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# Characterization of Soluble Organics in Produced Water

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Chemical Technology Division

**CHARACTERIZATION OF SOLUBLE ORGANIC? IN PRODUCED WATER**

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## LIST OF ABBREVIATIONS

DRO	Diesel range organic compounds
GC	Gas chromatography
GOM	Gulf of Mexico
FID	Flame ionization detector
GRO	Gasoline range organic compounds
HEX	Hexane fraction
HP	Hewlett Packard
IC	Ion chromatography
ICP-AES	Inductively coupled plasma-atomic emission spectrometry
IR	Infrared
ISA	Ionic adjustor solution
ISE	Ion selective electrode
K-D	Kudema-Danish
LC	Liquid chromatography
MS	Matrix spike sample
MSD	Mass selective detector
ORNL	Oak Ridge National Laboratory
PAH	Polynuclear aromatic hydrocarbons
PERF	Petroleum Environmental Research Forum
PI	Principal investigator
ppm	Parts per million (by weight or volume)
svocs	Semi-volatile organic compounds
SW	Statement of Work
TARO	Total aromatic fraction
TDM	Total dissolved material
TEM	Total extractable material
TPH	Total petroleum hydrocarbons
TPOL(gc)	Total polar fraction determined by gas chromatographic analysis
TPOL( sub)	Total polar fraction determined by subtraction from TEM fraction data
TSAT	Total saturate fraction
USEPA	United States Environmental Protection Agency
WAF	Water accumulated fraction
wso	Water soluble organic compounds



## EXECUTIVE SUMMARY

Soluble organics in produced water and refinery effluents represent treatment problems for the petroleum industry. Neither the chemistry involved in the production of soluble organics nor the impact of these chemicals on total effluent toxicity is well understood. The U.S. Department of Energy provides funding for Oak Ridge National Laboratory (ORNL) to support a collaborative project with Shell, Chevron, Phillips, and Statoil entitled "Petroleum and Environmental Research Forum project (PERF 98-04: "Manage Water-Soluble **Organics** in Produced Water"). The goal of this project, which involves characterization and evaluation of these water-soluble compounds, is aimed at reducing the future production of such contaminants.

To determine the effect that various drilling conditions might have on water-soluble organics (WSO) content in produced water, a simulated brine water containing the principal inorganic components normally found in Gulf of Mexico (GOM) brine sources was prepared. The GOM simulant was then contacted with as-received crude oil from a deep well site to study the effects of water cut, **produced-water pH**, salinity, pressure, temperature, and crude oil sources on the type and content of the WSO in produced water. The identities of individual semivolatile organic compounds (SVOCs) were determined in all as-received crude and actual produced water samples using standard USEPA Method (8270C) protocol. These analyses were supplemented with the more general measurements of total petroleum hydrocarbon (TPH) content in the gas ( $C_6-C_{10}$ ), diesel ( $C_{10}-C_{20}$ ), and oil ( $C_{20}-C_{28}$ ) carbon ranges as determined by both gas chromatographic (GC) and infrared (IR) analyses. An open liquid chromatographic procedure was also used to differentiate the saturated hydrocarbon, aromatic hydrocarbon, and polar components within the extractable TPH. Inorganic constituents in the produced water were analyzed by ion-selective electrodes and inductively coupled plasma (ICP)-atomic emission spectrometry (AES).

The WSO found in produced water samples was primarily polar in nature and distributed between the low and midrange carbon ranges. Typical levels of total extractable material (TEM) was about 20 **mg/L**; that associated with the aromatic fraction was present at 0.2 **mg/L** and that in the saturated hydrocarbon fraction was present at less than 0.02 **mg/L**. Formic, acetic, and propionic acids were also found in the produced water, occurring at a total concentration of 30 **mg/L**. It was estimated that the presence of 30 **mg/L** organic acids would artificially overstate TEM content by 2 **mg/L**.

Of the five tested parameters, the factor that most controlled the total WSO in produced water was that of aqueous phase **pH**. Beyond a value of **pH7** significant quantities of  $C_{10}-C_{20}$  range material become markedly soluble as they deprotonate in a basic aqueous phase. Both the absolute and relative volumes of GOM brine and crude additionally affected total WSO. Produced water appeared to reach a saturation level of WSO at a 50% water/oil ratio. Pressure slightly enhanced WSO by increasing the relative quantity of  $C_6-C_{10}$  range material. Temperature primarily altered the relative ratio of carbon ranges within the WSO without significantly elevating the total WSO in the GOM brine. Salinity had the least affect on the chemical character or the carbon size of WSO in produced water.



## 1. BACKGROUND

Soluble organics in produced water and refinery effluents represent treatment problems for the petroleum industry. Production **installations** and refineries take special efforts to meet increasingly stringent regulatory discharge requirements for dissolved organics. Early data from the Gulf of Mexico (GOM) wells indicate that the more polar hydrocarbons from deep-water drilling will significantly increase the dissolved hydrocarbon content in produced water from these sites. Neither the chemistry involved in the production of soluble organics nor the impact of these chemicals on total effluent toxicity is well understood.

The U.S. Department of Energy provides funding for Oak Ridge National Laboratory (ORNL) to support a collaborative project with Shell, Chevron, Phillips, and **Statoil**, entitled "Petroleum and Environmental Research Forum Project (PERF 98-04: "Manage Water-Soluble **Organics** in Produced Water"). The goal of this project, which involves characterization and evaluation of these water-soluble compounds, is aimed at reducing the future production of such contaminants. The information obtained in the project will increase our understanding of the generation of water-soluble organics (WSO) and ultimately result in strategies to minimize WSO production and lead to the development of guidelines for effluent treatment. This, information **will be used** to design the next generation of offshore platforms for deep-water wells; and could also help reduce construction costs.

The first phase of the task was to acquire characterization data for produced water. A subsequent task will be to use the data to construct a model that can be used to predict WSO content based on the composition of a specific crude source. This report describes 'results acquired for phase one of the task. Analysis of produced water was performed at both **ORNL and Statoil** Company. ORNL characterized WSO derived from sources of **GOM** crude, whereas **Statoil** accumulated characterization data from North Sea oil sources. The objective of this report is to provide the sponsor of the project with a description of the analytical procedures used and the data acquired by ORNL. The format of the document follows closely the document submitted by staff at **Statoil**. This was done in order to aid the customer in compiling data from both ORNL and **Statoil** studies. The format of the report is designed to:

- 1) briefly describe the scope of the study (Chapter 1);
- 2) report the sampling of the crude by industrial partners (Chapter 2) ;
- 3) provide a detailed discussion of analytical procedures (Chapters 3-6),
- 4) validate analytical results through a discussion of QC data (Chapter 7); and, finally,
- 5) present WSO characterization data derived from the contact **of the GOM crude** with brine (Chapters 8, 9 and appendix tables).

Because the primary objective of the project was to **provide a database** from which a second party would use to design a mathematical model to predict **WSO** content in produced water, the pertinent data was set aside as a separate group of tables in the appendix. All tables within the report text summarize information concerning objectives, analytical procedures, QC performance, results non-essential to the data base (although included as points of interest), etc. This approach in format should facilitate modeling by obviously separating types of data and clearly defining information that will be incorporated into the predictive model.

Finally, several limitations were placed on the format this document due to information **held** proprietary by the PERF committee. The well site from which the crude was **derived** is not specifically named; it is only stated as Deep 1 Well. This report tabulates WSO data information specific to this well site. The Deep 1 Well data amplifies a proprietary report written for the

PERF members that summarizes general levels and character of WSO found in national and international literature. Because the proprietary nature of the summary report and its availability to PERF committee members, no effort was made in this document to compare Deep 1 Well results to general trends. Comparisons were made in closed session with PERF committee members. The primary goal of this document is data compilation for the project.

## 1.1 ANALYTICAL APPROACH

Quantitative characterization data were needed as the first step in understanding the generation of WSO in produced water. To achieve this objective, industrial partners provided ORNL samples of GOM deep-water **crudes** and their associated produced water. Together with industrial partners, ORNL selected methods to characterize the WSO in both actual and simulated samples of produced water. The identities of individual semi-volatile organic compounds (**SVOCs**) were determined in all as-received crude and actual produced water samples using standard USEPA Method (8270C) protocol. These analyses were supplemented with the more general measurements of total petroleum hydrocarbon (TPH) content in the gas ( $C_6$ – $C_{10}$ ), diesel ( $C_{10}$ – $C_{20}$ ), and oil ( $C_{20}$ – $C_{28}$ ) carbon ranges as determined by both GC (USEPA SW-846 Method 8015B) and infrared (IR) analyses. An open liquid chromatographic procedure was used to differentiate the saturated hydrocarbon, aromatic hydrocarbon, and polar WSO components contained in the TPH extract. Inorganic constituents in the produced water were analyzed by ion-selective electrodes and inductively coupled plasma (ICP)-atomic emission spectrometry (**AES**).

Since the methods primarily follow those established by the U.S. Environmental Protection Agency (**USEPA**), the terminology used in this report follows that of **USEPA** regulatory nomenclature. The stated carbon ranges ( $C_6$ – $C_{10}$ ,  $C_{10}$ – $C_{20}$ ,  $C_{20}$ – $C_{28}$ ) refer to sample components eluting from a gas chromatograph between **n-alkane** standards within each of these carbon size ranges. The terminology does not define the identity of the sample component eluting, merely that it has a volatility within the given saturated **alkane** range. The regulatory nomenclature is specific to the **USEPA** method used, in this case **USEPA** SW-846 Method 8015B. A description of the analytical procedures is presented in chapters 3 through 6 to fully explain under what condition the data was acquired.

TPH is also regulatory nomenclature, referring to all components extracted using either **USEPA** SW-846 Method 8015B or **USEPA** SW-846 Method 1664. Since polar WSO is also extracted in these procedures, TPH results include some mass that is actually not petroleum hydrocarbons. One of the goals of this study was to estimate the positive bias introduced into TPH results from these non-TPH components.

## 1.2 EXPERIMENTAL APPROACH

Only a limited volume of produced water was present in the crude oil samples shipped from drilling sites. To determine the effect that various drilling conditions might have on WSO content in produced water, a simulated brine water containing the principal inorganic components normally found in GOM brine sources was prepared. The GOM simulant was then contacted with as-received crude oil to study the effects of water cut, produced-water **pH**, salinity, pressure, temperature, and crude oil sources on the type and content of the WSO in produced water. The test plan involved the analysis of the WSO content of simulated produced-water samples derived from oil/brine contacts covering the following operating ranges:

- pressures of 1-100 bar;
- temperatures of 90–200°F (25-100°C);
- pH level of  $6.5 \pm 2$ ;
- water cuts of 20, 50, and 80%; and
- salinities of 35,000–150,000 ppm.

The test matrix presented in Table 1.1 below was used for the Deep1 crude oil source. The specific name for each experiment associated with a given experimental variable is denoted as “PEW-” in the table. A complete description of each PEW experiment and its associated data are summarized in the appendix of this document.

**Table 1.1. Experimental test matrix for crude oil/GOM brine contacts**

Experimental variable	Pressure	Temperature (°C)	Produced H <sub>2</sub> O source	PH	Salinity (ppm)	Water cut (%)
As received (PERIL8, 16, 20)	Ambient	25	GOM simulant	7	63,000	80
Water cut (PERP-6)	Ambient	25	GOM simulant	7	63,000	67
Water cut (PERF-7, 10)	Ambient	2 5	GOM simulant	7	63,000	50
Water cut (PERF-9)	Ambient	25	GOM simulant	7	63,000	20
(PEE-13)	Ambient	25	GOM simulant	8.1	63,000	80
(PEE-1 1)	Ambient	2 5	GOM simulant	9.0	63,000	80
pH (PERP-12)	Ambient	2 5	GOM simulant	4.7	63,000	80
pH (PERF-19)	Ambient	2 5	GOM simulant	6.0	63,000	80
Salinity (PERP-14)	Ambient	25	GOM simulant	7	41,000	80
Salinity (PERF-15)	Ambient	25	GOM simulant	7	115,000	80
Temperature (PERF-17)	Ambient	47	GOM simulant	7	63,000	80
Temperature (PERF-18)	Ambient	75	GOM simulant	7	63,000	80

**Table 1.1. Experimental test matrix for crude oil/GOM brine contacts (cont.)**

Pressure (PERF-26)	14 bar	50	GOM simulant	7	63,000	80
Pressure (PERF-27)	35 bar	50	GOM simulant	7	63,000	80
Pressure (PEW-28)	60 bar	50	GOM simulant	7	63,000	80

The distribution and partitioning of WSO components were then calculated for the operating ranges. The resulting information was forwarded to industrial partners for incorporation into a model to predict the production of water-soluble **organics** as a function of crude composition and formation characteristics.

## **2. COLLECTION AND SAMPLING OF ACTUAL OIL AND PRODUCED WATER SAMPLES**

### **2.1 FIELD SAMPLING OF DEEP1 CRUDE OIL**

Two 5-gallon samples were collected directly from the Deep1 Well flowline; these samples were transferred directly into 5-gal DOT-approved containers. According to field notes, the crude oil had a 22.2 API gravity at a collection temperature of 75°F. Each sample was collected slowly over a 1-h interval to allow the oil to degas. While the samples were being collected, the operators performed a shakeout test, which gave a 10% water cut when heated with solvent at 140°F.

### **2.2 DETERMINATION OF WATER CONTENT IN AS-RECEIVED OIL SAMPLES**

Crude oil samples were shipped from the drilling site in 5-gal barrels equipped with two bung-hole ports in the lid. The ports were opened slowly to vent the slight overpressure of the barrels. Any actual produced water was collected by dropping weighted tubing (Tygon® F40440-A) down to the bottom of the slightly tipped barrel. A peristaltic pump, operating in the forward direction, was used to create a slight positive pressure to displace any oil from the tubing prior to withdrawing the water phase. Pumping was continued until all water had been collected; the volume of water was then estimated relative to the initial volume of sample present in the barrel.

In the case of Deep 1 crude, a separate aqueous phase could not be removed directly, so 0.5 to 1 L of oil was withdrawn from the bottom of the barrel and allowed to settle in a glass bottle. Again a separate aqueous phase did not form. Ten-mL aliquots of withdrawn oil were then transferred to 15-mL tapered centrifuge tubes. The oil was then acidified to 0.001 M HCl by the addition of 10 µL of 1 M HCl to separate any produced water that might have emulsified during shipment. When an aqueous phase was not observed at the base of the tube after 15 min, the acidified oil samples were centrifuged at 2000 rpm for 20 min to enhance phase separation. No distinct water phase was noted at the base of the tube, indicating the water content was too low to visually measure. Results acquired using an automated Karl Fischer iodometric titrator indicated  $8.9 \pm 0.4$  % water content, as opposed to field testing results of 10-12 %.



