



Ocean Acidification

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Summary

With increasing concentrations of carbon dioxide (CO₂) in the atmosphere, the extent of effects on the ocean and marine resources is an increasing concern. One aspect of this issue is the ongoing process whereby seawater becomes acidified (i.e., ocean acidification) as more CO₂ dissolves in it, causing hydrogen ion concentration in seawater to increase. Scientists are concerned that increasing hydrogen ion concentration could reduce growth or even cause death of shell-forming animals (e.g., corals, molluscs, and certain planktonic organisms) as well as disrupt marine food webs and the reproductive physiology of certain species. While not yet fully understood, the ecological and economic consequences of ocean acidification could be substantial.

Scientists are concerned that increasing acidification could alter biogeochemical cycles, disrupt physiological processes of marine organisms, and detrimentally alter marine ecosystems. This report does not discuss the effects of increasing thermal stress to marine organisms and ecosystems (e.g., coral bleaching) related to climate change. However, marine ecosystems are likely to be affected by the synergistic effects of factors involved in both thermal and acidification processes.

Congress is beginning to focus attention on better understanding ocean acidification and determining how this concern might be addressed. In the 111th Congress, the Federal Ocean Acidification Research And Monitoring Act of 2009 (Title XII, Subtitle D, of P.L. 111-11) directed the Secretary of Commerce to establish an ocean acidification program within NOAA, established an interagency committee to develop an ocean acidification research and monitoring plan, and authorized appropriations through FY2012 for NOAA and the National Science Foundation.

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Introduction

On January 30, 2009, a Monaco Declaration, signed by more than 150 marine scientists from 26 countries, called for immediate action by policymakers to reduce carbon dioxide emissions so as to avoid widespread and severe damage to marine ecosystems from ocean acidification.¹ The Monaco Declaration is based on the *Research Priorities Report* developed by participants in an October 2008 second international symposium on “The Ocean in a High-CO₂ World,”² organized by UNESCO’s Intergovernmental Oceanographic Commission, the Scientific Committee on Oceanic Research, the International Atomic Energy Agency, and the International Geosphere Biosphere Programme. While not yet fully understood, the ecological and economic consequences of ocean acidification could be substantial.

What Is Ocean Acidification?

The complex interplay between rising carbon dioxide (CO₂) levels in the atmosphere and the ocean’s chemistry is a process that scientists have recognized for decades. As increasing CO₂ from the atmosphere dissolves in seawater, seawater chemistry is altered. The prevailing pH of water near the ocean surface is around 8.1, or slightly alkaline.³ Ocean acidification is the name given to the ongoing process whereby pH decreases as seawater becomes acidified when increasing amounts of anthropogenic CO₂ from the atmosphere dissolve in seawater. When atmospheric CO₂ dissolves into the ocean, it forms carbonic acid (H₂CO₃). Some of the carbonic acid dissociates in ocean waters, producing hydrogen ions (H⁺). As the number of hydrogen ions increases, the pH of the ocean decreases, and the water becomes more acidified.

Scientists are concerned that increasing acidification could alter biogeochemical cycles, disrupt physiological processes of marine organisms, and detrimentally alter marine ecosystems. This report does not discuss the effects of increasing thermal stress to marine organisms and ecosystems (e.g., coral bleaching) related to climate change. However, marine ecosystems are likely to be affected by the synergistic effects of factors involved in both thermal and acidification processes.

At What Rate Is Ocean Acidification Occurring and What Factors Affect This Rate?

Over the past several decades, of the approximately 7 billion metric tons of carbon that all the countries in the world release as CO₂ into the atmosphere each year, the oceans take up about 2 billion tons.⁴ Between pre-industrial times and 1994, average pH of water near the ocean surface

¹ A copy of this declaration is available at <http://scrippsnews.ucsd.edu/Releases/doc/MonacoDeclaration.pdf>.

² The full report is available at http://ioc3.unesco.org/oanet/Symposium2008/ResearchPrioritiesReport_OceanHighCO2WorldII.pdf.

³ The pH scale is an inverse logarithmic representation of hydrogen proton (H⁺) concentration, indicating the activity of hydrogen ions (or their equivalent) in the solution. A pH of less than 7.0 is considered acidic, while a pH greater than 7.0 is considered basic (alkaline); a pH level of 7.0 is defined as “neutral.”

