

## The Inhibition of C-steel Corrosion in $H_3PO_4$ Solution by Some Furfural Hydrazone Derivatives

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## The Inhibition of C-steel Corrosion in $H_3PO_4$ Solution by Some Furfural Hydrazone Derivatives

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**요 약.** 3 M  $H_3SO_4$ (M: 몰농도) 용액에서 탄소강의 부식방지제로 Furfural hydrazone 유도체의 효과를 질량손실법 및 정전류극성법을 사용해 연구하였다. 이들 유도체 존재하에서 탄소강의 부식속도가 급격히 감소함을 관찰하였다. 이 연구로부터 부식방지효율은 부식방지제 농도가 증가함에 따라 증가하였고  $I^-$ 와  $SCN^-$ 을 첨가하면 부식방지효율은 더욱 증가되었다. 질량손실법을 사용해  $5 \times 10^{-6}$  M의 유도체가 있을 때와 없을 때 30-60 °C 사이에서 탄소강 부식에 미치는 온도 효과를 보았다. 부식과정에 대한 활성화에너지( $E_a^*$ )와 다른 열역학적 변수들을 계산하였고 이들에 대해 논의하였다. 정전류극성법을 통해 유도체들이 혼합형 방지제로 작용함을 알았고 외부전류를 흘려주었을 때 음극은 더욱 분극되었다. 3 M  $H_3SO_4$  용액에서 탄소강 표면에 이들 유도체들의 흡착은 Frumkin의 흡착등온을 따랐다. 이들 유도체들의 화학구조를 통해 부식방지 메커니즘을 설명하였다.

**주제어:**  $H_3PO_4$ , 부식, 탄소강, Furfural Hydrazone 유도체들

**ABSTRACT.** The effect of some furfural hydrazone derivatives (I -V) as corrosion inhibitors for C-steel in 3M  $H_3PO_4$  solution in which M indicates mol/l has been studied using weight-loss and galvanostatic polarization techniques. A significant decrease in the corrosion rate of C-steel was observed in the presence of the investigated compounds. This study revealed that, the inhibition efficiency increases with increasing the inhibitor concentration, and the addition of iodide and thiocyanate ions enhances it to a considerable extent. The effect of temperature on the corrosion behavior of C-steel was studied in the range from 30-60°C with and without  $5 \times 10^{-6}$  M investigated compounds using weight-loss method. Apparent activation corrosion energy ( $E_a^*$ ) and other thermodynamic parameters for the corrosion process were calculated and discussed. The polarization measurements show that the investigated compounds act as mixed-type inhibitors, but the cathode is more polarized when an external current was applied. The adsorption of these compounds on the surface of C-steel in 3 M  $H_3PO_4$  obeys Frumkin's adsorption isotherm. The mechanism of inhibition was discussed in the light of the chemical structure of the investigated compounds.

**Keywords:**  $H_3PO_4$ , Corrosion, C-steel, Furfural Hydrazone Derivatives

### INTRODUCTION

The corrosion inhibition of C-steel becomes interest because C-steel is widely applicable in industry

and domestic life and it is widely used as a constructional material in many industries due to its excellent mechanical properties and low cost.

The corrosion inhibition of iron and its alloys by

different organic inhibitors in acidic solutions has been studied by several authors.<sup>1-12</sup>

A little work appears to have been done on the corrosion inhibition of steel alloys in H<sub>3</sub>PO<sub>4</sub>. Some quaternary ammonium compounds were prepared from cycloaliphatic amines and have been tested as corrosion inhibitors for mild steel in 50% H<sub>3</sub>PO<sub>4</sub>.<sup>13</sup> 2-mercaptopyrimidine is effective for the corrosion inhibition of low C-steel over a wide concentration range of aqueous H<sub>3</sub>PO<sub>4</sub> solutions.<sup>14</sup> The effect of some aromatic hydrazone derivatives as corrosion inhibitors for C- steel in 1M H<sub>3</sub>PO<sub>4</sub><sup>15</sup> has been studied.