## 2.3 LABORATORY SAMPLING 'OF AS-RECEIVED CRUDE OIL

Consistent testing of the WSO contents in the oil/water mixing experiments was predicated on the use of homogeneous oil samples. Toward this end, a 5-gal barrel of oil was placed horizontally on a barrel roller and rotated at the rate of 8 rpm for 2 h. The barrel was then set upright and oil was withdrawn using submersed tubing and a peristaltic pump. No entrained air bubbles were observed in the withdrawn oil.

## 3. EXPERIMENTAL PROTOCOL FOR AMBIENT PRESSURE TESTING

Simulated produced water samples were prepared under ambient pressure conditions using a Water Accumulated Fraction (WAF) vessel. The WAF procedure is typically used for the aquatic toxicity testing of **acute** and chronic exposures of marine life to lubricants in **contact** with water sources. The glass vessel was fabricated according to dimensions stipulated in ASTM D6081, in which the height:diameter ratio of the liquid phases in the 5-L vessel can be varied from 1: 1 to 2: 1. Testing was initiated by first adding up to 3 L of GOM simulant to the vessel. Up to 4 L of oil were slowly drained down the side of the vessel so that it was layered over the water without undo turbulence. Typical water/oil ratios used in testing varied from 20–80%. A baseline value of 80% was used for testing all experimental variables except water cut. A Teflon®-coated stirring bar submersed in the GOM simulant was used to stir the oil/water interface such that the vortex of the oil phase descended to a depth of approximately 15–30% into the aqueous phase. Although the ASTM protocol indicates that a 24-h mixing time is sufficient to reach equilibrium extraction of WSO, a mixing time of 4 days was typically used to duplicate the experimental setup employed by Statoil in the testing of North Sea oil samples. Produced water was collected at the end of the equilibration time by opening a Teflon® stopcock located at the base of the WAF vessel. If necessary, any dispersed oil in the withdrawn water was removed using a separatory funnel. Several 20-mL aliquots of the produced water were stored in glass bottles at -4°C for subsequent analysis of carbonate/bicarbonate, inorganic metals, chloride, and low-molecular weight organic acids. One liter of produced water was then acidified with concentrated hydrochloric acid (HCl) to bring the pH level to less than 2. Surrogate recovery standards were added to the produced water fraction prior to the solvent extraction of WSO.

## 4. EXPERIMENTAL PROTOCOL FOR HIGH PRESSURE TESTING

A factory-assembled pressure vessel incorporating armored sight glasses on opposing walls, was used to generate produced water samples at pressures up to 1000 psi and 50°C. This vessel having an internal volume of 300 mL, is fabricated of mild steel, that has been certified as to chemical composition, hardness, and tensile strength. Additional hardware supplied by Pressure Products Company, Inc. (4540 W. Washington St., Charleston, W. VA 25313) included a 1200 psi pressure relief valve, a 1-1000 psi liquid-filled pressure gauge, an inlet for a digital thermometer, an inlet to pressurize the vessel with helium, and tubing and needle valves to allow filling and draining of solution from the vessel (Figs. 4.1 and 4.2). The 60-lb pressure vessel was clamped onto a Ro-Tap Model B Sieve Shaker (C. M. Tyler Co.), which rotated the system horizontally at approximately 60 rpm to produce a slight vortex at the oil/water solution interface. During mixing, the temperature was maintained at 50°C by controlling the voltage output to a heating tape wrapped around the circumference of the vessel.

Two hundred mL of brine simulant were added to the heated vessel; the system was then pressurized with helium, heated, and mixing for 4 days to determine blank concentration levels for TPH in the vessel. A Hoke® sample chamber was connected to the gas sampling port of the

vessel. at the end of the equilibration time. Pressure was vented into the evacuated gas sample vessel by opening the gas sampling valve located at the top of the pressure chamber. **Once** the chamber had been brought to ambient pressure, the **Hoke®** vessel was sealed. The exit valve located at the base of the pressure chamber was subsequently opened to collect the water phase.

The gases that collected in the **Hoke®** sampling vessel were vented through an -5 cm length of **¼-inch** stainless steel tube packed with sorbents to retain hydrocarbon components and then through a mass flow meter. The tube contained Carbotrap C to capture high-molecular-weight material, followed by Carbotrap for the majority of components and, finally, **Carbosieve SIII** for low (**C<sub>2</sub>–C<sub>3</sub>**) molecular-weight **organics**. These sample tubes have been sealed and archived should the need for their analysis exist.

Once the TPH blank levels of pressure vessel were suitably low, actual oil/brine experiments were conducted by loading the heated chamber with 200 mL of GOM brine simulant and 30 mL of crude oil. The first experiment was performed at ambient pressure and 50°C for comparison with data derived from oil/water contacts in the WAF vessel. The remaining experiments were conducted at pressures of 14, 35, and 60 bar and a temperature of 50°C. Again, the gases in the chamber were vented into a gas sampling vessel to depressurize the pressure chamber at the end of a 4-day equilibration time. Immediately after being depressurized, the produced water phase was collected at the drain port for subsequent analysis using the open LC fractionation protocol (see Sect. 6.4).

## 5. SYNTHETIC GOM BRINE

### 5.1 PREPARATION OF SIMULANT

The staff of the Phillips Company provided the recipe for the preparation of GOM brine simulant. Table 5.1 summarizes the average concentrations of actual GOM brine components as derived from data members of Phillips staff members have collected from 300 past and present GOM wells. These accumulated data provided the basis for the recipe formulation.

**Table 5.1. Average concentrations of GOM brine components<sup>a</sup>**

GOM brine Component	Concentration (mg/L)
Cl <sup>-</sup>	63,000
SO <sub>4</sub> <sup>2-</sup>	3
HCO <sub>3</sub> <sup>-</sup>	220
Na <sup>+</sup>	35,354
Ca <sup>2+</sup>	3,260
Mg <sup>2+</sup>	9 1 0
Ba <sup>2+</sup>	95
Sr <sup>2+</sup>	131
Fe <sup>3+</sup>	23
Dissolved CO <sub>2</sub>	68
Temperature	90°F
pH	6.7

<sup>a</sup>Data provided by D. Bourg, Philips Co., May 9, 2000.

Table 5.2 summarizes the results of a spreadsheet devised by Phillips staff that delineates the chemical components and their quantities required to prepare 1- and 4-L batches of produced water simulant. Once prepared, the simulant was stored in a refrigerator and then brought to the required experimental temperature just before use.

## 5.2 VERIFICATION OF GOM BRINE SIMULANT CONTENT

Ion chromatography (IC) and ion-selective electrodes were used to **determine anion** concentrations in the prepared simulant brine. ICP-AES was used to verify **dissolved** metal components. Complete descriptions of the analytical procedures used for cation/anion determinations can be found in Sects. 6.6 and 6.7 of this document. IC was initially selected for the analysis of brine salinity; however, results based on IC appeared to negatively bias the total chloride concentration in GOM brine simulant. The high sodium concentration in the simulant mandated that the brine samples be chemically treated with a cation exchange resin column to reduce the level of sodium ion before the sample was introduced into IC equipment. As can be seen in Table 5.3, significant error was introduced by the pretreatment procedure since the electrochemical balance of the cation content in the simulant (Table 5.4) indicated that the chloride content should be approximately 65,000 ppm. Chloride analysis based on the use of a chloride selective electrode was considered more accurate for this application because this analysis is typically performed in the presence of 5 M NaNO<sub>3</sub>, a matrix having an ionic strength similar to that of the simulant brine.

Ferrous ion was initially included at a concentration of 23 ppm to reflect the concentration present in actual GOM brine. However, precipitation of ferric oxide at near neutral pH required filtration of the simulant prior to use. The final soluble iron concentration in the filtrate was less than the detection limit of the ICP. Therefore, ferrous sulfate was not included in simulant preparation for any of the oil/water contact experiments in this study.

**Table 5.2. Preparation of Gulf of Mexico brine simulant**

Component	1-L Preparation requires (g)	4-L Preparation Requires (g)
NaCl	89.708	358.832
CaCl <sub>2</sub> • 2H <sub>2</sub> O	11.959	47.835
MgCl <sub>2</sub> • 6H <sub>2</sub> O	7.613	30.453
BaCl <sub>2</sub> • 2H <sub>2</sub> O	0 . 1 6 9	0.676
SrCl <sub>2</sub> • 6H <sub>2</sub> O	0 . 3 9 9	1.595
FeSO <sub>4</sub> • 7H <sub>2</sub> O	0.000	0.000
NaHCO <sub>3</sub>	0 . 3 0 3	1.212
Na <sub>2</sub> SO <sub>4</sub>	0.063	0.252
H <sub>2</sub> O	<b>889.740</b>	3,559.0
Final solution (g)	<b>1,000.0</b>	<b>3,999.8</b>
Total dissolved material (TDM, ppm)	102,971	102,971
Salinity as Cl <sup>-</sup> (ppm)	62,975	62,975

**Table 5.3. Anion concentrations in GOM brine**

Anion	Actual GOM brine concentration (ppm)	Simulant Analysis Results (ppm)	
		Ion Chromatography	Ion Selective Electrode
SO <sub>4</sub> <sup>2-</sup>	3	<10	NA <sup>a</sup>
Cl <sup>-</sup>	62,975	41,500	62,600 ± 1,700
HCO <sub>3</sub> <sup>-</sup>	220	NA	153 ± 1

<sup>a</sup>NA = not analyzed.

**Table 5.4. Cation concentrations (mg/L) in GOM brine**

	Cation (mg/L)					
	Na <sup>+</sup>	Ca <sup>2+</sup>	K <sup>+</sup>	Sr <sup>2+</sup>	Ba <sup>2+</sup>	Fe <sup>3+</sup>
Actual brine	35,400	3,260	0.0	0.131	0.095	0.023
Simulant by ICP	35,400	3,270	<0.004	0.139	0.044	<0.020

## 6. ANALYTICAL METHODS

### 6.1 IDENTIFICATION OF SEMI-VOLATILE ORGANIC COMPOUNDS (SVOC) IN ACTUAL OIL AND PRODUCED WATER BY GC/MS

Specific chemical compounds were identified and quantitated in as-received crude oils and actual produced water using USEPA SW-846 Method 8270 C. (This method is also identified as USEPA Method 525.2 of the “500 series” in the USEPA compendium for the analysis of contaminants in drinking water.) The protocol uses GC-MS to identify a core group of over 45 semi-volatile organic target compounds in environmental samples. Within the list of target compounds, those that are of particular interest to petrochemical applications are included in Table 6.1:

**Table 6.1. USEPA Semi-volatile organic compounds of interest to petrochemical applications**

Acenaphthene	Fluoranthene
Acenaphthylene	Fluorene
Anthracene	Indeno(1,2,3-c,d)pyrene
Benzo(a)anthracene	2-Methylnaphthalene
Benzo(a)pyrene	2-Methyl phenol
Benzo(b)fluoranthene	Naphthalene
<b>Benzo(e)pyrene</b>	Phenanthrene
Benzo(k)fluoranthene	Pyrene
Chrysene	

Although polynuclear aromatic hydrocarbons (PAHs) are included in the list due to the toxicity imparted by these compounds to environmental samples, the PERF activity reported here did not

address toxicity concerns as part of its scope of work. Therefore, PAH data were not acquired in samples derived from GOM wells.

Crude oil and actual produced water samples were analyzed by Mountain States Analytical, Inc. (1645 West 2200 South, Alt Lake City, Utah 84119). The protocol from this off-site testing facility includes the analyses of over 120 semi-volatile compounds pertinent to the petroleum industry. Protocol requirements include the addition of surrogate recovery standards (see Table 6.2) to samples prior to their extraction with methylene chloride. These standards are used to monitor the variability in extraction efficiency of the solvent for individual target compounds.

**Table 6.2. Surrogate recovery standards for SVOC analysis**

2,4,6-Tribromophenol  
2-Fluorophenol  
2-Fluorobiphenyl  
Nitrobenzene-*d*5  
Phenol-*d*6  
Terphenyl-*d*14

Neat oil samples were initially diluted by a factor of 25 with methylene chloride prior to GC-MS analysis. Produced water samples were acidified to pH 2 prior to extraction of WSOs. Three 30-mL contacts with methylene chloride were then used to serially extract organic compounds from 200 to 1000 mL of produced water in a separatory funnel. Internal standards were also added after solvent extraction as a means of checking the GC-MS instrument sensitivity and linearity. These compounds are listed in Table 6.3.

**Table 6.3. Internal standards for SVO compound analysis**

1,4-Dichlorobenzene  
Naphthalene-*d*8  
Acenaphthene-*d*10  
Phenanthrene-*d*10  
Chrysene-*d*12  
Perylene-*d*12

The combined extracts were dried with sodium sulfate and concentrated to 1 mL. A laboratory control spike standard provided an additional quality control check. The spike solution containing the compounds listed in Table 6.4, was added to an aliquot of the concentrated methylene chloride extract to check for the interference of sample matrix components with the quantitation of individual target compounds.

**Table 6.4. Compounds included in the laboratory control spike standard**

Acenaphthene	N-Nitroso-di-N-propylamine
4-Chloro-3-methylphenol	Pentachlorophenol
2-Chlorophenol	Phenol
1,4-Dichlorobenzene	Pyrene
2,4-Dinitrotoluene	4-Nitrophenol
1,2,4-Trichlorobenzene	

Prepared methylene chloride extracts were then injected into a GC-MS system (operated under the conditions listed in Table 6.5) to identify and **quantitate** individual **SVOCs**.

**Table 6.5. GC-MS conditions for SVOC analysis**

Instrumentation:	Hewlett Packard (HP) 6890 plus Gas Chromatograph with a 5973 MSD
GC column:	30 m X 0.25 mm (or 0.32 mm) ID 1- $\mu$ m-film-thickness silicone-coated fused-silica capillary column
Software:	HP MSD ChemStation with <b>EnviroQuant</b>
Ionization:	Electron impact, 70 eV
Ion source temperature:	230°C
Carrier gas	Helium, 30 cm/s
Initial pressure:	16 psi
Inlet:	Pulsed splitless, initial temperature 270 °C
Oven program:	50 °C (4 min), 10 °C/min to 350 °C, hold until <b>dibenz(a,e)pyrene</b> elutes
Sample Volume:	1 $\mu$ L

## 6.2 GC/FID ANALYSIS OF TOTAL PETROLEUM HYDROCARBONS

A modification of SW-846 Method 8015B was used to determine the concentrations of nonhalogenated volatile organic compounds and semivolatile organic compounds by GC using flame ionization detection (FID). This method analyzes total petroleum hydrocarbons (TPH) in specific carbon ranges defined as

- TPH-Gasoline (GRO): C<sub>6</sub>–C<sub>10</sub> ,
- TPH-Diesel (DRO): C<sub>10</sub>–C<sub>20</sub> , and
- TPH-Oil: C<sub>20</sub>–C<sub>28</sub> .