⁴ Richard A. Feely, Pacific Marine Environmental Laboratory, National Oceanic and Atmospheric Administration, U.S. (continued...)

is estimated to have decreased (i.e., acidity increased) by almost 0.1 pH unit.⁵ That increase sounds small, but it represents an increase of 26% in the concentration of hydrogen ions, because the pH scale is logarithmic (i.e., water with a pH of 7 is 10 times more acidified than water with a pH of 8, and 100 times more acidified than water with a pH of 9). Open ocean observational records of declining pH are available from the Hawaiian Ocean Time-series Station in the Pacific and the Bermuda Atlantic Time-series Station in the Atlantic.⁶ Up to a point, as atmospheric CO₂ continues to increase, the oceans will continue to become more acidified; one estimate suggests that the rate of CO₂ uptake by the oceans could stabilize at around 5 gigatons per year by 2100.⁷

All gases, such as CO₂, are less soluble with increasing water temperature. Thus, marine waters near the poles have a much greater capacity for dissolving CO₂ than do ocean waters in the tropics. In addition, dissolved CO₂ also is transported into ocean depths at these high latitudes (i.e., deep water formation mechanism) since the lower-temperature waters are of higher density, causing greater convection to occur than happens in the more stratified tropical oceans. If temperature were the only factor affecting the rate of ocean acidification and the appearance of impacts on physical and biological features, these impacts might be more likely to occur in marine waters nearer the poles. However, in addition to temperature, other factors modulate the impact of CO₂ on marine waters. Respiration adds CO₂ to seawater, and photosynthesis removes it. Deep oceanic water is enriched in CO₂ due to respiration in the absence of photosynthesis and, when brought to the surface by equatorial currents (i.e., upwelling), can place CO₂-enriched seawater in contact with the atmosphere where it can absorb even more CO₂. Hence, the tropics are also vulnerable to near-term effects, most notably tropical reefs. An additional factor is the potential increase in storm activity at higher latitudes as some climate models suggest.⁸ CO₂ and other acidic gasses such as nitrogen dioxide are also dissolved in rainwater.⁹ An increase in North Atlantic or western North Pacific storms could have significant implications for accelerating acidification of ocean waters near the surface in those regions.

Key scientific questions concern which factors may affect the future rate of acidification, especially whether the rate of increase will remain constant in direct relationship to the amount of CO₂ in the atmosphere or whether other factors will result in an acceleration or deceleration of this acidification rate. There is also the question of equilibrium—that is, how long it might take the acidification process of ocean waters, should the currently increasing atmospheric emission rate of CO₂ eventually be able to be stabilized or diminished, to come into equilibrium with the concentration of atmospheric CO₂. An adjunct to this question is how long might it take the rate of ocean acidification to slow (or even begin to decrease) in response to any measures that might be taken to slow, halt, or even reverse the increasing concentration of CO₂ in the atmosphere.

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Department of Commerce, World Ocean Forum, Nov. 13-14, 2006, at http://www.thew2o.net/events/oceans/oa_q_and_a.php and Richard A. Feely et al., "Impact of Anthropogenic CO₂ on the CaCO₃ System in the Oceans," *Science* (2004), vol. 305, pp. 362-366.

⁵ James C. Orr et al., "Anthropogenic Ocean Acidification over the Twenty-First Century and Its Impact on Calcifying Organisms," *Nature*, vol. 437 (2005): 681-686.

⁶ See http://hahana.soest.hawaii.edu/hot/hot_jgofs.html and <http://www.bios.edu/research/bats.html>.

⁷ David Archer, "Fate of Fossil Fuel CO₂ in Geologic Time," *Journal of Geophysical Research*, v. 110 (2005): C09S05, doi:10.1029/2004JC002625.

⁸ See the Intergovernmental Panel on Climate Change's *Technical Paper on Climate Change and Water*, available at http://www.ipcc.ch/meetings/session28/executive_summary.pdf.

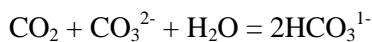
⁹ Rainwater is naturally acidic at a pH of around 5.6, and downwind of pollution sources has been measured as low as pH 3.0.

Additional questions relate to how ocean circulation, which eventually controls CO₂ uptake rate, might change in response to rising temperatures caused by greenhouse gas emissions.