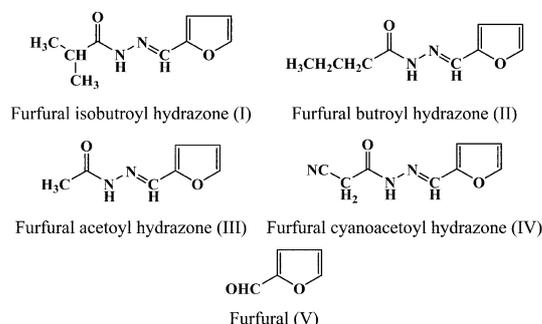
In the present paper the corrosion inhibition of C-steel in H<sub>3</sub>PO<sub>4</sub> solution by some furfural hydrazone derivatives using weight-loss and galvanostatic polarization technique was studied. The effect of addition of iodide and thiocyanate ions on the inhibition efficiency was tested. Moreover, the effect of temperature on the corrosion rate was studied to calculate the thermodynamic parameters related to the corrosion process.

## EXPERIMENTAL

### Materials

The experiments were performed with C-steel having the chemical compositions (weight %), C 0.2, Mn 0.35, P 0.24, Si 0.003 and Fe is the remainder.

The molecular structures of the investigated furfural hydrazone derivatives are shown below and were chosen because they are available, easily prepared and non toxic.



Hydrazone derivatives derived from furfural were prepared<sup>16</sup> by mixing equimolecular amounts of the

respective aldehyde and hydrazone in absolute ethanol. The mixture was refluxed on water bath for 2-6 hours and the solution was evaporated to about its half volume, left to cool until white, yellowish or orange crystals were separated out. These were filtered off, recrystallized from ethanol and finally dried in vacuum desiccator over anhydrous calcium chloride. Their structures were characterized by NMR, IR spectroscopy and element analyses before use.

100 ml stock solutions (10<sup>-3</sup> M) of compounds (I-V) were prepared by dissolving an accurately weighed quantity of each material in an appropriate volume of absolute ethanol, and then the required concentrations (1×10<sup>-6</sup> M to 11×10<sup>-6</sup> M) were prepared by dilution with double distilled water.

Approximately 5M H<sub>3</sub>PO<sub>4</sub> solution was prepared by diluting the appropriate volume of the concentrated chemically pure acid (BDH grade), with double distilled water. The concentration of the acid was checked by titrating an appropriately diluted portion with standard solution of sodium hydroxide, which was titrated in turn against standard solution of oxalic acid. From this stock concentrated solution, exactly 3M H<sub>3</sub>PO<sub>4</sub> solutions were prepared by dilution with double distilled water which was used throughout all experiments for the preparation of solutions.

Two different techniques have been employed for studying the corrosion inhibition of C-steel by hydrazone derivatives, these are:

- Chemical technique (Weight loss method).
- Electrochemical technique (Galvanostatic polarization method).

### Chemical Technique (Weight loss method)

The reaction basin used in this method was a graduated glass vessel 6 cm inner diameter having a total volume of 250 ml. 100 ml of the test solution were employed in each experiment. The test pieces were cut into 2×2×0.2 cm. They were mechanically polished with emery paper (a coarse paper was used initially and then progressively finer grades were employed), degreased in acetone,<sup>17</sup> rinsed with double distilled water and finally dried between two filter papers and weighed. The test pieces were sus-

pended by suitable glass hooks at the edge of the basin, and under the surface of the test solution by about 1 cm. After specified periods of time, three test pieces were taken out from the test solution, rinsed with double distilled water, dried as before and weighed again. The average weight loss at a certain time for each set of three samples was taken. The weight loss was recorded to the nearest 0.0001 gram.

#### Electrochemical Technique (Galvanostatic polarization method)

The electrochemical experiments were carried out in a three-electrode electrochemical cylindrical Pyrex glass cell with a platinum counter electrode and saturated calomel electrode (SCE) as reference. The working electrode had the form of a disc cut from C-steel sheet. The disc with an exposed area of 1 cm<sup>2</sup> was used and was weld from one side to a copper wire for electric connection and embedded in a glass tube of just larger diameter than the sample. Epoxy resin was used to stick the sample to the glass tube.<sup>18</sup> The exposed area was polished with different grades of emery papers in the normal way starting from coarser to finer, followed by degreasing in acetone and finally washing with double distilled water, just before insertion in the electrolytic cell.