TPH-GRO represent alkanes having a boiling point range of 60–170°C; the remaining compounds correspond to alkanes having a boiling point range of 170–430°C. Prior to solvent extraction, the surrogate recovery standards (n-dodecane-d<sub>26</sub>, **naphthalene-d<sub>8</sub>**, and n-phenol-d<sub>6</sub> or tribromophenol) were added at a **level** of 40  $\mu$ g each to produced water in order to estimate the variability in extraction efficiency of alkanes, aromatic, and polar compounds, respectively. The concentrated solvent extracts containing WSO were then injected into a gas **chromatograph** using the operating conditions summarized in Table 6.6.

**Table 6.6. GC/FID operating conditions for TPH analysis**

Instrumentation:	HP 5890 Series II Gas Chromatograph with an <b>FID</b>
GC column:	30 m X 0.53 mm ID fused silica capillary column bonded with 5% methyl silicone ( <b>HP-5</b> ), <b>0.88-<math>\mu</math>m</b> film thickness
Software:	HP ChemStation version 3365 Series II
Carrier gas flow rate:	Helium, 4 mL/min
Makeup gas flow rate:	Helium, 26 mL/min
Injector temperature:	200°C
FID temperature:	340°C
Hydrogen flow rate:	30 mL/min
Air flow rate:	375 mL/min

**Table 6.6. GC/FID operating conditions for TPH analysis (cont.)**

<b>Oven program</b>	
Initial temperature:	45°C (hold 3 mm)
Program:	45°C to 275°C at 12°C/min
Final temperature:	Hold 12 mm
Sample size:	0.5-3 µL

Detector response was externally standardized using a six-point calibration curve. The calibration data set was constructed from dilutions of a certified standard (Absolute Standards, catalog no. 908 14, lot 030700) containing 2000 µg/mL each of n-alkanes spanning the range C<sub>6</sub>-C<sub>28</sub>. The FID response was linear over the range of 0.01 – 0.1 µg injected mass of each alkane, corresponding to a TPH range of 0.2-1.4 µg/mL for each alkane in produced water. HP Chem Station® software was used to integrate peak area over the three carbon ranges: C<sub>6</sub>-C<sub>10</sub>, C<sub>10</sub>-C<sub>20</sub>, and C<sub>20</sub>-C<sub>28</sub> to obtain the TPH mass in each of the solvent extraction concentrates.

### **6.3 INFRARED ANALYSIS OF-TOTAL RECOVERABLE PETROCHEMICAL HYDROCARBONS**

The analysis of WSO in produced water samples was also supplemented using USEPA Method 418.1. The infrared (IR) method is more commonly used in the field or as a screening technique due to the simplicity of the procedure. As modified for this task, the protocol uses 10 µL of 4 M HCl to acidify 2 mL of produced water to a value less than pH 2. The oil-and-grease components are then extracted from the treated sample with 3 mL of IR-grade tetrachloroethylene. The organic phase is subsequently transferred to a 1-cm path length IR cell. The absorbance at a wavelength of 3.4 µm is measured on a single-beam IR spectrometer. Sets of oil-and-grease standards are prepared in the working range of 4 to 40 mg in tetrachloroethylene using either GOM crude oil or a synthetic oil [37.5 (v/v) % n-hexadecane, 37.5 (v/v) % isooctane, and 25 (v/v) % chlorobenzene]. A three milliliter volume of oil standard is acidified; 2 mL of 100 g/L NaCl is then added to match the salinity content of GOM brine simulant. After the shaking and phase separation steps, the solvent phase is withdrawn as the working oil-and-grease standard.

### **6.4 FRACTIONATION OF WSO IN SIMULATED PRODUCED WATER**

A more definitive approach to characterizing the WSO content in produced water samples is to fractionate compounds into primary classes of saturated hydrocarbon, aromatic hydrocarbon, and polar components. The TPH content of each fraction is then determined by GC/FID analysis. An open liquid chromatographic (LC) column was used to perform the fractionation of extractable TPH using a procedure similar to that described by Mills et al. In this procedure, a gross measurement of organic content is made by extracting total TPH from an aqueous sample with methylene chloride. The extract is then applied to a multi-sorbent column to separate the chemical fractions. The advantage of such an analytical approach is that the chemical character of WSO can be determined inexpensively as a function of general classes for a sample containing a myriad of organic constituents.

The procedure of Mills et al. was modified to accommodate the analytical requirements of the PERF 98-04 project. A primary change in the protocol was to replace the gravimetric analysis of column fractions with that of GC/FID analysis based on USEPA SW-846 Method 8015B. Gravimetric analysis did not provide the necessary detection limits and compositional information considered relevant to the project. A secondary procedural change was to estimate polar hydrocarbon content by eluting the bulk of the material from the open LC column with a

final methanol wash. The distribution of carbon content within molecular size ranges could then be estimated for this fraction by **GC/FID** analysis. The **GC/FID** data for the polar fraction were then compared the method by Mills et al. The organic mass of the polar fraction can also be estimated by subtracting the combined mass of saturated and aromatic fractions from the total TPH content originally present in the methylene chloride extract. Finally, the solvent exchange procedure used to remove methylene chloride and transfer total TPH to a hexane matrix was altered midway through the study to minimize the loss of more volatile WSO components. Rather than evaporation at **65°C** in Kudema-Danish glassware, the methylene chloride solvent was heated to **35°C** under a light flow of nitrogen gas to reduce the sample volume. Two sequential additions of hexane were then added and concentrated in the same manner.

An outline of the modified procedure is reproduced in diagrammatic form in Fig. 6.1. Prior to TPH extraction, aliquots of produced water were withdrawn for analyses of inorganic constituents and low molecular weight organic acids. Surrogate recovery standards (*n*-hexanoic acid, *n*-dodecane-*d*26, *naphthalene-d*8, and *n*-phenol-*d*6 or tribromophenol) are then added to 1 L of produced water. Finally, WSOs were then serially extracted with three **30-mL** methylene chloride contacts.

The methylene chloride solvent was concentrated to **2 mL** at **65°C** using Kudema-Danish (K-D) evaporation glassware to yield the total extractable material (TEM) fraction. A **100-μL** aliquot of the TEM fraction was dried on a tared glass-fiber filter to estimate the **initial TPH** mass in the fraction gravimetrically (generally a few milligrams). A **second 100-μL** aliquot of the TEM fraction was placed in a sealed ampoule for subsequent **GC/FID** analysis.

In the first 14 oil/water contact experiments, **10 mL** of hexane was added to **1-mL** TEM aliquot and evaporated at **65°C** to a final volume of **2 mL** in order to transfer TPH into a hexane matrix (HEX fraction). This step was repeated to ensure that all methylene chloride had been removed from the sample. Beginning with experiment PERF-15, solvent substitution was modified to reduce organic losses at elevated temperatures. Instead, **2 mL** of hexane was added to a **1-mL** TEM fraction and evaporated to near dryness at **35°C** under a light flow of nitrogen gas. This step was repeated twice for complete solvent substitution. The residue was then brought to a **2-mL** volume with hexane. The hexane fraction, denoted as HEX, was refrigerated overnight to precipitate asphaltenes. A **100-μL** aliquot of the HEX fraction was dried on a tared glass-fiber filter to estimate the initial TPH mass in the fraction gravimetrically (generally a few milligrams). A **second 100-μL** aliquot of the HEX fraction was placed in a sealed ampoule for subsequent **GC/FID** analysis.

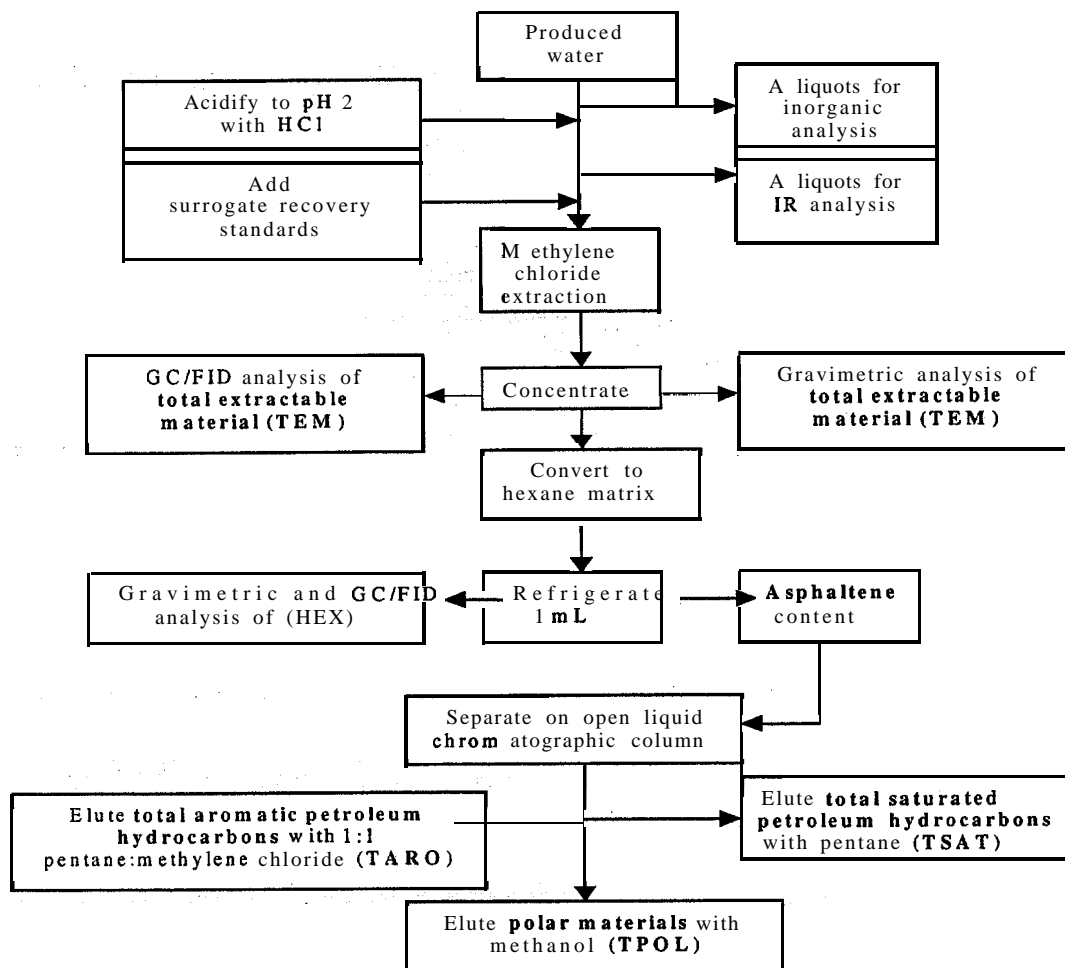
The remainder of the clarified HEX fraction was added to the open LC column. The multilayered column was packed sequentially with 0.5 cm of fired glass wool, 5 g of fired sand prewashed with methanol, 7 g of fired activated alumina, 13.5 g of fired activated silica, 2 g of anhydrous sodium sulfate, and a final top layer of 0.5 cm of fired glass wool (Fig. 6.2). The packed column was prewashed with **50 mL** of methanol, **200 mL** of methylene chloride, and finally **40 mL** of pentane before the HEX fraction was added. The saturated hydrocarbons from the applied HEX fraction were eluted with **40 mL** of pentane. This fraction was concentrated to **1 mL** by evaporation at **65°C** in the K-D glassware, and is denoted as the TSAT fraction. The concentrated TSAT fraction was transferred to a sealed ampoule for **GC/FID** analysis.

A **200 mL** volume of 1: 1 pentane:methylene chloride was used to elute the aromatic compounds from the LC column. This fraction, denoted as TARO, was also concentrated to **1 mL** by



evaporation at 65°C. The concentrated **TARO** fraction was transferred to a sealed vial for GC/FID analysis.

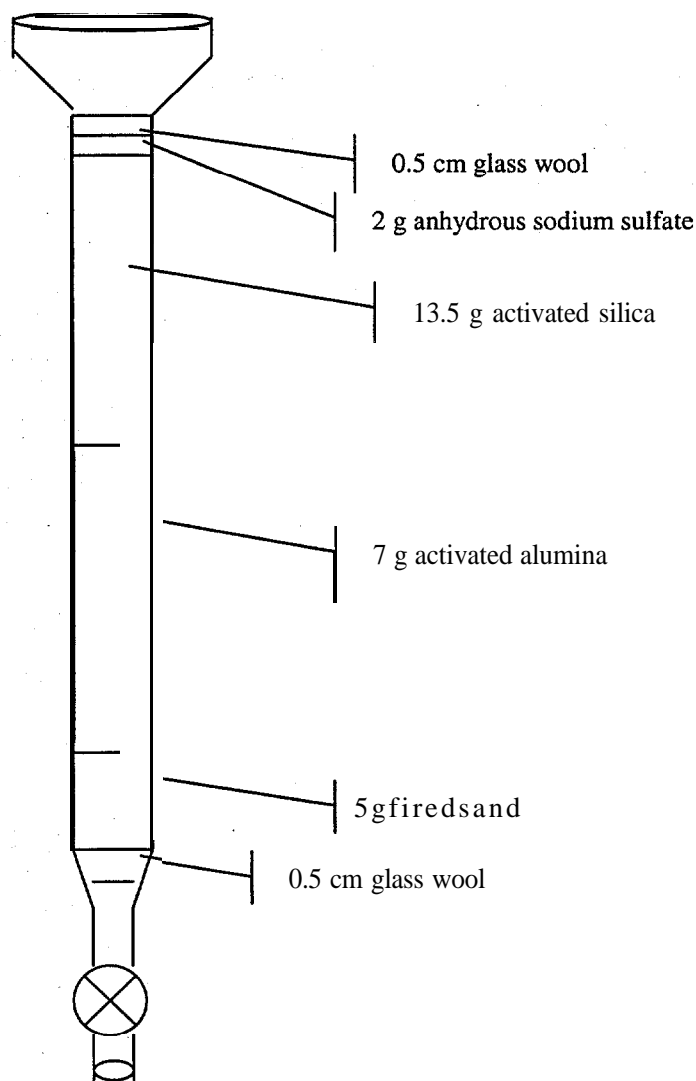
The LC column was finally washed with 150-mL of methanol to displace the majority of the polar compounds. This fraction, denoted as the **TPOL**, was then evaporated to 1 mL at 85°C. Each LC fraction was analyzed for organic mass content by GC/FID.