What Are Some of the Potential Effects of Ocean Acidification?

Since the marine environment is complex and some of the likely changes may be years in the future, the potential effects identified in this section, although many are supported by laboratory experimentation, are primarily conjecture and/or forecasts. However, field studies are beginning to provide a more direct view of potential ocean acidification problems.¹⁰

A lower pH affects marine life in the oceans and is related to other changes in ocean chemistry. In addition to the lower pH, another consequence of the increased amount of dissolved CO₂ in the ocean is the production of more bicarbonate ions (HCO₃¹⁻). As more CO₂ dissolves into the ocean, bicarbonate ions form at the expense of carbonate ions (CO₃²⁻), which scientists often describe by the following reaction:



The abundance and availability of carbonate ions are critical to many shell-forming marine organisms. At current average ocean pH levels (about 8 or above), ocean waters near the surface have ample carbonate ions to support shell formation and coral growth. However, as increased amounts of carbonic acid form in the ocean from higher CO₂ levels in the atmosphere, pH gets lower and the amount of carbonate ions in the oceans decreases, resulting in fewer carbonate ions available for making shells.

Organisms make biogenic calcium carbonate for their shells by combining calcium ions (Ca²⁺)—which are abundant in the oceans—with carbonate ions to form solid calcium carbonate (CaCO₃). Marine organisms such as corals and pteropods precipitate one mineral type of calcium carbonate called aragonite, and other organisms such as foraminifera and coccolithophorids use another type called calcite. A third type of calcium carbonate—high magnesium calcite—is precipitated by echinoderms (sea urchins, starfish, and sea cucumbers) and some coralline algae. Under present conditions of ocean chemistry, these forms of calcium carbonate are relatively stable in waters near the ocean surface, except for certain areas of high upwelling activity.¹¹ Water near the ocean surface currently is supersaturated with respect to calcite, high magnesium calcite, and aragonite,¹² meaning that organisms easily can form shells from all of these mineral types. However, as more carbonic acid is formed in water near the ocean surface from higher levels of CO₂ in the atmosphere, the level of saturation decreases. If the ocean waters become

¹⁰ See, for example, K. K. Yates and R. B. Halley, “Diurnal Variations in Rates of Calcification and Carbonate Sediment Dissolution in Florida Bay,” *Estuaries and Coasts*, v. 29 (2006): 24-39; and K. K. Yates and R. B. Halley, “CO₃²⁻ Concentration and pCO₂ Thresholds for Calcification and Dissolution on the Molokai Reef Flat, Hawaii,” *Biogeosciences*, v. 3 (2006): 357-369.

¹¹ Results from a 2007 National Oceanic and Atmospheric Administration survey along the U.S. West Coast documented the first undersaturated waters at the ocean surface along the California coast, brought to the surface by seasonal upwelling. See Richard A. Feely et al., “Evidence for Upwelling of Corrosive “Acidified” Water onto the Continental Shelf,” *Science*, v. 320, no. 5882 (June 13, 2008): 1490-1492.

¹² Surface ocean waters in the tropics are currently supersaturated by a factor of more than 3.

undersaturated, then shells made from all of these minerals would tend to dissolve. Shells made from high magnesium calcite would tend to dissolve first, at lower concentrations of carbonic acid (and thus at higher pH values) than would shells made from aragonite. Shells made from calcite would dissolve at higher concentrations of carbonic acid than those made from aragonite.¹³ Thus, organisms that incorporate high magnesium calcite (i.e., echinoderms and some coralline algae) are likely to be the “first responders” to ocean acidification.¹⁴ Ocean waters at depths of thousands of feet are undersaturated with respect to all forms of biogenic calcite, which is why most of the shells from dead organisms that “rain” down from the ocean surface dissolve before reaching the ocean floor.¹⁵ Some suggest that marine surface waters closer to the poles may become undersaturated within the next 50 years.¹⁶ Researchers at the Antarctic Climate and Ecosystems Cooperative Research Centre have demonstrated significant reductions in shell mass and thickness of several Southern Ocean marine plants and animals that appear consistent with the projected effects of recent increased acidification of the ocean.¹⁷