A constant quantity of the test solution (100 ml) was taken in the polarization cell. A time interval of about 30 minutes was given for the system to attain a steady state and the open circuit potential (OCP) was noted. Both cathodic and anodic polarization curves were recorded galvanostatically using an Amel Galvanostat (Model-549) and digital Multimeters (Fluke-73) were used for accurate measurements of potential and current density. Solutions were not deaerated to make the conditions identical to weight loss measurements. All the experiments were carried out at 30±1 °C by using ultra circulating thermostat.

## RESULTS AND DISCUSSION

### Weight-Loss Measurements

The Corrosion Inhibition Behavior- Weight-loss

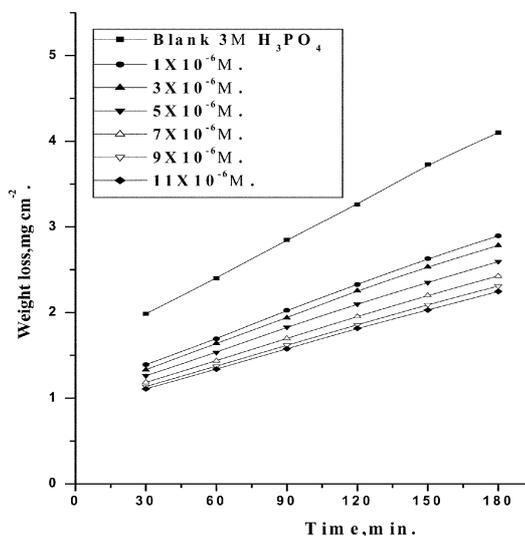


Fig. 1. Weight loss-time curves for C-steel dissolution in 3 M H<sub>3</sub>PO<sub>4</sub> in absence and presence of different concentrations of compound (1) at 30 °C.

of C-steel was determined at various time intervals in absence and presence of different concentrations of furfural hydrazone derivatives (I-V). The obtained weight loss-time curves are represented in Fig. 1 for inhibitor (I), the most effective one. Similar curves were obtained for other inhibitors (not shown).

As shown from this Figure, by increasing the concentration of compound (1), the weight loss of C-steel samples is decreased. This means that the presence of this compound retards the corrosion of C-steel in 3 M H<sub>3</sub>PO<sub>4</sub> or in other words, this compound acts as inhibitor.

The linear variation of weight loss with time in uninhibited and inhibited 3 M H<sub>3</sub>PO<sub>4</sub> indicates the absence of insoluble surface films during corrosion. In the absence of any surface films, the inhibitors are first adsorbed onto the metal surface and thereafter impede corrosion either by merely blocking the reaction sites (anodic and cathodic) or by altering the mechanism of the anodic and cathodic partial processes.

The percentage inhibition efficiencies (%I) of the investigated furfural hydrazone derivatives were determined using the equation (1):

Table 1. Inhibition Efficiency (%I) of C-Steel Dissolution After 120 Minutes Immersion in 3 M H<sub>3</sub>PO<sub>4</sub> in Presence of Different Concentrations of Furfural Hydrazone Compounds at 30 °C.

Concentration M	Inhibition efficiency (%I)				
	I	II	III	VI	V
1×10 <sup>-6</sup>	28.59	24.99	21.30	17.01	13.24
3×10 <sup>-6</sup>	30.82	27.29	23.29	18.58	14.89
5×10 <sup>-6</sup>	35.62	31.41	27.80	23.32	18.18
7×10 <sup>-6</sup>	40.11	36.07	32.06	27.78	22.36
9×10 <sup>-6</sup>	43.08	39.19	34.80	30.26	25.75
11×10 <sup>-6</sup>	44.28	40.37	36.41	31.39	27.06

$$\%I = [(W_1 - W_2) / W_1] \times 100 \quad (1)$$

where  $W_1$  and  $W_2$  are the corrosion rate of C-steel in the absence and presence the investigated furfural hydrazone derivatives, respectively at given time period and temperature. From the calculated values of the % inhibition efficiency (%I) given in Table 1 at 30 °C, the order of decreased inhibition efficiencies of the investigated furfural hydrazone derivatives is: I > II > III > IV > V.