Total petroleum hydrocarbons (TPH) = Total saturated (TSAT) + total aromatic (TARO) + total polar (TPOL) fractions  
 True TPH = Total saturated (TSAT) + total aromatic (TARO) fractions

**Fig. 6.1. Fractionation of WSO in simulated produced water.**

Thus, the produced water sample is characterized by total TPH content as determined from the TEM fraction. Again, it is important to recognize that total TPH mass in actuality is a summation of real petroleum hydrocarbons and water soluble organic constituents that are also extracted using methylene chloride solvent. The HEX fraction contains total TPH that was transferred from the methylene chloride matrix to the non-polar hexane fraction. Actual TPH is a combination of saturated hydrocarbon mass found in the TSAT fraction and the aromatic hydrocarbon mass found in the **TARO** fraction. The **TPOL** fraction contains the balance of organic mass from the produced water that is not petroleum hydrocarbons, but was coextracted when the water was extracted with methylene chloride.



**Fig. 6.2. Open liquid chromatography column for fractionation of TPH .**

#### Column Preparation and Sample Elution

1. Pack the sorbents using a methanol solvent.
2. Once packed, elute the column successively with 50 mL of methanol, 200 mL of methylene chloride, and 40 mL of pentane.
3. Add the HEX sample to the top of the column. Rinse with several 1 mL pentane additions to transfer the sample to the head of the column.
4. Elute the TSAT fraction with 40 mL of pentane at 1 mL/min.
5. Rinse the K-D tube with 1: 1 methylene chloride:pentane and transfer rinsates to the head of the column. Elute the **TARO** fraction with 200 mL 1: 1 methylene chloride :pentane.
6. Rinse the K-D tube with methanol and transfer the rinsates to the head of the column. Elute the TPOL fraction with 150 mL of methanol.

## 6.5 DETERMINATION OF WATER CONTENT IN CRUDE OIL

The water content of as-received crude oil was determined by the standard Karl Fischer titration, using a Brinkman Model 652 KF-Coulometer. In this procedure, water reacts with iodine in the presence of an alkyl sulfurous acid and imidazole base to form the corresponding alkyl sulfuric acid. Iodine is generated at a constant rate by anodic oxidation of iodide present in the titration medium. When all the water present in the sample has reacted, excess iodine is indicated coulometrically. The total current multiplied by the time required to reach the titration end point is proportional to the water content of the sample. The anode compartment of the coulometric cell was filled with AQUASTAR® Coulomat A (EM Science) anode generator solution. The cathode compartment was filled to a level of 2–3 mm below that of the anode generator solution with AQUASTAR® Coulomat C (EM Science) counter solution. The instrument was then turned on, and the cell matrix was conditioned to remove the trace water present in cell reagents. When the instrument response no longer drifted, 10–50 µL of diluted oil was introduced into the coulometric cell. Equivalent water introduced to the cell from the diluted oil samples was 100–300 µg. Instrument response was calibrated using a 1 mg/mL water standard (HYDRANAL® Water Standard 1.00, Allied Signal).

## 6.6 DETERMINATION OF pH AND BICARBONATE/CARBONATE CONTENT IN WATER

A Metrohm 717 DMS Titrino automatic titrator was used to determine the pH and the hydroxide/bicarbonate/carbonate content of produced water samples (Franson, 1992). The instrument was standardized with two NIST-traceable buffer solutions (pH 7 and pH 10, respectively). The temperature of the solution was entered digitally before the pH of the sample was measured. The OH<sup>-</sup>, CO<sub>3</sub><sup>2-</sup>, and HCO<sub>3</sub><sup>-</sup> concentrations were determined by titrating 20 mL of produced water with standard 0.01 N HCl. End points measured at pH 8.3 (i.e., volume end point A) and 3.7 (i.e., volume end point B) were then entered into the following calculations:

If  $2A > B$ , the solution contains OH<sup>-</sup> and CO<sub>3</sub><sup>2-</sup>,

$$(2A - B) \times \text{normality of HCl} \times 17.0073 / (\text{sample volume}) = \text{ppm OH}^-$$

$$2(B - A) \times \text{normality of HCl} \times 30.0046 / (\text{sample volume, mL}) = \text{ppm CO}_3^{2-}$$

If  $B > 2A$ , the solution contains CO<sub>3</sub><sup>2-</sup> and HCO<sub>3</sub><sup>-</sup>,

$$2A \times \text{normality of HCl} \times 30.0046 / (\text{sample volume, mL}) = \text{ppm CO}_3^{2-}$$

$$(B - 2A) \times \text{normality of HCl} \times 61.0171 / (\text{sample volume, mL}) = \text{ppm HCO}_3^-$$

Total alkalinity is determined from the total volume of acid required to achieve a pH of 3.7. It is calculated as:

$$\frac{(A + B) \times \text{normality of HCl} \times 1000}{\text{sample volume, mL}} = \text{Alkalinity to pH 3.7, mg CaCO}_3/\text{L}$$

## 6.7 DETERMINATION OF CHLORIDE CONTENT IN PRODUCED WATER

Ion chromatography (IC) and chloride ion-selective electrodes (ISE) were used to determine the salinity in produced water. Performance of IC protocol followed that of USEPA Method 9056 “Determination of Inorganic Anions in Ion Chromatography”. Chloride analysis was performed by the sequential elution of 0.01 to 50 ppm of anions from an IonPac AJ11 analytical column using a gradient elution. The eluting anion concentration was detected with a temperature-compensated electrical conductivity cell. The instrument was calibrated by the introduction of standard anion solutions using IC operating conditions summarized in Table 6.7.

**Table 6.7 IC operating conditions for anion analysis**

Sample loop volume:	4mmID; 10 $\mu$ L
Trap column:	ATC-1
Guard column:	IonPac <sup>®</sup> AG11
Analytical Column:	IonPac <sup>®</sup> AJ11
Eluents:	El: Type I deionized water E2: 5.0 mM NaOH E3: 100 mM NaOH
Eluent flow rate:	2.0 mL/min
Suppressor:	Anion self-regenerating suppressor Autosuppression recycle mode

The majority of the sodium in brine samples had to be removed prior to anion analysis by IC. To accomplish this, a portion of the sample was passed through a hydrogen-form cation-exchange resin. The acidified sample was then introduced to the IC.

Determination of the chloride concentration was more commonly achieved using an Orion Model 94-17B chloride electrode and an Orion Model 90-02 Double-Junction Reference Electrode in combination with an Orion 520A pH/ISE meter. Response of the electrode system was calibrated using NaCl standards (100-3000 ppm Cl<sup>-</sup>) in which an Ionic Adjustor Solution (ISA) of 5 M NaNO<sub>3</sub> was added to modify the ionic strength of the solution. ISA was similarly added to a 1: 100 dilution of produced water to measure the chloride content.

## 6.8 DETERMINATION OF METAL CONTENT IN PRODUCED WATER

ICP-AES was used to determine the trace-element content of the GOM brine simulant and the produced water. All of the brine simulant samples were acidified to 10% HNO<sub>3</sub> concentration prior to analysis. In the case of both actual and simulated produced water samples, the dissolved organic constituents were first wet-ashed using microwave-assisted acid digestion (USEPA Method 3015). In this preparation technique, of 50 mL sample was digested in 5 mL concentrated HNO<sub>3</sub> in a Teflon® digestion vessel for 20 min using microwave heating. After the sample has cooled, it was filtered prior to analysis.

A Thermo-Jarrel Ash 61E ICP-AES system was used to measure trace metal concentration according to USEPA Method 6010B. Table 6.8 shows an example of instrument output for the analysis of GOM brine simulant, which lists the metals that are quantified by the equipment, as well as the instrument detection levels. Instrumental response and elemental emission correction factors were established by analyzing blanks and certified mixed-element standards.

**Table 6.8. ICP inorganic analysis of GOM brine simulant using ICP-AES**

<b>Metal</b>	<b>IDL (mg/L)</b>	<b>Quantification limit (mg/L)</b>	<b>Raw conc. (mg/L)</b>	<b>data S<sub>x</sub> (mg/L)</b>	<b>Sample conc (mg/L)</b>	<b>Analytical error (mg/L)</b>
Silver	0.002744	0.008	-0.00086	0.00035	8.00	± 0.87
Aluminum	0.046404	0.06	-0.01643	0.01023	60.0	± 25.4
Arsenic	0.008304	0.02	-0.00649	0.00092	20.0	± 2.3
Boron	0.021937	0.05	-0.03 163	0.00094	50.0	± 2.3
Barium	9.26842E	0.003	0.04387	0.00125	43.9	± 3.1
Beryllium	5.32885E	0.002	-0.00046	0.00013	2.00	± 0.32
Calcium	0.025085	0.075	3.3131	0.07225	3313,	± 179
Cadmium	6.34906E	0.002	-0.00004	0.0001	2.00	± 0.25
Chromium	0.00 1673	0.005	-0.00051	0.00015	5.00	± 0.37
Cesium <sup>a</sup>	17.10689	20	-0.61237	3.15486	20,000	± 7838
Copper	0.005883	0.02	-0.00186	0.00076	20.0	± 1.9
Iron	0.113441	0.2	0.04355	0.01663	200	± 41
Potassium	0.023321	0.04	-0.06447	0.00535	40	± 13
Magnesium	0.040085	0.05	0.883 16	0.02262	883	± 56
Manganese	0.001003	0.003	-0.00038	0.00007	3.00	± 0.17
Molybdenum	0.00407 1	0.01	-0.00042	0.00078	10.0	± 1.9
Sodium	0.092114	0.15	35.14379	0.84618	35,144	± 2102
Niobium	0.00273 1	0.008	0.0018	0.00424	8.0	± 10.5
Nickel	0.006758	0.02	-0.00475	0.00034	20.0	± 0.8
Lead	0.002946	0.009	-0.00114	0.00136	9.0	± 3.4
Selenium	0.010864	0.03	-0.00189	0.00169	30.0	± 4.2
Antimony	0.017710	0.04	-0.00373	0.00158	40.0	± 3.9
Silicon	0.029039	0.06	-0.0279	0.00489	60.0	± 12.1
Strontium	3.90035E	0.001	0.13402	0.00322	134.0	± 8.0
Thorium	0.553 176	0.8	0.01249	0.09979	800	± 248
Titanium	6.18243E	0.002	-0.0013	0.00002	2.00	± 0.05
Thallium	0.008663	0.025	0.0002	0.00337	25.0	± 8.4
Uranium	0.073958	0.18	0.12958	0.06027	180	± 150
Vanadium	0.007891	0.02	-0.00026	0.00013	20.00	± 0.32
Zinc	0.004324	0.01	-0.00797	0.00016	10.00	± 0.40
Zirconium	0.031518	0.07	0.00154	0.00561	7.0 0 0	± 13.9

<sup>a</sup> Cesium concentration overestimated due to the interference of large quantities of sodium ion in the sample.

## 6.9 ANALYSIS OF C<sub>1</sub>–C<sub>6</sub> ORGANIC ACIDS

Concentrations of organic acids were determined by injecting produced water directly into a Dionex LC20 ion chromatograph fitted with an IonPac<sup>®</sup> ICE-AS6 (Dionex Corp.) separation column. The acids were separated from the high chloride sample matrix using a 0.4 *mM* heptafluorobutyric acid eluent; a 1-ppm detection limit was typical. Instrument response was calibrated using formic, acetic, propionic, malonic, and oxalic acids in the concentration range of 10–50 ppm. Specific operating conditions are listed below.

**Table 6.9 IC operating conditions for C<sub>1</sub>–C<sub>6</sub> organic acid analysis**

Sample loop volume:	0.037-mm ID X 9.75 in length; 50 µL
Analytical Column:	IonPac <sup>®</sup> ICE-AS6
Eluents:	E1: 1.8 <i>mM</i> carbonate/1.7 <i>mM</i> bicarbonate E2: 0.1 <i>M</i> sodium hydroxide E3: 0.4 <i>mM</i> heptafluorobutyric acid E4: dionized water
Eluent flow rate:	1.0mL/min
Suppressor:	Anion micromembrane self-regenerating suppressor-ICE II Autosuppression regenerant mode

## 7. ASSESSMENT OF DATA QUALITY

### 7.1 CORRESPONDENCE BETWEEN ANALYTICAL METHODS FOR WSO CONSTITUENTS

Correspondence testing of chemical procedures used to analyze the chemical contents of neat oil and produced water was performed to establish equivalency in data derived through various analytical protocols within the PERF 98-04 project. Statoil provided characterization data derived from North Sea samples. Their protocol relied on the analysis of individual semi-volatile compounds, as defined by USEPA Method SW-846 8270C, and total petroleum hydrocarbon content, as defined by USEPA Method SW-846 8015B. ORNL supplied data derived from GOM crude oil and brine. Rather than identifying individual chemical compounds, ORNL characterized samples based on carbon-range content and general chemical classes. It will be important to understand the equivalency of the North Sea and GOM data when all the information is combined into a single mathematical model that will predict organic contamination in produced water.

ORNL submitted samples of GOM crude and synthetic produced water to Mountain States Analytical, Inc. (1645 West 2200 South, Salt Lake City, Utah 84119) for standard USEPA Methods testing. These samples were also analyzed at ORNL, using open-column LC separation to fractionate WSO into general chemical classes. The resulting fractions were then analyzed by GC/FID to determine the relative distribution of carbon content in each of these classes. Comparative results of deep-well crude by various protocols are summarized in Table 7.1. The TPH content in the GOM crude, as defined by USEPA Method SW-846 8015B, was 300 g/kg TPH-DRO C<sub>10</sub>–C<sub>28</sub>; 180 g/kg TPH-GRO C<sub>6</sub>–C<sub>10</sub>; and 200 g/kg TPH-Oil C<sub>20</sub>–C<sub>28</sub>. Approximate ratios of carbon size by the USEPA method were 37, 21, and 42% for C<sub>6</sub>–C<sub>10</sub>, C<sub>10</sub>–C<sub>20</sub>, and C<sub>20</sub>–C<sub>28</sub> ranges, respectively. The relative distribution using a similar GC/FID procedure at

ORNL was found to be 50, 42, and 8% in the carbon series. Differences in relative distribution of carbon content may reflect error introduced by variances in sampling time and frequency of sampling. Off-site data were the result of duplicate analyses from a single grab sample taken on arrival of Deep1 crude. ORNL data were derived from the average results of five grab samples taken over a several month time frame.