In response to ocean acidification, scientists have projected that mussel and oyster calcification, and thus shell strength, could decrease significantly by the end of the 21st century, according to the Intergovernmental Panel on Climate Change’s IS92a scenario.¹⁸ There is also the concern that increased acidification may cause marine calcium carbonate sediments to dissolve with potential detrimental effects on species and communities residing in and on these sediments.¹⁹ Since many of these organisms provide food or modify habitat for other organisms, the well-being of these carbonate-dependent species will affect other species. Because of these interrelationships, the potential indirect effects of acidification on other marine organisms is not well understood. While some have raised concerns that ocean acidification, by negatively affecting calcifying plankton species, could shift ecological balances so as to increase populations of some noncalcifying species, there appears to be no significant relationship between jellyfish abundance and acidified conditions, and any role of pH in structuring zooplankton communities is believed to be tenuous.²⁰

There are also concerns that increasing acidification of ocean waters could alter the ability of some organisms to conduct essential biochemical and physiological processes.²¹ For example,

¹³ See Railsback, *Some Fundamentals of Mineralogy and Geochemistry*, at <http://www.gly.uga.edu/railsback/Fundamentals/820HMC-LMCSolubilities05LS.pdf>.

¹⁴ Andreas J. Andersson, Fred T. Mackenzie, and Nicholas R. Bates, “Life on the Margin: Implications of Ocean Acidification on Mg-Calcite, High Latitude and Cold-Water Marine Calcifiers,” *Marine Ecology Progress Series*, v. 373 (2008): 265-273.

¹⁵ Recent research suggests that no more than about 30% of the calcium carbonate produced in the surface ocean each year is buried in shallow or deep sea sediments; the rest dissolves on its way down the water column. See Feely et al. (2004), p. 365.

¹⁶ Orr et al. (2005).

¹⁷ Bruce Mapstone, “Acid Oceans in the Spotlight,” *Antarctic Climate and Ecosystem News*, edition 4 (August 2008): 1.

¹⁸ Frederic Gazeau et al., “Impact of Elevated CO₂ on Shellfish Calcification,” *Geophysical Research Letters*, v. 34, no. 7 (Apr. 16, 2007): L07603 (5 p.).

¹⁹ M. Gehlen, L. Bopp, and O. Aumont, “Short-term Dissolution Response of Pelagic Carbonate Sediments to the Invasion of Anthropogenic CO₂: A Model Study,” *Geochem. Geophys. Geosyst.*, v. 9 (Feb. 16, 2008): Q02012.

²⁰ A. J. Richardson and M. J. Gibbons, “Are Jellyfish Increasing in Response to Ocean Acidification?,” *Limnology and Oceanography*, v. 53, no. 5 (2008):2040-2045.

²¹ Hans-O. Pörtner, “Ecosystem Effects of Ocean Acidification in Times of Ocean Warming: A Physiologist’s View,” *Marine Ecology Progress Series*, v. 373 (2008): 203-217.

scientists have found that, when exposed to water of pH 7.7, roughly equivalent to acidity levels predicted for the year 2100, sea urchin sperm swam much more slowly. Overall, fertilization fell by 25%, and in almost 26% of cases where the eggs were fertilized, they did not survive long enough to develop into larvae.²² While marine invertebrates in general, and their early developmental stages in particular, are believed to be more sensitive to environmental disturbance, available data to assess their vulnerability to ocean acidification is contradictory.²³ Although evidence suggests that larval and juvenile fish are more susceptible to changes in ocean water pH than adults, larval and juvenile fish exposed to exceedingly high CO₂ concentrations (more than 100 times current levels) suffered little apparent harm.²⁴ Fish appear to be among the more tolerant marine animals.²⁵ These scientists believe that “the relative tolerance of fish may relate to high capacity for internal ion and acid-base regulation via direct proton excretion, and an intracellular respiratory protein that results in a high oxygen-carrying capacity and substantial venous oxygen reserve.”