#### Synergistic effect

The corrosion behavior of C-steel in 3M H<sub>3</sub>PO<sub>4</sub> solution in the presence of 1×10<sup>-4</sup> M potassium iodide and potassium thiocyanate at different concentrations of inhibitors (I-V) was studied. The % inhibition efficiencies were increased in the presence of these anions than in its absence. In case of KI the % inhibition efficiency in presence of I<sup>-</sup> and SCN<sup>-</sup> ions are presented in Table 2. From these results, it is observed that (%I) is enhanced due to synergistic effects.<sup>19</sup> The synergistic effect of these anions (I<sup>-</sup> and SCN<sup>-</sup>) has been observed.<sup>20</sup> The data in Table 2 indicate that this co-operative effect

decreased in order I<sup>-</sup> > SCN<sup>-</sup> suggesting a possible role by the I<sup>-</sup> and SCN<sup>-</sup> which also decrease in the order I<sup>-</sup> (0.135 nm) > SCN<sup>-</sup> (0.095 nm). The strong chemisorption of (I<sup>-</sup> and SCN<sup>-</sup>) anions on the metal surface are responsible for the synergistic effect of iodide and thiocyanate anions in combination with cation of the inhibitor. The cation is then adsorbed by columbic attraction on the metal surface where iodide and thiocyanate anions are already adsorbed by chemisorption. Stabilization of adsorbed iodide and thiocyanate anions with cations leads to a greater surface coverage and therefore greater inhibition. One can conclude that the addition of iodide and thiocyanate ions enhances the inhibition efficiency to a considerable extent, due to the increase of the surface coverage in the presence of iodide and thiocyanate ions. The order of investigated compounds remains unchanged, as before. From the previous results it is known that KI could be considered as one of the effective anions for synergistic action within the investigated inhibitors.

#### Effect of Temperature on Corrosion Process

The effect of temperature on corrosion of C-steel

Table 2. Inhibition Efficiency (%I) of C-Steel Dissolution After 120 Minutes Immersion in 3 M H<sub>3</sub>PO<sub>4</sub> in Presence of 1×10<sup>-4</sup> M KI & KSCN at Different Concentrations of Furfural Hydrazone Derivatives at 30 °C.

Concentration (M)	Inhibition efficiency (%I)				
	I	II	III	VI	V
1×10 <sup>-6</sup>	66.9(62.0)	62.3(55.5)	58.0(48.5)	46.3(32.8)	37.1(23.9)
3×10 <sup>-6</sup>	69.6(65.3)	65.6(59.5)	61.0(52.2)	48.1(35.0)	40.2(28.5)
5×10 <sup>-6</sup>	74.2(69.0)	69.3(63.4)	65.2(55.2)	51.8(38.7)	44.2(31.6)
7×10 <sup>-6</sup>	77.3(72.4)	72.1(67.5)	69.3(58.6)	54.0(42.0)	47.2(36.5)
9×10 <sup>-6</sup>	80.2(76.4)	75.2(71.5)	72.4(62.6)	57.4(45.4)	50.6(40.2)
11×10 <sup>-6</sup>	82.6(79.5)	78.2(74.5)	74.3(66.3)	59.8(47.9)	53.7(43.3)

%I values in presence of 1×10<sup>-4</sup> M KSCN are given in parenthesis.

