) of $[\text{Cr}(\text{Me}_2\text{tn})_2\text{F}_2]^+$ in aqueous solution at room temperature is represented in Figure 2.

The observation of four quartet bands in the *d-d* transition region is indicative of the *trans* configuration. In order to have some point of reference for the splitting and position of the bands, we have fitted the band profiles to four Gaussian curves, as seen in Figure 2.¹⁸ The contribution from outside bands was corrected for fine deconvolution. A deconvolution procedure on the experimental band pattern yielded maxima at 18600 , 21620 , 25290 and 28590 cm^{-1} for the noncubic (D_{4h}) split levels ($^4E_g^a + ^4B_{2g}$ and $^4A_{2g} + ^4E_g^b$) of $^4T_{2g}$ and $^4T_{1g}(\text{O}_h)$, respectively. These band positions were used as the observed spin-allowed transition energies in the ligand field calculation.

In the framework of the angular overlap model the metal-ligand interactions are described in terms of localized bonding parameters of σ - and π -type. Throughout the ligand field optimization, we assumed the tetragonal symmetry for the title complex. The parameters varied during the optimization were the interelectronic repulsion parameter B plus the AOM parameters $e_\sigma(\text{F})$ and $e_\pi(\text{F})$ for the fluoride-chromium, and $e_\sigma(\text{N})$ for the Me_2tn nitrogen-chromium. The π -interaction of the fluoride with the metal ion was considered to be isotropic. The π -interaction of amine nitrogens with sp^3 hybridization in the Me_2tn ligand was assumed to be negligible. However, it is noteworthy that the peptide nitrogen with sp^2 hybridization has a weak π -donor character.¹⁹ The Powell parallel subspace optimization procedure²⁰ was used to find the global minimum. The results of the optimization and the parameter set used to generate the best-fit energies are listed in Table 1.

The following values were finally obtained for the ligand field parameters: $e_\sigma(\text{N}) = 7175$; $e_\sigma(\text{F}) = 7510$; $e_\pi(\text{F}) = 1710$ and

Table 1. Observed band positions and calculated transition energies (cm^{-1}) in the electronic spectrum of *trans*- $[\text{Cr}(\text{Me}_2\text{tn})_2\text{F}_2]^+$

Assignment	Observed ^a	Calculated ^b
$^4B_{1g} \rightarrow ^4E_g^a$	18600	18600
$^4B_{1g} \rightarrow ^4B_{2g}$	21620	21619
$^4B_{1g} \rightarrow ^4A_{2g}^a$	28590	28592
$^4B_{1g} \rightarrow ^4E_g^b$	25290	25289
$^4B_{1g} \rightarrow ^4A_{2g}^b$?	40395
$^4B_{1g} \rightarrow ^4E_g^c$?	43275

^aObtained from the Gaussian component deconvolution, $^b e_\sigma(\text{N}) = 7175$; $e_\sigma(\text{F}) = 7510$; $e_\pi(\text{F}) = 1710$; $B = 660\text{ cm}^{-1}$

Table 2. The AOM parameters (cm^{-1}) for $\text{trans-[Cr(Me}_2\text{tn)}_2\text{L}_2]^+$

L	$e_\sigma(\text{Me}_2\text{tn})$	$e_\pi(\text{Me}_2\text{tn})$	$e_\sigma(\text{L})$	$e_\pi(\text{L})$	Ref
F^-	7 175	0	7 510	1 710	This work
Cl^-	7 358	0	5 263	859	6
Br^-	7 281	0	4 983	946	5
NCS^-	7 132	0	6 160	32	7

$B = 660 \text{ cm}^{-1}$. The AOM parameters are plausible and reproduce the spectrum pretty well. Table 2 contains the comparison of the AOM parameters with the values reported in the literature⁵⁻⁹ for the $\text{trans-[Cr(Me}_2\text{tn)}_2\text{L}_2]^+$ with anionic ligands. A ligand field analysis of absorption spectrum indicates that the fluoride is a strong σ - and π -donor. The $e_\sigma(\text{N})$ value of 7175 cm^{-1} for amine nitrogen of Me_2tn is located in the normal range, and slightly smaller than other amine nitrogens.²¹ The strong coordination of two fluorine atoms to axial position may be caused a slight weakness of the Cr-N(Me_2tn) bond strength in equatorial coordination plane. However, it is suggested that the four nitrogen atoms of the Me_2tn ligand still have strong σ -donor properties toward chromium(III) ion. The value of Racah parameter, B is about 72% of the value for a free chromium(III) ion in the gas phase.¹⁶ These factors plus AOM parameters can be used in determining the photolabilized ligand of the photoreaction in the chromium(III) complexes and predicting the relative efficiency of the 3d-4f energy transfer in the heterometal dinuclear complexes.^{22,23}

In conclusion, the infrared and electronic spectroscopic properties of the newly prepared complex show that the chromium atom is an octahedral environment, coordinated by two bidentate Me_2tn ligands and two fluorine atoms in *trans* positions. The ligand field analysis shows that the fluoride is strong σ - and π -donor and the nitrogen atoms of the Me_2tn ligand also have a strong σ -donor properties toward the chromium(III). Current studies in our laboratory on this compound indicate that the suitable crystals for X-ray analysis are not isolated until now. The mass spectrum, 2D [^1H , ^{13}C] HSQC spectrum and crystal structural work including magnetic and conductance data will be reported later.

Experimental Section

The UV-visible absorption spectrum was recorded with a HP 8453 diode array spectrophotometer. The mid-infrared spectrum was obtained with a Mattson Infinities series FT-IR spectrometer using a KBr pellet. Analysis for C, H, and N was performed on a Carlo Erba 1108 Elemental Vario EL analyzer. The ligand field parameters were calculated using the AOMX program.²⁴ The ligand, 2,2-dimethylpropane-1,3-diamine was purchased from Aldrich Chemical Company. All other chemicals were of reagent grade or better quality and used without further purification.

Caution! Although we have experienced no difficulty with the perchlorate salt of the complex described in this paper, this should be regarded as a potentially explosive compound and treated with care.

As starting material, $\text{trans-[Cr(py)}_4\text{F}_2]\text{ClO}_4$ was prepared

according to the literature.¹¹ The complex $\text{trans-[Cr(py)}_4\text{F}_2]\text{ClO}_4$ (1.0 g) was dissolved in 2-methoxyethanol (8 mL). 2,2-Dimethyl-1,3-propanediamine (0.4 g) was added to this solution, which is then heated under reflux for 45 min. The resulting red solution was concentrated and the residue was passed through a Sephadex G-10 column, eluting with water. The first band was blue, and the second band was red. The desired red fractions were combined, and freezing-dried to give the pink powder product. The compound was recrystallized three times for spectroscopic measurements from water. *Anal.* Found: C, 32.93; H, 5.47; N, 12.88%. Calc. for $\text{C}_{12}\text{H}_{24}\text{ClCrN}_4\text{O}_8$: C, 32.77; H, 5.50; N, 12.74%.

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