The fossil record indicates that marine organisms may be quite sensitive to ocean acidification—about 55 million years ago, a large injection of CO₂ into the deep ocean, presumably resulting from a massive methane release, was followed by the extinction of some species of benthic foraminifera.²⁶ Others caution that these paleo-events may be imperfect analogs to current conditions.²⁷

Some have raised questions downplaying the potential harm to coral reefs from ocean acidification. Differences of opinion exist on the relative effects of climate change as expressed in increased CO₂ when compared to increased ocean temperature. Opinion has been expressed that, in marine systems, increased temperature may have detrimental effects comparable to or larger than those seen from increased CO₂ concentration, for corals and for phytoplankton.²⁸ Although calcification rates in massive *Porites* coral were reported to have declined over a 16-year study period by approximately 21% in two regions on Australia’s Great Barrier Reef, these findings were consistent with studies of the synergistic effect of elevated seawater temperatures and CO₂ concentrations on coral calcification.²⁹ While ocean acidification may not appear currently to be killing corals, such acidification is slowing development of coral larvae into juvenile colonies.³⁰

²² Jon N. Havenhand, Fenina-Raphaëla Buttler, Michael C. Thorndyke, and Jane E. Williamson, “Near-Future Levels of Ocean Acidification Reduce Fertilization Success in a Sea Urchin,” *Current Biology*, v. 18, no. 15 (August 2008): 651-652.

²³ S. Dupont and M. C. Thorndyke, “Impact of CO₂-Driven Ocean Acidification on Invertebrates Early Life History—What We Know, What We Need to Know and What We Can Do,” *Biogeosciences Discussions*, v. 6 (2009): 3109-3131.

²⁴ A. Ishimatsu et al., “Effects of CO₂ on Marine Fish: Larvae and Adults,” *Journal of Oceanography*, v. 60, no. 4 (2004): 731-741.

²⁵ V.J. Fabry et al., “Impacts of Ocean Acidification on Marine Fauna and Ecosystem Processes,” *ICES Journal of Marine Science*, v. 65 (2008): 414-432.

²⁶ J. C. Zachos et al., “Rapid Acidification of the Ocean During the Paleocene-Eocene Thermal Maximum,” *Science*, v. 308 (2005): 1611-1615.

²⁷ Scott C. Doney et al., “Ocean Acidification: The Other CO₂ Problem,” *Annual Review of Marine Science*, v. 1 (January 2009): 169-192.

²⁸ Clinton E. Hare et al., “Consequences of Increased Temperature and CO₂ for Phytoplankton Community Structure in the Bering Sea,” *Marine Ecology Progress Series*, v. 352 (2007), p. 14.

²⁹ T. F. Cooper et al., “Declining Coral Calcification in Massive *Porites* in Two Nearshore Regions of the Northern Great Barrier Reef,” *Global Change Biology*, v. 14 (2008): 529-538.

³⁰ R. Albright et al., “Effect of Aragonite Saturation State on Settlement and Post-Settlement Growth of *Porites* (continued...) ”

In support of the ability of certain corals to survive decreasing pH, stony and soft corals have been grown successfully in open systems with water from a saltwater well at a pH between 7.5 and 7.8 since the 1970s.³¹ However, given the high level of adaptation in corals to facilitate calcification via complex processes, at least some corals may be sensitive to changes in pH because of adaptation to invariant pH, with evidence coming primarily from the discovery that periods of high CO₂ in the geological past were often also periods of low aragonite-coral abundances and diversity.³² Others have found that certain species of coral survive in the laboratory at a pH 7.3 to 7.6 after their calcified structure dissolves by functioning similar to sea anemones, and retaining the ability to recalcify when pH is increased.³³ However, in the natural marine environment, predation could be a significant factor in limiting the viability of such “naked” corals, and it is unlikely that such organisms could form reefs and attract the diverse community that constitutes a coral reef.

Concern has also arisen that lower ocean water pH will diminish low-frequency (below 10 KHz) sound absorption in the ocean, increasing noise levels within the auditory range critical for environmental, military, and economic interests.³⁴ Frequency-dependent decreases to sound absorption related to the current decrease in pH of about 0.1 pH unit may exceed 12%, and an anticipated pH decrease of 0.3 pH units by mid-century may result in an almost 40% decrease in sound absorption. It is unknown how marine mammals might be affected by and adapt to an ocean increasingly transparent to sound at low frequencies.

What Are the Natural and Human Responses That Might Limit or Reduce Ocean Acidification?

Several natural feedback mechanisms can act to moderate the process of acidification. The more acidified the ocean becomes, the less CO₂ will be taken up by dissolution. In addition, the warmer the seawater becomes, the less CO₂ will dissolve. Speculative questions exist related to what might occur should the oceans reach an equilibrium in their ability to take up CO₂ and atmospheric CO₂ levels continue to increase. Even with increasing concentrations of atmospheric CO₂, the oceans are not likely to reach pH values of less than 7 (neutral).