The principal semi-volatile compounds in the deep-well oil, as identified by USEPA Method SW-846 8270C, included acetophenone (0.6 g/kg) and methylnaphthalene (0.4 g/kg). These particular constituents represented less than 2% of the total organic content in either GOM or North Sea sources of crude oil. Chemical classification by open LC protocol helps to account for the remaining 98% of the organic character in either oil or produced water samples. Methylene chloride-extractable material (TEM) in deep-well oil was 200–400 g/kg. Approximately 35% of this material could be exchanged into a hexane “matrix. The relative composition of the hexane matrix is 35% saturated hydrocarbons, 15% aromatic components, and 50% polar material. The visibly colored red-brown material was primarily present in the polar fraction.

**Table 7.1 Correspondence of analytical methods in the characterization of deep-well crude**

	Deep-well oil analysis (g/kg)			
	Offsite	Offsite duplicate	ORNL	ORNL duplicate
USEPA Method SW-846 8015B: , ,				
Carbon analysis in total extractable material:				
TPH-DRO C <sub>10</sub> –C <sub>28</sub>	314	310	192	219
TPH-GRO	179	179		
TPH-OIL	192	186		
TPH: C <sub>6</sub> –C <sub>10</sub>			98	120
TPH: C <sub>10</sub> –C <sub>20</sub>			82	80
TPH: C <sub>20</sub> –C <sub>28</sub>			12	19
USEPA Method SW-846 8270C				
Semivolatile organics (SVO):				
Acetophenone	0.60	0.58		
Methylnaphthalene	0.40	0.43		
Open LC analysis (ORNL):				
Total extractable material (TEM)			193	220
Total hexane extractable (HEX)/% TEM			50/ 26	88/ 40
Total saturated (TSAT) / %TEM			70 / 36	84/ 38
Total aromatic (TARO) / %TEM			36 / 19	20/ 9
Total polar (TPOL) / %TEM			88 / 46	116/ 53

Note: Total TEM = TSAT +TARO +TPOL

The total concentration of WSO in produced water derived from oil/brine simulant contacts was 20-30 ppm, using either USEPA Method SW-846 8015B or ORNL GC/FID procedures. Identified semi-volatile compounds were 1-methylnaphthalene (10 ppb), 2-methylnaphthalene (9 ppb), naphthalene (14 ppb), and phenol (30 ppb). Again, compounds identified by USEPA Method SW-846 (8270C) accounted for very little of the total WSO content. Chemical

fractionation of the total extractable materials suggested that **80–90%** of WSO is present as polar compounds; the next largest fraction is that of aromatic materials. The colored material was primarily present in the lightly yellow aromatic fraction.

In combining data from the North Sea and GOM samples, it appears that data generated by USEPA Method SW-846 8015B or ORNL GC/FID procedures correlate well to indicate the total organic content and relative carbon ranges from these two drilling sites. The toxicity of the WSO can be estimated using data generated by USEPA Method SW-846 8270C from North Sea samples and the chemical nature of the material can be estimated using the open LC method derived from GOM samples.

## **7.2 SURROGATE RECOVERY AND BLANK LEVELS FOR OPEN LC FRACTIONATION AND GC/FID ANALYSIS**

The analysis of produced water generated by contacting shipped oil samples with simulant GOM brine probably reflects a negative bias in total WSO content. The oil had been contacted with injection water during its collection **onsite**; therefore, a portion of the TPH had already been stripped from the sample of crude prior to testing in this project. An estimate of the analytical bias might be achieved by analyzing WSO in actual produced water from this particular site.

The gravimetric analysis of evaporate residue provided a gross estimate of TPH content. This analysis was usually performed on either crude oil dissolved directly in methylene chloride or WSO present in the concentrated methylene chloride extracts (TEM fractions) of produced water. The lower limit of detection for gravimetric analysis was dictated by the **0.2-mg** weight limit of an electronic balance. This measurement corresponded to a gravimetric limit of detection of 0.2 g/L in 650 g/L WSO found in crude. The WSO content in produced water was significantly lower (20-30 ppm). Since the TEM fraction is derived from the extraction of one liter of produced water, the balance weight limit corresponded to a gravimetric limit of detection of 12 ppm in aqueous samples.

GC/FID provided a much more sensitive estimate for WSO present in oil or produced water. Quantitation was based on instrument response relative to that of a series of certified **n-alkane** standards. Identification of mass within carbon size ranges is dependent on retention times of the various **alkane** standard components (Fig. 7.1). Both **C<sub>6</sub>–C<sub>10</sub>** and **C<sub>10</sub>–C<sub>20</sub>** ranges had similar FID response factors; FID response over the **C<sub>20</sub>–C<sub>28</sub>** range was slightly lower due to column bleed at the elevated GC temperatures in this portion of the chromatogram. The relatively constant FID response factor over all carbon ranges indicated that GC/FID analysis provided a relatively accurate estimate of true TPH mass in the LC fractions. The material present in the TPOL fraction is made up of compounds containing nitrogen, oxygen and sulfur atoms. FID response to non-hydrocarbon components is more variable. Therefore, GC/FID analysis of this fraction in particular provides an estimate of organic mass.

The successive solvent extractions and multiple concentration steps in either the USEPA protocol or the LC fraction protocol contributed to losses of WSO material. Although negative bias is inherent to the WSO analysis, discharge permitting is based on the same USEPA method. Therefore, data derived from such procedures will still reflect a value for WSO content that can be used for modeling and subsequent development of engineering practices that will reduce the generation of WSO. The extraction recovery of WSO were typically estimated by the introduction of surrogate recovery standards to the produced water prior to manipulating the sample. Deuterated forms of the spike standards were selected to provide the option for future



sample analysis by GC-MS if the latter was considered necessary. These compounds are listed in Table 7.2, together with the recovery found for each LC fraction. Tribromophenol was substituted for phenol midway through the experimental campaign. Phenol identification was difficult due to the proximity of its retention peak to the peaks of the majority of WSO compounds found in the TEM and HEX fractions.

**Table 7.2. Surrogate spike recovery compounds for LC fractionation protocol**

Recovery standard	Constituent surrogate	Spike level ( $\mu\text{g}$ )	Recovery in each LC fraction (%)				
			TEM	HEX	TSAT	TARO	TPOL
n-Hexanoic acid	Organic-acids	4,635	100	55	0	0	1-10
<i>n</i> -Dodecane- <i>d</i> <sub>26</sub>	Saturated TPH	115	100	60	5-30	0	0
Naphthalene- <i>d</i> <sub>8</sub>	Aromatic TPH	100	100	55	0	20-60	5-40
N-phenol& tribromophenol	Polar TPH	70	---	---	0	0	0-40
	polar TPH	196	100	75	0	0	50-90

Positive bias can be introduced into the analytical procedure through the contamination present in solvents and labware. Therefore, a procedural blank containing all surrogate recovery standards was run to determine the WSO blank level for each LC fraction. Figure 7.2 presents a chromatogram of the blank levels and recovery standards associated with the TEM fraction derived from the methylene chloride extraction of 1 L simulant brine spiked with recovery standards. The solvent front completely eluted after 2 min. After solvent and recovery surrogate data was stripped from the chromatogram, regression analysis of the resulting peak area for this fraction yielded a calculated blank level of -0.2  $\mu\text{g}$  TPH. This level compares with a 0.2- $\mu\text{g}$  TPH detection limit for the GC/FID. Complete processing through the LC protocol yielded the blank levels shown in Table 7.3. Comparable WSO blank levels were calculated for 1 L of produced water. A positive blank was associated with the TPOL fraction: A low level of apparent WSO resulted from methanol extraction of the organic matter from fired sand. This necessitated an additional wash of fired sand before a packed LC column was loaded with sample.

**Table 7.3 Calculated blank levels associated with each LC fraction**

LC fraction	TPH ( $\mu\text{g}$ )	Equivalent WSO in produced water (ppm)
TEM	-0.2	-0.02
HEX	-0.3	-0.02
TSAT		0.0
TARO	0.08	0.08
TPOL	0.2	0.02

### 7.3 BIASES ASSOCIATED WITH LC FRACTIONATION PROTOCOL

A series of chromatograms were prepared to illustrate the typical quantities of TPH found in individual LC fractions, as well as the relative losses encountered with the fractionation protocol. These chromatograms followed the processing of produced water generated from contacting 750 mL of deep-well crude with 3000 mL of GOM brine simulant at pH 7, 65,000 TDS, 25°C, and ambient pressure (denoted as experiment "PERF-20" in appendix data tables). As in chromatograms of diluted oils, the WSO in produced water contained a multitude of organic components. Figure 7.3 presents the WSO contained in the concentrated methylene chloride extract of produced water (i.e., the TEM fraction). Because of the complexity of the composition,

it is evident that a general classification, rather than specific component identification, should be used to characterize WSO in produced water economically. The majority of WSO elute at retention times comparable to  $C_6$ – $C_{10}$  and  $C_{10}$ – $C_{20}$  n-alkanes. Because of the quantity of material present in TEM fractions, manual baseline construction was performed to determine the organic mass in these fractions.

Most of the TEM fraction concentrate was converted to a hexane (HEX) matrix prior to its addition to the LC column. This procedural step accomplished two objectives. The first was to acquire TPH data equivalent to data obtained from extractions with hexane, a typical solvent used to determine WSO in field testing. Second, the LC protocol required an initial nonpolar matrix. Organic fractionation was then accomplished by serially increasing the polarity of the eluent. Figure 7.4 presents the chromatogram of WSO in the hexane matrix. Concentration of hexane requires a heating temperature of 85°C, which is 20°C higher than that needed for methylene chloride evaporation. Therefore, some of the more volatile WSO constituents in the  $C_6$ – $C_{10}$  range were less prevalent in the chromatogram of HEX extracts, yielding a carbon range distribution observed using USEPA methods based on hexane extractions. Approximately 50% of the surrogate recovery standards are lost in the solvent exchange. Beginning with experiment PERF-15, solvent substitution was modified to reduce organic losses as a result of elevated temperatures. The volume of hexane was reduced by evaporation at 35°C under a flow of nitrogen gas. Again, the quantity of material present in TEM fraction was such that the chromatograms were typically not baseline resolved. This may have contributed to a slight negative bias in the calculation of percent TPH transferred to the HEX fraction.

Pentane was used to elute saturated organic compounds from the LC column. Typically, the TSAT chromatogram (Fig.7.5) contained only the n-decane surrogate standard at a recovery of approximately 25%. Aromatic compounds were present in the produced water as indicated in the number of peaks found in the chromatogram of the TARO fraction. Naphthalene was usually recovered at 40% of initial spike levels. The TARO chromatogram (Fig. 7.6) contained column contaminants in the  $C_6$ – $C_{10}$  carbon range equivalent to about 0.1 ppm WSO in produced water. This contaminant level was subtracted from all TARO fraction data.

The LC protocol developed by Mills et al. was used to calculate the quantity of polar constituents in produced water by subtracting the combined WSO mass found in TARO and TSAT fractions from that found in the TEM fraction. This type of calculation was made for each contact experiment performed in the study reported here. These values are denoted as TPOL(subtract) in the appendix data tables. This form of calculation may have overestimated the contribution of polar material, considering that only 40% of TEM material was actually transferred to the LC column in the form of the HEX extract. If it is assumed that losses in the solvent exchange from methylene chloride to hexane were primarily due to the evaporation of low molecular weight organic acids, then the TPOL(subtract) data might adequately compensate for this loss and thus would not significantly bias the calculation of total WSO found in this fraction.

The polar material in produced water was also estimated by adding a final elution step to the LC protocol defined by Mills et al. After the TARO fraction had been collected, a polar solvent was employed to elute the polar WSO from the column. Initially, methylene chloride was used as the eluent; however, methanol was found to be more effective in removing polar constituents and was used in experiment PERF-11 and all experiments thereafter. Even with a methanol solvent, the presence of colored material remaining on the fired sand at the head of the column from the LC separation of diluted crude indicated that some polar constituents were irreversibly adsorbed to the column. Nonetheless, the chromatogram (Fig. 7.7) of the eluted TPOL fraction, denoted as

TPOL(gc) in the appendix data tables, implies that most of the WSO compounds were present as polar constituents in the C<sub>10</sub>–C<sub>20</sub> carbon range. The hexanoic acid and tribromophenol recovery standards in this fraction were present at 5 and 70%, respectively.

#### 7.4 COMPARISON OF LC FRACTIONATION PROTOCOL WITH INFRARED SPECTROMETRIC ANALYSIS,

A secondary comparison of methods is presented below. The TPH content of a portion of produced-water samples was determined by the GC/FID analysis of TEM fractions and, secondarily, by the IR analysis of their tetrachloroethylene extracts. Results based on IR data tended to underestimated the TPH content; the salinity present in the produced water depressed IR absorbance readings in the samples (see Table 7.4).

**Table 7.4. Correspondence of TPH results from LC fractionation and IR spectroscopy**

Experiment	TPH (mg/L)	
	LC fractionation - TEM fraction	IR spectroscopy
PERF-6	12	1
PERP-7	37	
PERF-8	25	6
PERF-9	30	13
PERF-10	22	13
PERF-11	2	6 25
PERF-12	11	
PERF-13	21	16
PERF-14	10	49.8
PERF-15	12	12.3
PERP-16	14	10.5
PERP-17	18	17.6
PERF-18	2	1 23.1
PERP-19	10	8.7
PERF-20	21	10.6

## 8. RESULTS

### 8.1 CHEMICAL CHARACTERIZATION OF DEEP-WELL CRUDE

The LC column fractionation procedure was used to characterize deep-well crude according to the quantity of organic matter in each chemical class, as well as carbon size distribution. Results presented in Table 8.1 summarize the data for three to five deep crude analyses, in which 1g of mixed oil was diluted 25-fold with methylene chloride. Data were calculated in units of grams organic mass per liter of oil but can be converted to a weight basis using a density value of 0.8432 g/L at 25°C. The TEM fraction was made up of 45% each of C<sub>6</sub>–C<sub>10</sub> and C<sub>10</sub>–C<sub>20</sub> material. The C<sub>6</sub>–C<sub>28</sub> data indicate that 35% of the organic mass was transferred to the HEX fraction; the major loss of material was from the C<sub>6</sub>–C<sub>10</sub> carbon range. Of the organic material transferred to the HEX fraction roughly 75% of the mass was contained in the saturated hydrocarbon class. The remaining mass was equally divided between aromatic and polar constituents. The standard

deviation in the data reflects a combination of error introduced in sampling the mixed oil and in preparing the LC fractions.