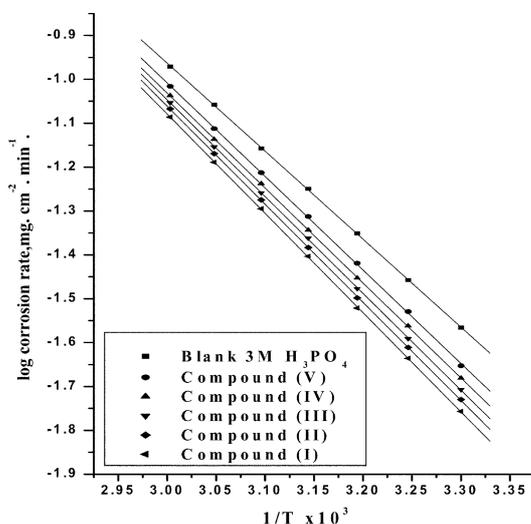


Fig. 2.  $\log k - 1/T$  curves for C-steel dissolution in 3 M  $H_3PO_4$  in absence and presence of  $5 \times 10^{-6}$  M furfural hydrazone derivatives.

in 3 M  $H_3PO_4$  solution in absence and presence of different concentrations of inhibitors (I-V) at different temperatures ranging from 30 °C to 60 °C was investigated.

Arrhenius plot, equation (2), of logarithm of the corrosion rate ( $\log W$ ) with the reciprocal of absolute temperature ( $1/T$ ) in absence and presence of  $5 \times 10^{-6}$  M inhibitors (I-V) was shown graphically in Fig. 2.

$$W = k \exp (E_a^*/RT) \quad (2)$$

where  $k$  is the Arrhenius preexponential constant, and  $R$  is the universal gas constant.

From the slopes of the plots, the respective apparent activation corrosion energy ( $E_a^*$ ) was calculated and tabulated in Table 3. The results show that, the

values of apparent activation corrosion energy ( $E_a^*$ ) increases in the same order of increasing inhibition efficiency of the inhibitors. It is also indicated that the whole process is controlled by surface reaction, since the energy of activation of the corrosion process is over 20  $\text{kJ mol}^{-1}$ .<sup>20</sup>

From the results of the effect of temperature, it was observed that the inhibition efficiency increases with increasing the inhibitor concentration at one hand, and decreases with the rise of temperature on the other hand. This means that these derivatives are adsorbed physically on C-steel surface.

From transition state theory<sup>21</sup>

$$W_{\text{corr}} = RT/N h \exp (\Delta S^*/R) \exp (-\Delta H^*/RT) \quad (3)$$

Where  $h$  is Planck's constant,  $N$  is Avogadro's number, we can calculate the  $\Delta H^*$  &  $\Delta S^*$  from plotting  $\log (k/T)$  vs. ( $1/T$ ), equation (3). Straight lines were obtained as shown in Fig. 3 for C-steel dissolution in 3 M  $H_3PO_4$  in absence and presence of  $5 \times 10^{-6}$  M of inhibitors (I-V). The order of inhibition efficiency of the investigated derivatives as gathered from the increase in  $E_a^*$  and  $\Delta H^*$  and the decrease in  $\Delta S^*$ , values remains unchanged and follows the order : I > II > III > IV > V.

The enthalpy change of activation ( $\Delta H^*$ ) ranged from 35 to 38  $\text{kJ mol}^{-1}$ , and the entropy change of activation ( $\Delta S^*$ ) ranged from -157 to -144  $\text{J mol}^{-1}\text{K}^{-1}$ , for corrosion of C-steel in 3 M  $H_3PO_4$  solution in absence and presence of  $5 \times 10^{-6}$  M of each additive. From the results obtained, it is clear that the presence of investigated derivatives increased the activation energy values and consequently decreased the corrosion rate of the C-steel. Moreover, activa-

Table 3. Activation Parameters of the Dissolution of C-Steel in 3 M  $H_3PO_4$  in the Absence and Presence of  $5 \times 10^{-6}$  M of Furfural Hydrazone Derivatives.