Our ability to reduce ocean acidification through artificial means is unproven. Proposals have suggested the addition of chemicals to the ocean, such as (1) using iron compounds to stimulate planktonic algae growth whereby the increased photosynthesis might capture/remove dissolved CO₂, or (2) using limestone to neutralize (i.e., buffer) the more acidified streams and rivers near where they enter oceans and close to sources of limestone or add limestone powder directly to the

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astreoides Larvae,” *Coral Reefs*, v. 27, no. 3 (2008): 485-490.

³¹ M. J. Atkinson, B. Carlson, and G. L. Crow, “Coral growth in high nutrient, low-pH seawater: a case study of corals cultured at the Waikiki Aquarium, Honolulu, Hawaii,” *Coral Reefs*, v. 14, no. 4 (1995): 215-223.

³² Personal communication, John W. McManus, Director, National Center for Coral Reef Research, Rosenstiel School of Marine and Atmospheric Science, University of Miami, February 21, 2009.

³³ Maoz Fine and Dan Tchernov, “Scleractinian Coral Species Survive and Recover from Decalcification,” *Science*, v. 315, no. 5820 (March 30, 2007): 1811.

³⁴ Keith C. Hester et al., “Unanticipated Consequences of Ocean Acidification: A Noisier Ocean at Lower pH,” *Geophysical Research Letters*, v. 35 (2008): L19601, doi:10.1029/2008FL034913.

ocean where deeper, lower pH water upwell.³⁵ Other measures might include habitat restoration/creation, such as planting seagrass. Unless a massive global effort is mounted, these techniques will at best be effective only on a very local scale. In addition, manipulation of ocean chemistry has the potential to damage or otherwise alter the marine environment and ecosystems. Reducing CO₂ emissions to the atmosphere and/or removing CO₂ from the atmosphere (i.e., carbon sequestration) currently appear to be the only practical ways to minimize the risk of large-scale and long-term changes to the acidity of marine waters. Because of the continuing increase in CO₂ levels in the atmosphere, and its residence time there, acidification of the oceans will likely continue for a long time. Even if atmospheric CO₂ were to return to pre-industrial levels, it would likely take tens of thousands of years for ocean chemistry to return to a condition similar to that occurring at pre-industrial times more than 200 years ago.³⁶

What Is the Federal Government Doing About Ocean Acidification?

Much of the current federal attention to ocean acidification focuses on research to better understand the chemical processes involved and to become better able to predict how ocean ecosystems might respond to increasing acidification.

The National Science Foundation (NSF) was the first federal agency to become involved in research related to ocean acidification. The modern surveys of CO₂ status in the oceans can be traced to the NSF-sponsored Joint Global Ocean Flux Study (JGOFS), which originated in recommendations from a National Academy of Sciences workshop in 1984.³⁷ The more modern concerns over ocean acidification arose from a May 2004 Paris workshop chaired by the now-president of the National Academy of Sciences, Ralph Cicerone.³⁸ In April 2005, NSF, the National Oceanic and Atmospheric Administration (NOAA), and the U.S. Geological Survey sponsored a workshop on the impacts of ocean acidification on coral reefs and other marine calcifiers.³⁹ In October 2008, NOAA and NSF commissioned an 18-month comprehensive national study by the National Research Council of the National Academy of Sciences of how CO₂ emissions absorbed into the oceans may be altering fisheries, marine mammals, coral reefs, and other natural resources.⁴⁰ This study was required by Section 701 of P.L. 109-479.⁴¹

³⁵ L. D. D. Harvey, "Mitigating the Atmospheric CO₂ Increase and Ocean Acidification by Adding Limestone Powder to Upwelling Regions," *Journal of Geophysical Research*, v. 103 (2008): C04028, 21 p.

³⁶ The Royal Society, *Ocean Acidification due to Increasing Atmospheric Carbon Dioxide*, Policy Document 12/05 (June 2005), 60 p.

³⁷ For additional background, see <http://www1.whoi.edu/jgofMission.html>.

³⁸ Ralph Cicerone, "The Ocean in a High CO₂ World," *Eos*, v. 85, no. 37 (September 14, 2004): 351, 353.