**Table 8.1. Chemical fraction data for GOM deep-well crude samples**

LC fraction	C <sub>6</sub> -C <sub>28</sub> (g/L)	C <sub>6</sub> -C <sub>10</sub> (g/L)	C <sub>10</sub> -C <sub>20</sub> (g/L)	C <sub>20</sub> -C <sub>28</sub> (g/L)
TEM	201 ± 26	114 ± 21	90. ± 14	9.7 ± 6.1
HEX	70. ± 19	5.3 ± 3.5	54 ± 19	10. ± 4.7
TSAT	74 ± 9.1	8.3 ± 4.0	61 ± 11	3.5 ± 0.7
TARO	21 ± 14	4.0 ± 3.4	15 ± 9.6	4.0 ± 3.0
TPOL(gc)	16 ± 3.6	14 ± 4.2	0.57 ± 0.06	0.67 ± 0.57

## 8.2 WSO AS A FUNCTION OF PERCENT WATER CUT

Methylene chloride-extractable material in the equilibrated produced water was typically present at 20-30 ppm. Its chemical character differed significantly from that of the deep crude. The values listed in Table 8.2 reflect the average of three oil/brine contacts in which the experimental baseline conditions were set at 80% GOM synthetic brine/oil, equilibration for 4 days at 25°C, pH 7 brine, and 65,000 TDS. The brine volumes added to the WAF vessel varied from 1 to 3 L, while the oil volume varied from 0.25 to 0.75 L. TEM organic matter was distributed between the low and midrange carbon masses, similar to the oil data. However, the WSO eluted from the LC column are primarily polar compounds with a minor contribution by aromatic constituents. The slight yellow coloration of produced water appears to be associated with the aromatic fraction. As would be expected, saturated compounds were not noticeably extracted into the produced water.

**Table 8.2. Chemical fraction data for GOM synthetic produced water\***

LC Fraction	C <sub>6</sub> -C <sub>28</sub> (mg/L)	C <sub>6</sub> -C <sub>10</sub> (mg/L)	C <sub>10</sub> -C <sub>20</sub> (mg/L)	C <sub>20</sub> -C <sub>28</sub> (mg/L)
TEM	21 ± 4	7 ± 0.4	13 ± 3	0.8 ± 1.3
HEX	11 ± 5	0.4 ± 0.4	10 ± 4	0.01 ± 0.01
TSAT	0.02 ± 0.02	0.02 ± 0.01	0.00 ± 0.00	0.00 ± 0.00
TARO	0.2 ± 0.2	0.07 ± 0.08	0.04 ± 0.02	0.00 ± 0.00
TPOL(gc)	2.4 ± 1.7	0.00 ± 0.00	2.4 ± 1.7	0.00 ± 0.00

\*80% GOM synthetic brine/oil, 4 days at 25 °C, pH 7 brine, 65,000 TDS WAF vessel

Figures 8.1 and 8.2 summarize the data trends for various water/oil ratios. There was a slight negative trend in WSO content in all chemical fractions as the water cut increased. This effect was a combination of decreasing solubility of the C<sub>10</sub>-C<sub>20</sub> mass and increasing solubility of the C<sub>6</sub>-C<sub>10</sub> carbonaceous material.

## 8.3 WSO DISTRIBUTION COEFFICIENTS AS A FUNCTION OF PERCENT WATER CUT

The analytical results obtained in the characterization of crude oil were used to determine values of the approximate distribution coefficients ( $K_{ds}$ ) of WSO in GOM brine as a function of percent water cut and brine volume. The  $K_{ds}$  were calculated for brine simulant at baseline conditions of 65,000 TDS, pH 7 brine, ambient pressure, and 25°C. Equilibrium conditions were assumed after

a 4day contact time. Calculations based on the following equation were performed for each chemical class:

$$K_d = \frac{(\text{initial oil concentration}) \times (\text{oil volume}) - (\text{WSO in brine}) \times (\text{brine volume})}{(\text{WSO in brine}) \times (\text{oil volume})}$$

The data presented in Fig. 8.3 imply that the value for Log  $K_d$  of each chemical fraction remained constant with water cut. Average values for Log  $K_d$  were 4.0, 3.8, 5.0, 5.1, and 6.3 for TEM, HEX, TPOL, TARO and TSAT fractions, respectively. The value of Log  $K_d$  for each carbon range was also independent of water cut, as borne out in Fig. 8.4 as well as in the data summarized in Tables 8.3 and 8.4. The only deviation from this trend was observed with slightly higher values for the data set obtained for a 73% water, cut. These data were generated using the small pressure vessel. Although all data acquired with the pressure vessel were consistent in the subset, WSO values in the produced water were lower than those obtained with the WAF vessel.

#### 8.4 WSO AS A FUNCTION OF SALINITY

The variation of TDS (as salinity) did not significantly alter the WSO content of the produced water. As seen in Fig. 8.5, the total WSO remained at 10-15 ppm over the chloride concentration range of 40,000 to 115,000 ppm. Neither the chemical character nor the carbon content (Fig. 8.6) varied with TDS.

#### 8.5 WSO AS A FUNCTION OF pH

The pH of the GOM brine simulant. was-buffered to 7-7.5 by the presence of an oil layer. In order for the oil/water contact to be performed at a pH outside this range, NaOH or HCl had to be added continuously to maintain the desired pH level. The use of a 0.25 M borate buffer was not sufficient to maintain the pH of the aqueous phase at more alkaline values. The required pH adjustment was reduced with each contact day, such that no further adjustment was required on the fourth day. Thus, equilibrium was assumed to have been reached at this point. As expected, the pH of the brine simulant affected the concentrations of  $\text{CO}_3^{2-}$  and  $\text{HCO}_3^-$ . Acidic solutions reacted with  $\text{HCO}_3^-$  anion. Basic solutions increased the initial carbonate concentration sufficiently to precipitate alkaline earth metals, thereby reducing the equilibrium concentration of both  $\text{CO}_3^{2-}$  and  $\text{HCO}_3^-$ .

The WSO content almost doubled for pH values above 7 (Fig. 8.7). This was primarily due to the enhanced solubility of polar organic compounds that ionize at higher pH levels. An enhanced solubility of the  $\text{C}_{10}$ - $\text{C}_{20}$  carbon-range material and a decline in  $\text{C}_6$ - $\text{C}_{10}$  mass for pH values greater than 7 were also observed (Fig. 8.8).

#### 8.6 WSO AS A FUNCTION OF TEMPERATURE

The chemical character of the WSO did not appear to be dependent on brine temperature (Fig. 8.9). However, the percentage mass in each carbon range was altered; the mass of WSO in the  $\text{C}_6$ - $\text{C}_{10}$  carbon range decreased while that in the  $\text{C}_{10}$ - $\text{C}_{20}$  range increased (Fig. 8.10). Thus, the net quantity of organics in the produced water increased by only 5 ppm for a 60°C elevation in solution temperature.

## 8.7 WSO AS A FUNCTION OF PRESSURE

The distribution of organic compounds in GOM brine was studied as a function of pressure over the range of 1 to 58 bar. The results are presented in Fig. 8.11 and 8.12. It was found that WSO content in the TEM fraction increased by less than ten percent over the pressure range. The total WSO content for this subset of experiments were less than expected, based on previous data obtained with the WAF vessel. Experiments performed with the WAF vessel were typically conducted with 300 mL oil in contact with 1 L of brine. The total WSO levels were generally 16-20 ppm for WAF experiments. The WSO of the 200 mL brine drawn from the pressure vessel consistently averaged only 4-6 ppm. At this point, the reason for this variance is unknown. Pressure data presented in these graphs also exhibit a fair amount of scatter, attributed to both the low level of WSO in the pressurized samples and the difficulty in separating the phases after contact in the opaque pressure vessel.

The slight increase observed in total WSO as a function of pressure appears to be associated with the polar fraction. The influence of pressure on the measured fractions of saturated hydrocarbons (TSAT) and aromatic hydrocarbons (TARO) were investigated, but no trends were observed. The enhanced quantity of WSO observed in the polar fraction was the result of additional dissolved C<sub>6</sub>-C<sub>10</sub> material (Fig. 8.12). A slight decrease in quantities in the C<sub>20</sub>-C<sub>28</sub> range was also evident.

## 8.8 COMPOSITION OF C<sub>1</sub>-C<sub>6</sub> ORGANIC ACIDS IN PRODUCED WATER

Three organic acids were typically found in the produced-water samples: formic, acetic, and propionic acids. No higher-carbon acids were found. Acetic acid, which made up the bulk of the acid mass, was present at levels similar to that of TPH (10-30 ppm). Propionic and formic acids were present at approximately 60 and 20%, respectively, of the acetic acid concentration.

It is important to understand the extent to which low-molecular-weight acids might elevate the apparent TPH concentration found in methylene chloride and HEX extracts. The NPDES limit set for the oceanic discharge of GOM produced water is 42 ppm TPH. If the acid and TPH concentrations in produced water were roughly equivalent, the TPH analysis might be artificially overstated by the presence of high levels of organic acids. To estimate the proportion of acids extracted by methylene chloride and hexane, 1 L of GOM brine simulant was spiked with 1000 µg each of formic, acetic, propionic, and n-butyric acids. Both a TEM fraction and a HEX fraction were generated from the spiked brine using the standard extraction protocol for produced water. The fractions were analyzed by GC/FID and IC. Only propionic and n-butyric acids were sufficiently resolved from the solvent fronts to quantitate these particular acids by GC/FID. These acids fell within GC retention times of C<sub>7</sub> and C<sub>8</sub> n-alkanes. Less than 5% total acid was recovered in the TEM fraction, corresponding to an equivalent TPH content of 51 ppm in a produced water sample. Less than 1% organic acids were recovered in the HEX fraction, corresponding to a value of 23 ppm TPH in produced water. The organic acid concentration observed in PERF oil/water contacts indicates that no more than 30 ppm organic acids were typically present. Estimating that no more than 5% of the acids are recovered in the TEM fraction, TPH would typically be overestimated by less than 2 ppm.

The TEM and HEX fractions were analyzed by IC by first diluting the samples by a factor of 20 with acetonitrile. Relative ratios for formic, acetic, propionic, and n-butyric acids were 1: 100:500:2500, respectively. The total acid concentration found in the TEM fraction was approximately four times that determined by GC/FID, due to the ability to the ability of the

method to resolve the formic and acetic acid peaks from the solvent front. Only propionic and butyric acids were observed in the HEX fraction at a concentration ratio of 1: 10, respectively. Eight percent of the total acid mass found in TEM was transferred into the HEX fraction. Losses of formic and acetic acids from this fraction are attributed to volatilization in the solvent exchange necessary to produce the HEX fraction. Assuming that the typical TPH concentration in HEX fractions derived **from the** processing of produced water samples was approximately 12 ppm and that 30 ppm of organic acids was present in the water, the TPH concentration would be elevated by no more than 0.5 ppm in the HEX fraction.

Several experimental parameters were found to affect the concentrations of organic acids found in produced water. The oil/water ratio did not consistently affect the levels of organic acids in produced water (Fig. 8.13). However, the acid levels decreased as the brine simulant became more alkaline, a trend opposite to that observed for hydrocarbon content in the TEM fraction (Fig. 8.14). The organic acid concentration was near zero at 42,000 ppm chloride; acetic and propionic acids were consistently present at about 10 ppm at the higher salinity levels (Fig. 8.15). Concentrations, of both acetic acid and TPH **were** elevated at temperatures greater than 50°C (Fig. 8.16). Elevating pressure did not significantly affect organic acid content in produced water.

In addition to organic acids, the chromatograms of produced water included elution peaks for the primary anion, chloride, as well as a minor phosphate constituent (Fig. 8.17). Phosphate was present at slightly greater than 1200 ppm at room temperature; however it was lost as the temperature increased. Only 300 ppm of phosphate was contained in the produced water of a 75°C oil/water contact experiment. Phosphate **was** most prevalent at near-neutral pH levels and was rapidly lost when the pH was increased or decreased. Phosphate concentration showed no consistent trends with either the salinity or the water cut.

## **8.9 PRESENCE OF INORGANIC CONSTITUENTS IN CRUDE AND PRODUCED WATER**

Alkaline and alkaline-earth metals were the primary metal components in the deep-well crude according to results of ICP-AES analyses; their total concentration was about 7 g/kg. The primary transition metals are nickel (25 mg/kg); silicon (248 mg/kg); iron (22 mg/kg); vanadium (67 mg/kg) and zinc (6 mg/kg). Only selenium (3.5 mg/kg) approached the level to be considered a toxicity characteristic of the oil.

Results of inorganic chemical analysis of produced water samples indicated that only cadmium and antimony were occasionally extracted into the water phase at the level of 0.1 ppm. The concentrations of these metals did not appear to be affected by variations in pH or water/oil cut. The alkaline and alkaline-earth **metal concentrations** in produced water were equivalent to those found in the starting brine simulant.

**Table 8.3. WSO distribution coefficients within chemical classes**

Experiment	Water cut (%)	Simulant volume (mL)	Oil volume (mL)	Distribution coefficients, $K_d$				
				TEM	HEX	TSAT	TARO	TPOL (GC)
PERF-9	20	1000	4000	6,719	5,095	1,457,002	88,523	NA <sup>a</sup>
PERF-7	50	1000	1000	5,471	3,208	1,746,309	46,785	2,953
PERF-10	50	1100	1100	8,257	7,339	1,267,370	144,889	2,560
PERF-6	67	1000	500	17,291	6,671	1,188,170	627,449	NA
PERF-8	80	1000	250	8,132	4,522	2,156,409	139,427	NA
PERF-16	80	2400	600	12,097	6,606	4,794,989	128,948	4,468
PERFPO	80	3000	750	9,698	12,009	no extraction	no extraction	14,358
PERF-25	73	200	73	52,006	20,220	4648566	128,756	3,927

<sup>a</sup>NA = not analyzed.

**Table 8.4. WSO distribution coefficients within carbon ranges for TEM fractions**

Experiment	Water cut (%)	Simulant volume (mL)	Oil volume (mL)	$K_d$ in TEM Fractions		
				C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>
PERF-9	20	1000	4000	12,195	4,375	No extraction
PERF-7	50	1000	1000	9,525	3,637	No extraction
PERF-10	50	1100	1100	18,654	4,937	No extraction
PERF-6	67	1000	500	21,306	14,346	No extraction
PERF-8	80	1000	250	16,025	5,876	No extraction
PERF-16	80	2400	600	16,792	9,153	No extraction.
PERF-20	80	3000	750	14,941	6,873	No extraction
PERF-25	73	200	73	51.388	54.557	No extraction



## 9 . CONCLUSIONS

Quantitative characterization data were collected for the WSO content of produced water as a function of several experimental parameters. The WSO content was defined on the basis of an open LC column procedure in which the total organic material was fractionated into methylene chloride-soluble, hexane-soluble, saturated hydrocarbon, aromatic hydrocarbon, and polar hydrocarbon classes. The proportions of gas-, diesel- and oil-TPH within each fraction were then determined by GC/FID analysis. Data from this protocol were supported by the results using USEPA SW-846 Method 8270 C for semi-volatile organic analytes and infrared spectroscopy for TPH content. Analyses were also performed for a number of in-organic constituents .