Inhibitor	Activation parameters		
	$E_a^*$ , $\text{kJ mol}^{-1}$	$\Delta H^*$ , $\text{kJ mol}^{-1}$	$-\Delta S^*$ , $\text{J mol}^{-1}\text{K}^{-1}$
Blank ( Free acid )	38.3	35.7	157.0
(I)	43.3	40.7	144.4
(II)	42.7	40.1	145.8
(III)	42.2	39.6	147.2
(IV)	41.4	38.7	149.3
(V)	40.7	38.2	156.6

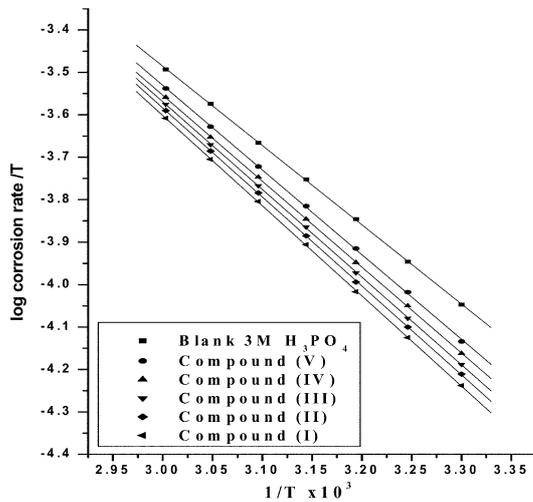


Fig. 3. log k/T-1/T curves for C-steel dissolution in 3 M H<sub>3</sub>PO<sub>4</sub> in absence and presence of 5×10<sup>-6</sup> M furfural hydrazone derivatives.

tion energy increased by increasing the concentration of the inhibitors. These results indicate that these investigated derivatives are acting as inhibitors by increasing the activation energy of C-steel dissolution and by making a barrier to mass and charge transfer due to their adsorption on C-steel surface. The values of ΔS<sup>\*</sup> in absence and presence of the investigated derivatives are negative; this indicates that the activated complex is the rate determining step representing an association rather than dissociation step.<sup>22</sup> This means that the activated molecules are in higher ordered state than that at the initial state.

**Adsorption Isotherms**

The variation of surface coverage determined by weight loss measurements, θ, with the logarithm of the inhibitor concentration, log C, is represented in Fig. 4. These curves are of S- shape obeying the Frumkin’s adsorption isotherm<sup>23</sup> and are in good agreement with the Frumkin’s equation (4). As shown from this Figure, one can conclude that the degree of surface coverage increases as the concentration of the inhibitor increases and hence, the inhibition efficiency increases.

Frumkin’s equation is:

$$[\theta/(1-\theta)] \exp (-2a\theta)=K C \quad (4)$$

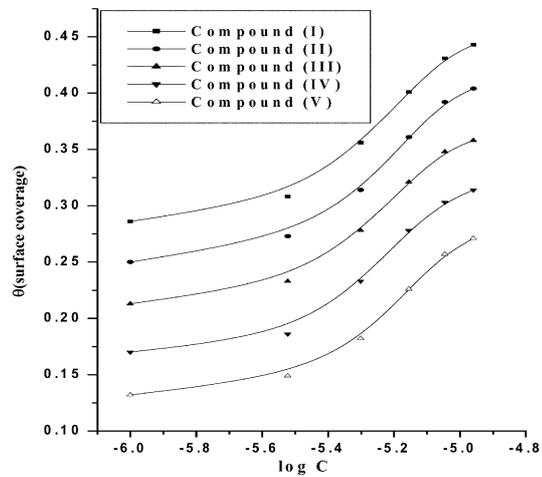


Fig. 4. θ-log C curves for C-steel dissolution in 3 M H<sub>3</sub>PO<sub>4</sub> in presence of furfural hydrazone derivatives.

where (a) is a molecular interaction parameter depending on the molecular interaction in the adsorption layer and on the degree of heterogeneity of the surface, C is the molar concentration of inhibitor, K is the equilibrium constant of the adsorption reaction.

