

EMSP Project ID Number 60123

Agreement Number: DE-FG07-97ER14819

POTENTIAL MODULATED INTERCALATION OF ALKALI CATIONS INTO METAL
HEXACYANOFERRATE COATED ELECTRODES

PROGRESS REPORT: 2/18/02

Principal Investigator

Daniel T. Schwartz
University of Washington
Department of Chemical Engineering
Box 351750
Seattle, WA 98195-1750
TEL: 206 685-4815; E-Mail: dts@u.washington.edu

Co-Investigator(s): None

Research Personnel Supported (in part) by Project in Final Year

Graduate students

- Kavita M. Jeerage
- William A. Steen
- Haixia Dai

Postdoctorate

- Qiuming Yu

Specific DOE Problems Addressed by Project

The separation of radiocesium from other cations has important cost and safety implications for the processing of various cesium contaminated wastes forms such as basin waters and sludges (e.g., at Hanford and Savannah River sites), as well as groundwater and more acidic tank wastes (e.g. at the Idaho Site). The key feature of these sites being their cesium contamination in a moderate to low pH aqueous solvent.

Overall Research Objectives

Nickle hexacyanoferrate is a polynuclear inorganic ion intercalation material that loads (intercalates) and elutes (deintercalates) alkali cations from its structure when electrochemically reduced and oxidized, respectively. Nickel hexacyanoferrate (NiHCF) is known to preferentially intercalate cesium over all other alkali cations, thus providing a basis for a separation scheme that can tackle DOE's radiocesium contamination problem. This program studied fundamental issues in alkali cation intercalation and deintercalation in nickel hexacyanoferrate compounds, with the

goal of (1) quantifying the ion exchange selectivity properties for cation mixtures, (2) enhancing ion exchange capacities, and (3) and understanding the electrochemically-switched ion exchange process (ESIX).

Year 4 Progress Report and Implications

The final year (a no-cost extension) of this project has been extremely productive. Eight peer reviewed publications have appeared, are in press, or are at some state of the editorial process (see below). Below, we separately address our progress toward each of the goals of this research.

Most of our effort has been directed toward goal (1), where the overriding question has been, "what materials do we make by electrodeposition, and how do their properties depend on stoichiometry and structure?" Publications, (1)-(5), (7), and (8) have tackled the issue by developing appropriate measurement techniques that can get at properties and structure of thin nickel hexacyanoferrate films. A combined electrochemical, Raman spectroscopy, and energy dispersive x-ray spectroscopy method has been developed for analyzing the quantity and type of ion loaded into NiHCF films. In journal article (2), (3), and (5) cited below, we describe the methodology and show the range over which this technique can be used to quantify separation factors in NiHCF coated electrodes. We find that the distribution coefficient for NiHCF coated electrodes can exceed 3×10^4 ml/g in a solution of 10^{-6} M Cs^+ with excess sodium nitrate (1 M Na^+). However, cited paper (5) shows that the ion selectivity is strongly dependent on the oxidation state of nickel hexacyanoferrate, where the highest cesium selectivities are displayed by the oxidized matrix and the lowest selectivities are found in the reduced matrix. The origin of the charge-density dependence is $\text{Cs}^+ - \text{Cs}^+$ repulsion in the reduced (fully loaded) matrix. These experimental results are complemented by molecular dynamic simulations (paper (4)), where we explore the role of solid structure on the nature of ion and solvent intercalation. Together, these papers point to the role of size exclusion in nickel hexacyanoferrate, with more ideal structure-function relationships existing in nickel hexacyanoferrate thin films that possess a high density of ferrocyanide defects in the solid.

For goal (2), all of the papers presented below (except (1)) used the cathodic deposition of nickel hexacyanoferrate on noble metal electrodes as a high capacity means of synthesis. In this approach, NiHCF is deposited cathodically by cycling a platinum electrode between +850 and 0 mV vs. SCE at a scan rate of 25 mV/S in freshly prepared solution of 4 mM NiSO_4 , 4 mM $\text{K}_4\text{Fe}(\text{CN})_6$ and 500 mM Na_2SO_4 . The resulting thin films can routinely be produced with ion intercalation capacities exceeding 8 mC/cm^2 . Previously, our anodic deposition scheme (used in paper (1)) produced a typical capacity of 3 mC/cm^2 , so cathodic deposition represents a roughly 3 fold improvement in capacity.

Our effort to address Goal (3) has been to develop macroscopic batch contacting models for the ESIX process. This work is represented by paper (6) below. The model we've developed takes into account the electrochemical nature of the experiment, and all physical processes that are involved in an equilibrium process (in this case, rate parameters are ignored). We show that ESIX has the possibility of recovering a high fraction of the target ion from a binary mixture, and concentrate it into a product stream (something conventional ion exchange cannot do), but to do

so requires a combination of selectivity and ion elution efficiency that has not yet been achieved, but is clearly possible.

Information Access

- (1) D.T. Schwartz and S.M. Haight, Transport and chemistry at electroactive interfaces studied using line-imaging Raman spectroscopy, *Colloid Surf. A* **174**, 209-219 (2000).
- (2) K.M. Jeerage and D.T. Schwartz, Characterization of cathodically deposited nickel hexacyanoferrate for electrochemically switched ion exchange, *Sep. Sci Tech.* **35**, 2375-92 (2000).
- (3) K.M. Jeerage, W.A. Steen, and D.T. Schwartz, Correlating nanostructure with ion intercalation in electrodeposited nickel hexacyanoferrate thin films, *Chem. Mat.* **14**, 530-535 (2002).
- (4) Q. Yu, W.A. Steen, K.M. Jeerage, S. Jiang, and D.T. Schwartz, Structure-dependent solvent and ion intercalation in reduced and oxidized nickel hexacyanoferrates, *J. Electrochem. Soc.*, **in print** (2002).
- (5) K.M. Jeerage, W.A. Steen, and D.T. Schwartz, Charge-density-dependent partitioning of Cs⁺ and K⁺ into nickel hexacyanoferrate matrices, *Langmuir*. **accepted** (2002).
- (6) K.M. Jeerage and D.T. Schwartz, An equilibrium stage model for electrochemically-switched ion exchange separations, *Ind. Eng Chem. Res.* **submitted** (2001).
- (7) W.A. Steen, K.M. Jeerage, and D.T. Schwartz, Spectroscopic determination of nickel hexacyanoferrate oxidation state, *Appl. Spec.*, **submitted** (2002)
- (8) W.A. Steen, Q. Yu, and D.T. Schwartz, Structure of electrodeposited nickel hexacyanoferrate by x-ray diffraction and extended x-ray absorption fine-structure, *J. Electrochem. Soc.*, **in preparation** (2002).