The GOM crude was obtained from a single deep-well site. The TPH contents in this sample were 50, 42, and 8% for the  $C_6-C_{10}$ ,  $C_{10}-C_{20}$ , and  $C_{20}-C_{28}$  ranges, respectively. The principal SVOCs in the deep-well oil were acetophenone (0.6 g/kg) and methylnaphthalene (0.4 g/kg), which represent less than 2% of the total organic content. Chemical classification by the open LC protocol was used to account for the remaining 98% of the organic character of the oil. Methylene chloride-extractable material (TEM) in the deep-well oil was present at 200-400 g/kg. Approximately 35% of this material could be exchanged into a hexane matrix. The relative composition of the hexane matrix was 35% saturated hydrocarbons, 15% aromatic components, and 50% polar material. The visibly colored red-brown material was primarily present in the polar fraction. Alkaline and alkaline-earth metals were the primary ICP metal components in the deep-well crude; their total concentration was about 17 g/kg. Several RCRA metals were detected slightly above the ICP detection limit; however, they were not present above the USEPA-defined toxicity limits.

The sample of GOM crude described above was contacted with a GOM brine simulant in a WAF vessel to generate produced water samples. The total concentration of WSO observed in the oil/brine simulant contacts ranged from 20-30 ppm. Four SVOCs were identified: 1-methylnaphthalene (10 ppb), 2-methylnaphthalene (9 ppb), naphthalene (14 ppb), and phenol (30 ppb). The distribution within the chemical fractions differed significantly from that of the deep crude. Chemical fractionation of the total extractable materials suggested that 80 to 90% of the WSO was present as polar compounds; the next largest fraction was that of aromatic materials. The visibly colored material was primarily present in the light-yellow aromatic fraction. TEM organic matter was distributed between the low and midrange carbon masses, similar to oil data.

Variation in the water/oil ratio indicated that there was a slight negative trend in the WSO content in all chemical fractions as the water cut increased. This effect was a combination of, the decreasing solubility of the  $C_{10}-C_{20}$  mass and the increasing solubility of the  $C_6-C_{10}$  carbon material. Analytical results obtained in the characterization of crude oil were used in combination with data from water cut experiments to determine approximate distribution coefficients ( $K_d$  values) of WSO in GOM brine as a function of percent water cut and brine volume. Distribution data imply that that brine was saturated with WSO for water cuts of less than 50%.

The variation of TDS with salinity did not significantly alter the WSO content in produced water; however, the pH of the brine did affect TPH content. Normally, the pH of the GOM brine simulant was buffered to 7-7.5 by the presence of an oil layer. When the buffering capacity of oil was exhausted with base, the WSO content almost doubled for pH values above 7. This was primarily due to the enhanced solubility of polar organic compounds that deprotonate at higher

pH values. There was an enhanced solubility of the C<sub>10</sub>–C<sub>20</sub> carbon range material and a decline in C<sub>6</sub>–C<sub>10</sub> mass for pH values greater than 7.

The chemical character of WSO did not appear to be significantly dependent upon solution temperature, although the percentage mass in each carbon range was altered. The mass of WSO in the C<sub>6</sub>–C<sub>10</sub> carbon range decreased while that in the C<sub>10</sub>–C<sub>20</sub> range increased. Thus, the net quantity of organics in the produced water increased by 5 ppm for a 60°C elevation in solution temperature.

Over the pressure range of 1-60 bar, total WSO content increased by less than 10%. The additional quantity of WSO was primarily associated with the polar fraction. The most notable difference in the content of produced water derived from high pressure contacts was the increased amounts of C<sub>6</sub>–C<sub>10</sub> range material, particularly in the polar fraction and to a lesser extent in the aromatic fraction. There was considerable scatter in pressure data due to the low volumes of oil and water used in this series of experiments. The low quantities of saturated hydrocarbon fractions made it difficult to discern any trends for this chemical fraction.

Three organic acids were typically found in the produced-water samples: formic, acetic, and propionic. Acetic acid, which was present at 10-30 ppm (similar to that of TPH), made up the bulk of the acid mass. However, acid levels decreased as the brine simulant became more alkaline, a trend opposite from that observed for the hydrocarbon content in the TEM fraction. Both salinity and temperature slightly elevated the organic acid concentration, particularly that of acetic acid. Less than 5% total acid in produced water was recovered in the TEM fraction; less than 1% organic acids was recovered in the HEX fraction. For 30 ppm of total acid present in produced water, it was estimated that the TPH content would be overestimated by less than 2 ppm if methylene chloride were used to extract WSO from produced water. The TPH content would be overestimated by 0.5 ppm if hexane were used for solvent extraction. Assuming a constant distribution coefficient, overestimation would then increase in a linear fashion relative to total acid concentration in the produced water.

Of the five experimental parameters tested (percent water cut, pH, salinity, temperature, and pressure), the factor that most controlled the total WSO in produced water was that of aqueous phase pH. Beyond a value of pH 7 significant quantities of C<sub>10</sub>–C<sub>20</sub> range material became markedly soluble as they deprotonated in basic GOM brine. Both the absolute and relative volumes of GOM brine and crude additionally affected total WSO. Produced water appeared to reach a saturation level of WSO at a 50% water/oil ratio. Pressure slightly enhanced WSO by increasing the relative quantity of C<sub>6</sub>–C<sub>10</sub> range material. Temperature primarily altered the relative ratio of the carbon ranges within the WSO without significantly elevating the total WSO in the GOM brine. Salinity had the least affect on the chemical character or the carbon size of WSO in produced water.

## 10. REFERENCES

M. A. Mills, T. J. McDonald, J. S. Bonner, M. A. Simon, and R. L. Autenrieth, "Method for Quantifying the Fate of Petroleum in the Environment," *Chemosphere* **39**, 2563-82 (1999).

M. A. Franson, (ed.), *Standard Methods for the Examination of Water and Wastewater*, 18th ed., American Public Health Association, Washington, D.C., 1992, pp. 2-23 to 2-28.

## Figures for Text of Report



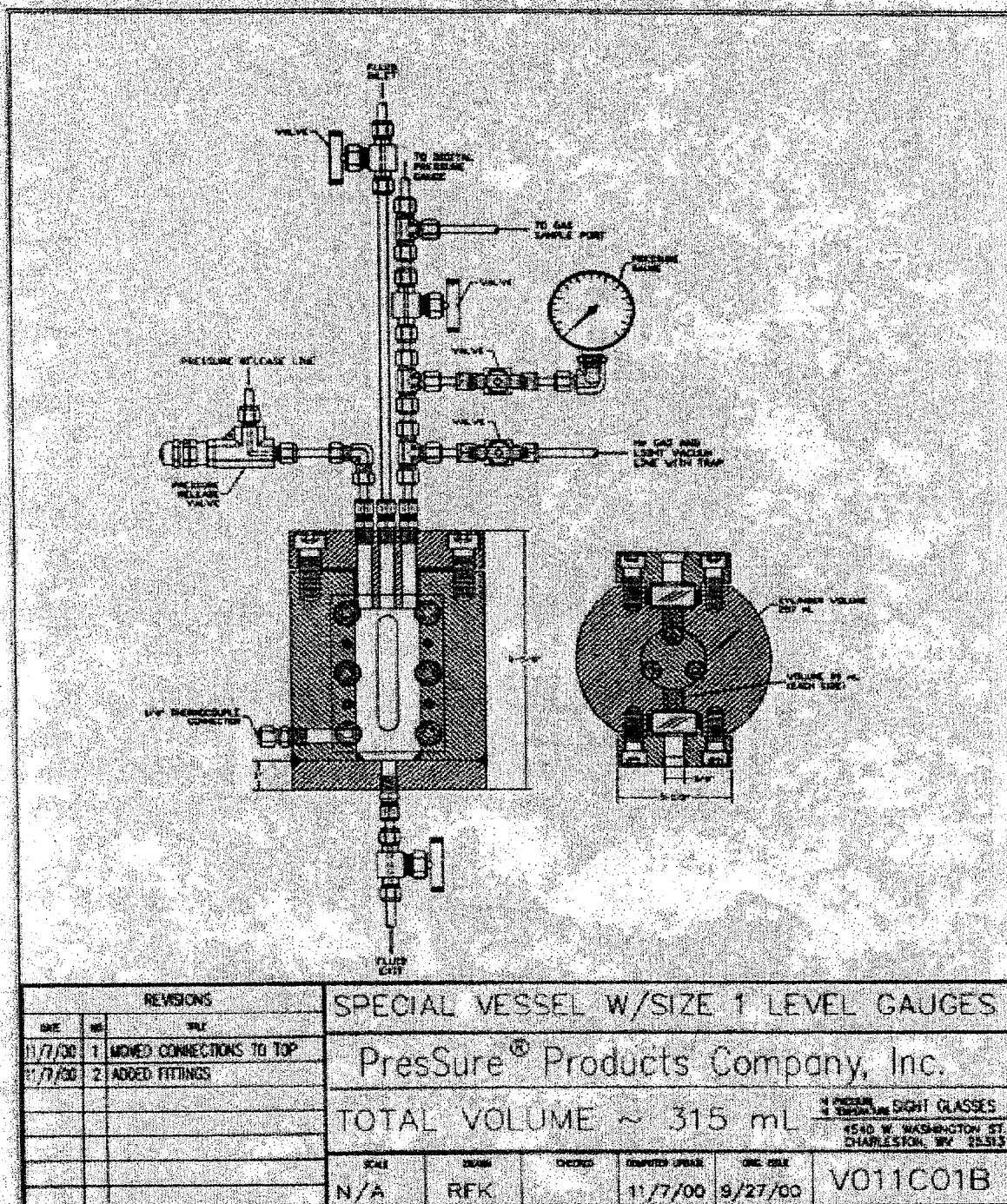
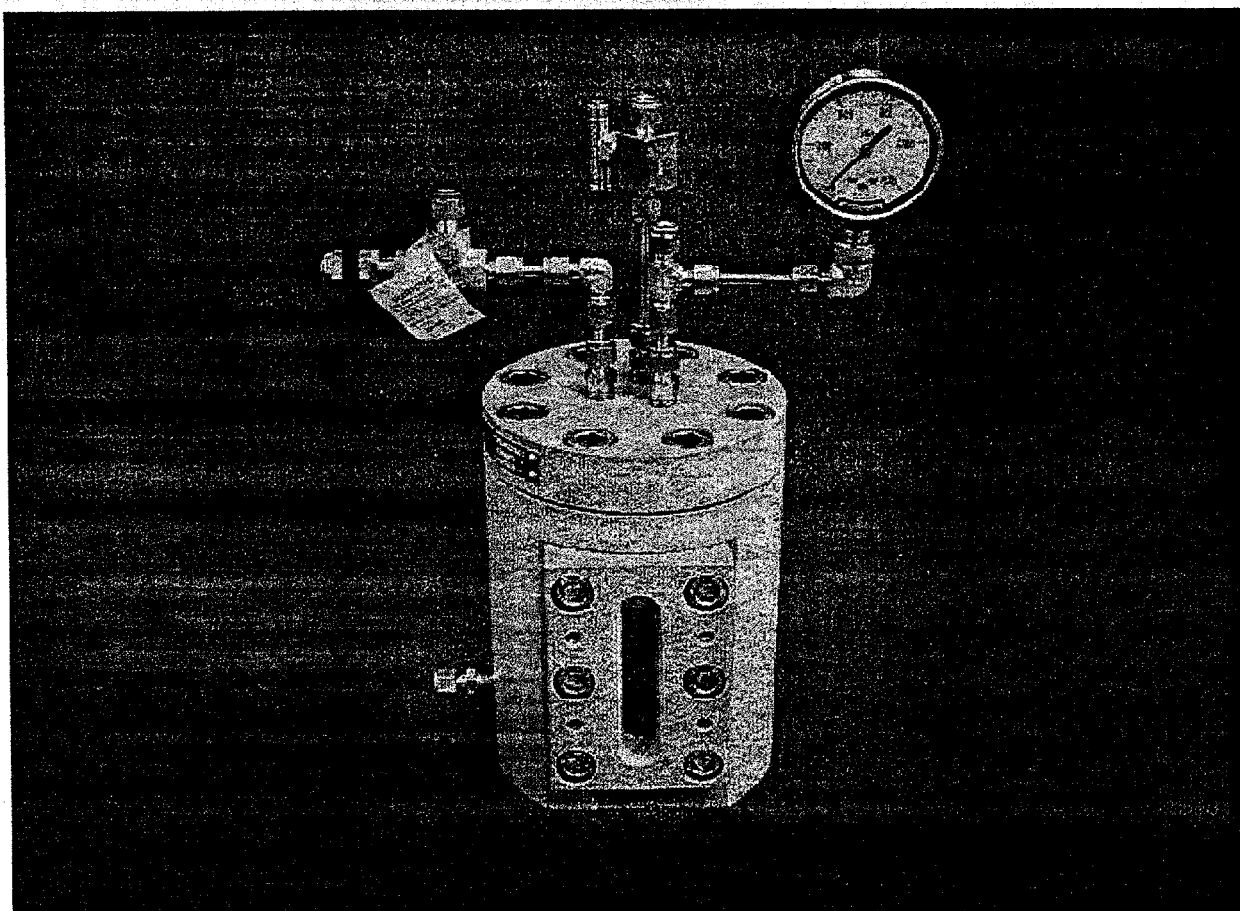
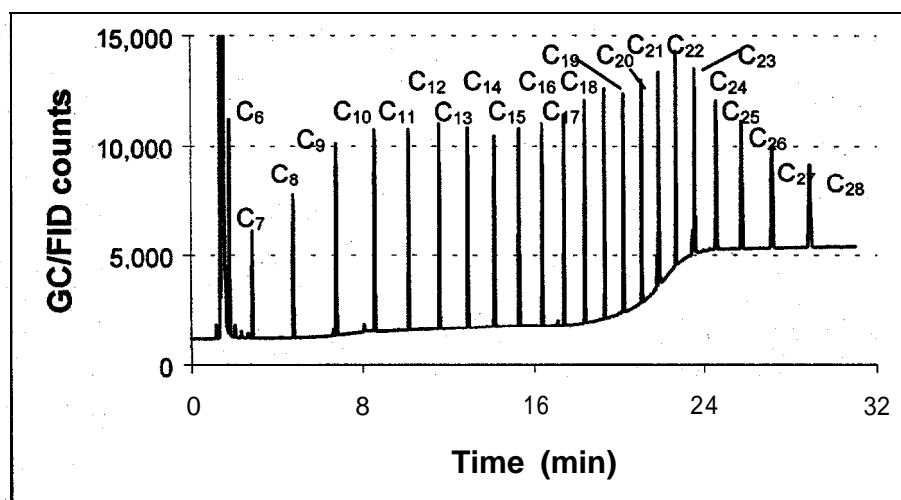


Fig. 4.1. Schematic diagram of oil/water contact pressure vessel.