**Galvanostatic Polarization Technique**

Fig. 5 shows the galvanostatic polarization curves

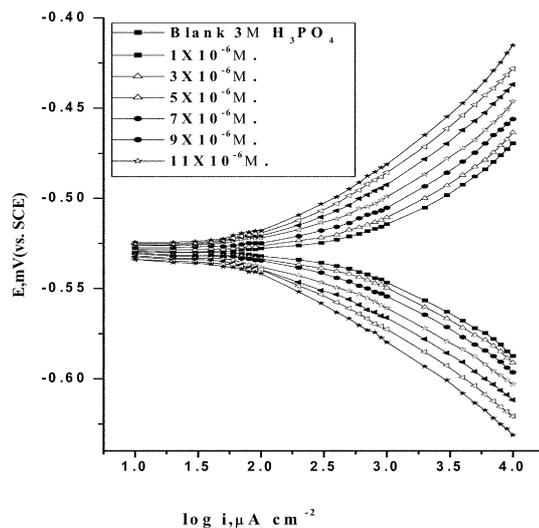


Fig 5. Galvanostatic polarization curves for C- steel in 3 M H<sub>3</sub>PO<sub>4</sub> in absence and presence of different concentrations of inhibitor (I) at 30 °C.

Table 4. Effect of Inhibitor (I) Concentrations on the Corrosion Potential ( $E_{\text{corr}}$ ), Corrosion Current Density ( $i_{\text{corr}}$ ), Tafel Slopes ( $\beta_c$  and  $\beta_a$ ), Degree of Surface Coverage ( $\theta$ ) and Percentage Inhibition (%I) on C- Steel in 3 M  $\text{H}_3\text{PO}_4$  at 30 °C.

Concentration, M	$-E_{\text{corr}}$ (mV)	$i_{\text{corr}}$ $\mu\text{A cm}^{-2}$	$\beta_c$ $\text{mV dec.}^{-1}$	$\beta_a$ $\text{mV dec.}^{-1}$	$\theta$	% I
Blank	531	384	152	157	-	-
$1 \times 10^{-6}$	533	279	159	169	0.266	26.56
$3 \times 10^{-6}$	533	272	163	172	0.292	29.20
$5 \times 10^{-6}$	538	267	166	175	0.349	34.90
$7 \times 10^{-6}$	540	260	170	181	0.388	38.80
$9 \times 10^{-6}$	541	255	173	186	0.412	41.20
$11 \times 10^{-6}$	544	249	185	188	0.431	43.10

for C-steel dissolution in 3 M  $\text{H}_3\text{PO}_4$  in absence and presence of different concentrations of inhibitor (I) at 30 °C. Similar curves were obtained for other inhibitors (not shown).

The numerical values of the variation of corrosion current density ( $i_{\text{corr}}$ ), corrosion potential ( $E_{\text{corr}}$ ), Tafel slopes ( $\beta_a$  and  $\beta_c$ ), and percentage of inhibition efficiency (%I) degree of surface coverage ( $\theta$ ) with the concentrations of different furfural hydrazone derivatives are given in Table 4. In the case of polarization method, the relation determines the inhibition efficiency (%I) is:

$$\%I = [1 - (i_{\text{corr}}/i_{\text{corr}}^o)] \times 100 \quad (5)$$

Where  $i_{\text{corr}}$  and  $i_{\text{corr}}^o$  are the inhibited and uninhibited corrosion current densities, respectively, determined by extrapolation of cathodic and anodic Tafel lines.

The results indicate that:

1. The cathodic and anodic curves obtained exhibit Tafel-type behavior. Addition of furfural hydrazone derivatives increased both the cathodic and anodic overpotentials and the presence of furfural hydrazone derivatives in solution inhibits both the hydrogen evolution and the anodic dissolution processes i.e. they are mixed-type inhibitors.
2. The corrosion current density ( $i_{\text{corr}}$ ) decreases with increasing the concentration of furfural hydrazone derivatives which indicates that these derivatives act as inhibitors, and the degree of inhibition efficiency depends on the concentration and type of inhibitors present.
3. The slopes of the anodic and the cathodic Tafel

lines ( $\beta_a$  and  $\beta_c$ ) were slightly changed on increasing the concentration of the tested derivatives. This indicates that there is no change of the mechanism of corrosion in presence and absence of inhibitors.

4. The order of inhibition efficiency of all inhibitors at different concentrations, as given by polarization measurements are listed in Table 4, and is in the order of:

$$I > II > III > IV > V.$$

5. The results obtained from weight-loss measurements are in good agreement with that obtained in polarization measurements.