³⁹ J. A. Kleypas et al., *Impacts of Ocean Acidification on Coral Reefs and Other Marine Calcifiers: A Guide for Future Research*, report of a workshop held April 18-20, 2005, St. Petersburg, FL, sponsored by NSF, NOAA, and the U.S. Geological Survey (June 2006), 88 p.

⁴⁰ See http://www.noaa.gov/stories2008/20081020_oceanacid.html.

⁴¹ This measure requires the Secretary of Commerce to request that the National Research Council study acidification of the oceans and how this process affects the United States.

A variety of programs conducted within the NOAA help to gain a better understanding of ocean acidification.⁴² The Pacific Marine Environmental Laboratory's CO₂ shipboard measurements and monitoring buoys provide data that help discern seasonal changes in the oceanic carbon system. The Atlantic Oceanographic and Meteorological Laboratory monitors changes in CO₂ and pH through the use of chemical sensors on ships and moorings. NOAA's Repeat Hydrography Program provides data on the large-scale changes to the carbon system and rates of ocean acidification over decadal time scales. Sea Grant supports research on the effects of ocean acidification on coral reefs in Hawaii. NOAA's Geophysical Fluid Dynamics Laboratory participated in the Ocean-Carbon Cycle Model Intercomparison Project (OCMIP2) to develop an international collaboration to improve the predictive capacity of carbon cycle models. NOAA's Alaska Fisheries Science Center has been conducting exposure studies of blue king crab larval survival due to reduced pH and has developed an ocean acidification research plan.⁴³ NOAA's Southwest Fisheries Science Center has been evaluating the long-term impacts of low pH on marine plankton in the California Current and off Antarctica. Projects funded by the NOAA Global Carbon Cycle program at NOAA laboratories and universities provide information to address the CO₂ and pH changes in the ocean. NOAA estimates that its annual expenditures for these ocean acidification research and related programs in FY2008 were about \$4.3 million; for FY2009, NOAA has requested \$4.06 million.⁴⁴

The National Aeronautics and Space Administration (NASA) has a number of space projects that contribute to a better understanding of ocean acidification.⁴⁵

What Is the Congressional Interest in Ocean Acidification?

Congressional attention is focused primarily on addressing the cause of ocean acidification—increasing atmospheric CO₂.⁴⁶ Legislative attention to ocean acidification focuses on authorizing and funding research to increase knowledge about ocean acidification and its potential effects on marine ecosystems.

In the 111th Congress, the Federal Ocean Acidification Research And Monitoring Act of 2009 (Title XII, Subtitle D, of P.L. 111-11) directed the Secretary of Commerce to establish an ocean acidification program within NOAA, established an interagency committee to develop an ocean acidification research and monitoring plan, and authorized appropriations through FY2012 for NOAA and the National Science Foundation.

⁴² Testimony by Dr. Richard A. Feely, NOAA Office of Oceanic and Atmospheric Research, before the House Committee on Science and Technology, Subcommittee on Energy and Environment, June 5, 2008.

⁴³ Alaska Fisheries Science Center, *Forecast Fish, Shellfish, and Coral Population Responses to Ocean Acidification in the North Pacific Ocean and Bering Sea*, National Marine Fisheries Service (Juneau, AK: August 2008), AFSC Processed Report 2008-7, 35 p.

⁴⁴ Personal communication from Jonathan Kelsey, Congressional Affairs Specialist, Office of Legislative Affairs, NOAA, (202) 482-0809, Dec. 11, 2008.

⁴⁵ See <http://oco.jpl.nasa.gov/>, <http://modis.gsfc.nasa.gov/about/>, <http://oceancolor.gsfc.nasa.gov/SeaWiFS/>, and <http://so-gasex.org/media.html>.

⁴⁶ See CRS Current Legislative Issue "Climate Change", available at http://apps.crs.gov/cli/cli.aspx?PRDS_CLI_ITEM_ID=2645&from=3&fromId=2522.

In comparison to previous sessions of Congress, legislative interest in ocean acidification expanded significantly in the 110th Congress.⁴⁷

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⁴⁷ For more details on ocean acidification legislation in the 110th Congress, see the section “Climate Change” in CRS Report RL33813, *Fishery, Aquaculture, and Marine Mammal Legislation in the 110th Congress*, by Eugene H. Buck.