#### Chemical Structure of the Inhibitors and Corrosion Inhibition

Inhibition of corrosion of C-steel in 3M  $\text{H}_3\text{PO}_4$  by the investigated furfural hydrazone derivatives as determined from weight loss and polarization measurements was found to depend on the concentration and the nature of the inhibitor.

The observed corrosion data in the presence of the selected furfural hydrazone derivatives showed in the following:

- i) The corrosion rate and corrosion current density decrease with increasing the concentration of the inhibitors.
- ii) The linear variation of weight loss with time.
- iii) The parallel shift in Tafel lines to higher potential regions.
- iv) The decrease in % inhibition efficiency (%I) with increasing the temperature.

This indicates that the corrosion inhibition occurs by adsorption of the inhibitor at the electrode / solution interface. The nature of inhibitor interaction on

metal surface during corrosion inhibition has been known from its adsorption characteristics. However, inhibition efficiency (%I) of the investigated derivatives depends on many factors which include the number of adsorption sites in the molecule and their charge density, molecular size, stability of these derivatives in acidic solution, mode of adsorption and formation of metallic complexes.

It was found that, the order of decreasing inhibition efficiency (%I) of the tested compounds in 3 M H<sub>3</sub>PO<sub>4</sub> for C-steel is:

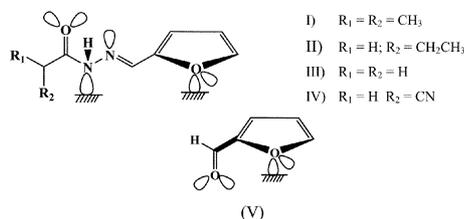


Mode of adsorption of furfural hydrazone derivatives on the surface of C-steel is shown below and clearly indicates the adsorption centers. As shown from this presentation, the compound (V) is the least efficient inhibitor in this series because of the presence of only one active adsorption center (one O-atom), but the other used investigated compounds have two active adsorption centers (one N-atom and one O-atom).

Variation in structure of inhibitors molecules (I-IV) takes place through the substituent attached to the carbonyl group. So, the inhibition efficiency will depend on this part of the molecule. This order of decreased inhibition efficiency (%I) of the additives can be accounted for in terms of increase the length of the chain<sup>24</sup> within the given homologous series (I-IV). Compounds (I-IV) have two adsorption active centers. Compound I has two branched -CH<sub>3</sub> groups which increase the electron density on the molecule, compound II has one -CH<sub>3</sub> group therefore it comes after compound I in %I. Compound III comes after compound II because it has no -CH<sub>3</sub> group. Compound IV comes after compound III because it has -CN group which is a reducing group (decrease the electron density on the molecule). Compound V has one active adsorption center, so it is the least effective one.

The order of inhibition efficiency (%I) of furfural hydrazone derivatives derived from weight loss method is further supported by galvanostatic measurement. The observed agreement among these independent techniques prove the validity of the results obtained and supports the explanation

given for the effect of chemical structure on the inhibition action of these investigated compounds.



In our Lab we use some aromatic furfural hydrazone derivatives<sup>15</sup> as corrosion inhibitors for C-steel in 1M H<sub>3</sub>PO<sub>4</sub> at the same concentrations, the % inhibition was found to be 10-46% which is less than % inhibition in our case. This may be due to, the aromatic moiety does not lie flat on the C-steel surface so, less surface coverage obtained and hence less % inhibition.

## CONCLUSIONS

1. Furfural hydrazone derivatives are fairly efficient inhibitors for C-steel dissolution in 3M H<sub>3</sub>PO<sub>4</sub>.
2. Percentage inhibition efficiency (% I) are increased in the presence of 1×10<sup>-4</sup> M KI, and KSCN due to the synergistic effect.
3. Efficiency of these investigated compounds decreased with the rise of temperature.
4. These investigated compounds show an adsorption on C-steel surface according to Frumkin isotherm.
5. Polarization measurements show that these investigated furfural hydrazone derivatives act as a mixed inhibitors without modifying the mechanism of hydrogen evolution.

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