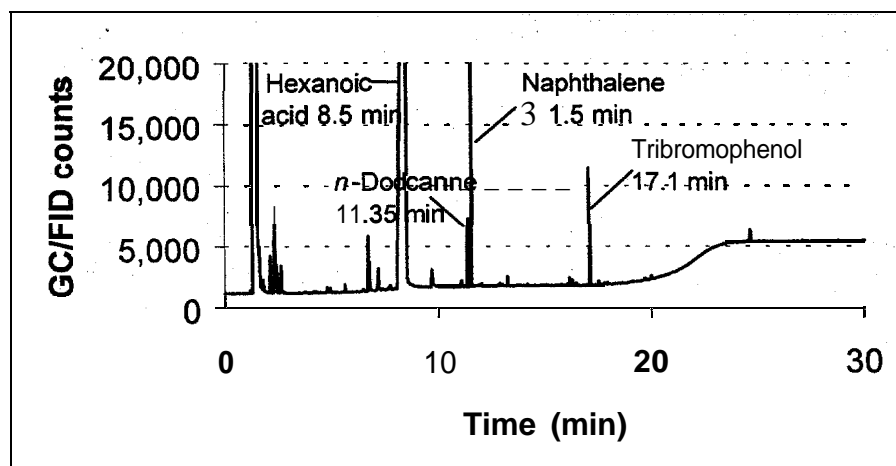




**Fig. 4.2. Oil/brine contact pressure vessel.**  
P.

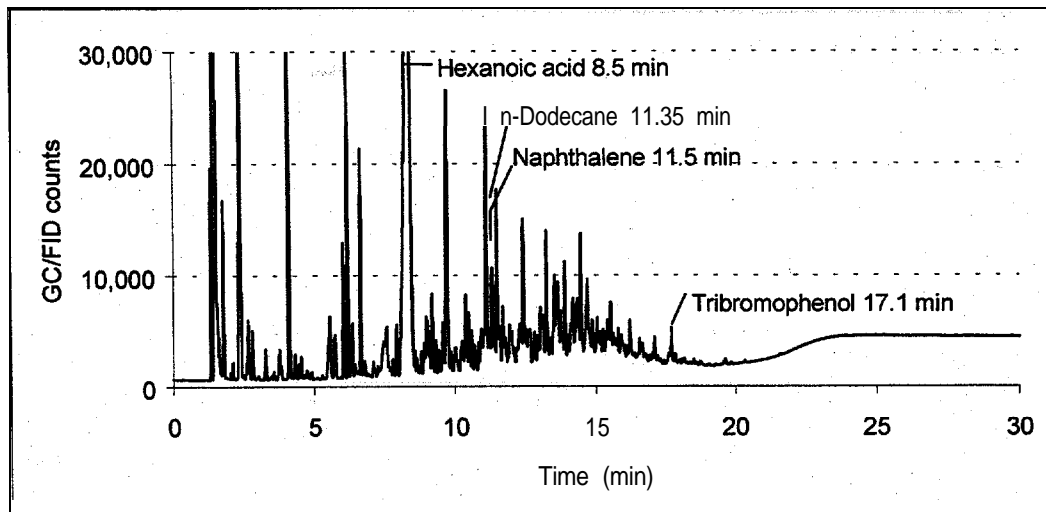


**Fig. 7.1. GC/FID chromatogram of *n*-alkane standards.**  
(2- $\mu$ L injection of 15 ppm of each alkane)

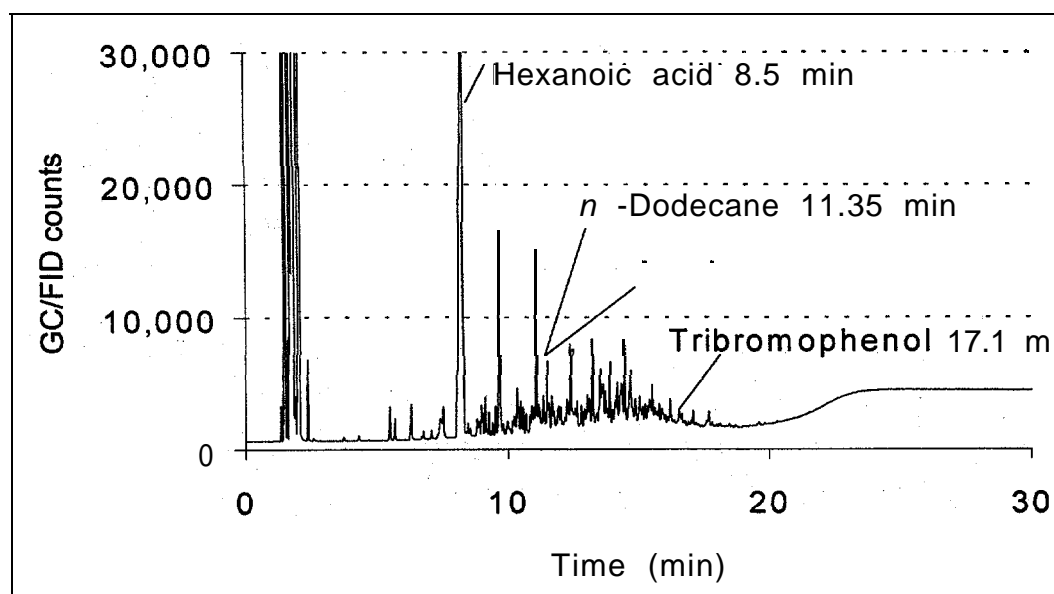


**Fig. 7.2. GC/FID chromatogram of surrogate recovery standards.**  
(2- $\mu$ L injection of methylene chloride (TEM) extract)

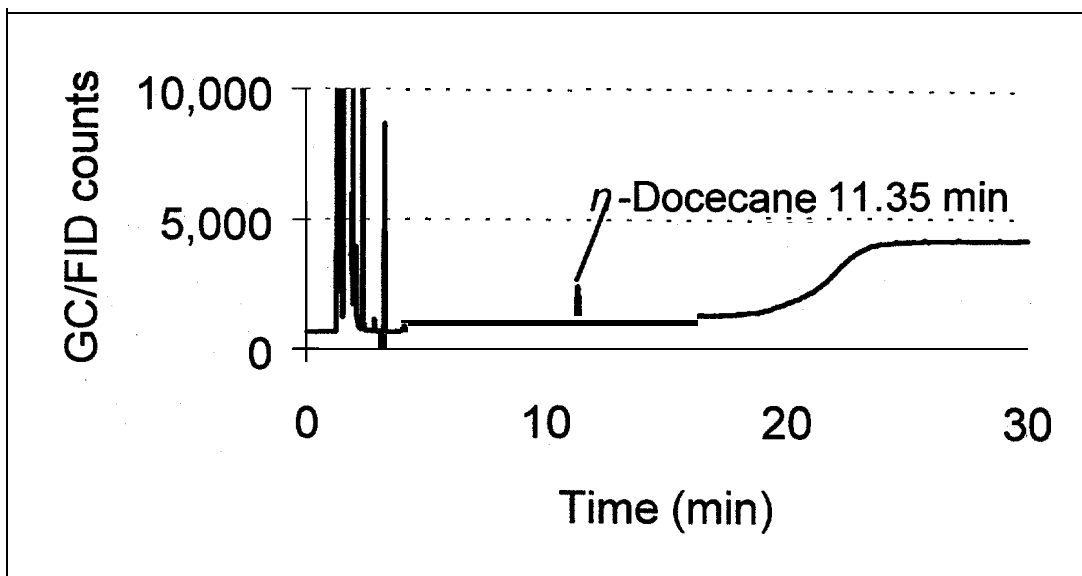




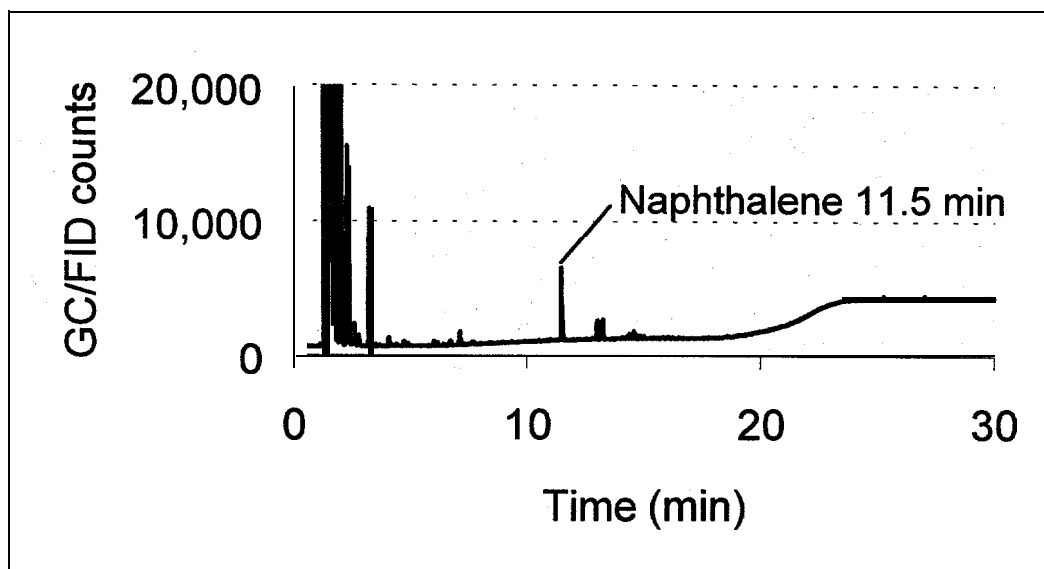
**Fig. 7.3. GC/FID chromatogram of TEM fraction.**  
(0.5- $\mu$ L injection of 1.2 mL of TEM-20 fraction)



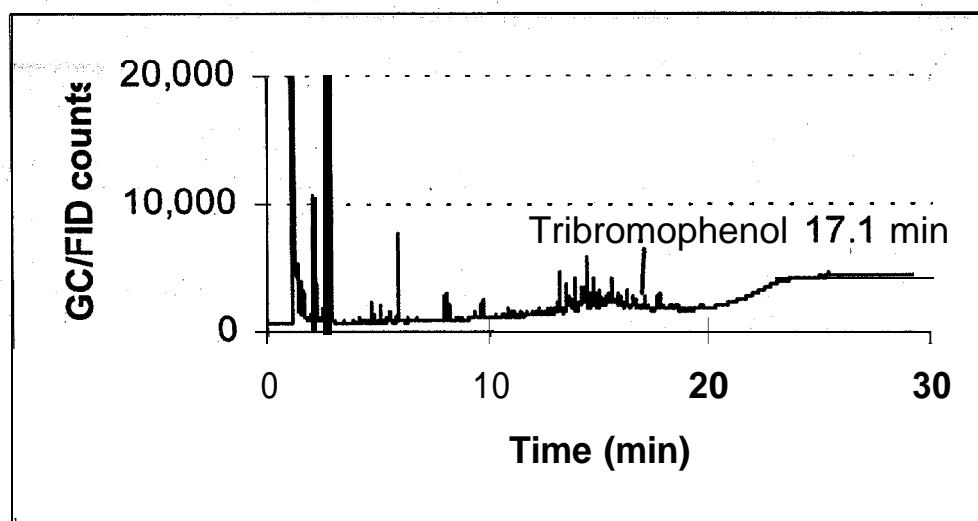
**Fig. 7.4. GC/FID chromatogram of HEX fraction.**  
(0.5- $\mu$ L injection of 1.0 mL of HEX-20 Fraction)



**Fig. 7.5. GC/FID chromatogram of TSAT fraction.**  
(3- $\mu$ L injection of 1.1 mL of TSAT-20 fraction)



**Fig. 7.6. GC/FID chromatogram of TARO fraction.**  
(3- $\mu$ L injection of 0.95 mL of TARO-20 fraction)



**Fig. 7.7. GC/FID chromatogram of TPOL fraction.  
(1- $\mu$ L injection of 1.0 mL of TPOL-20 fraction)**

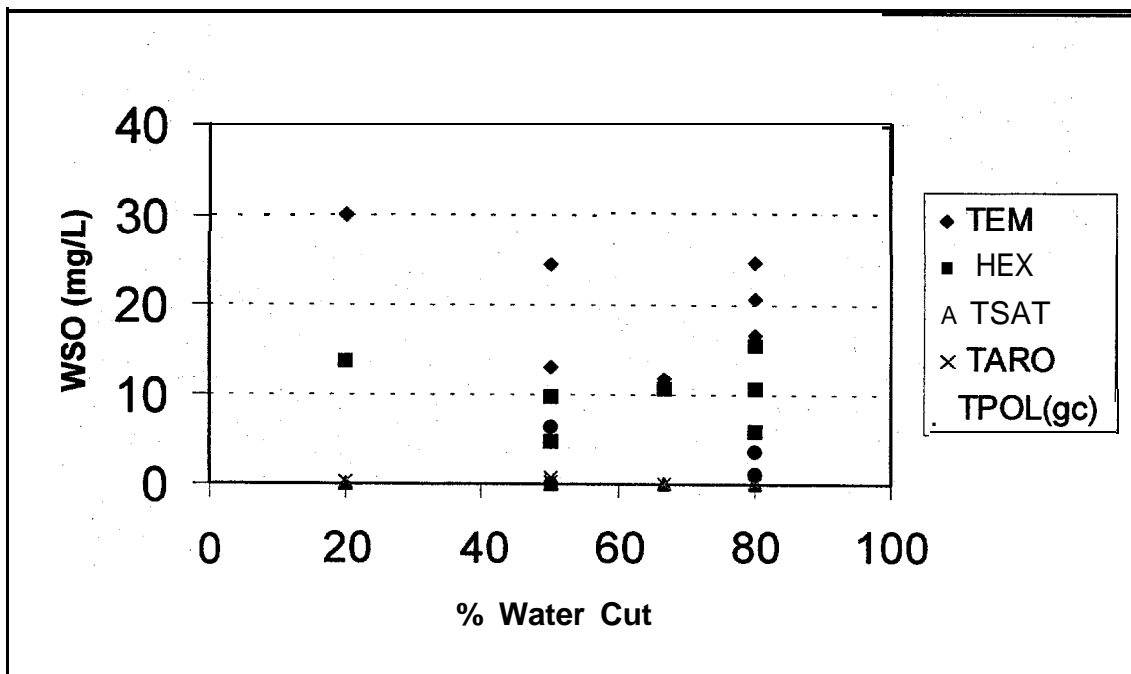


Fig. 8.1. WSO chemical character versus percent water cut.

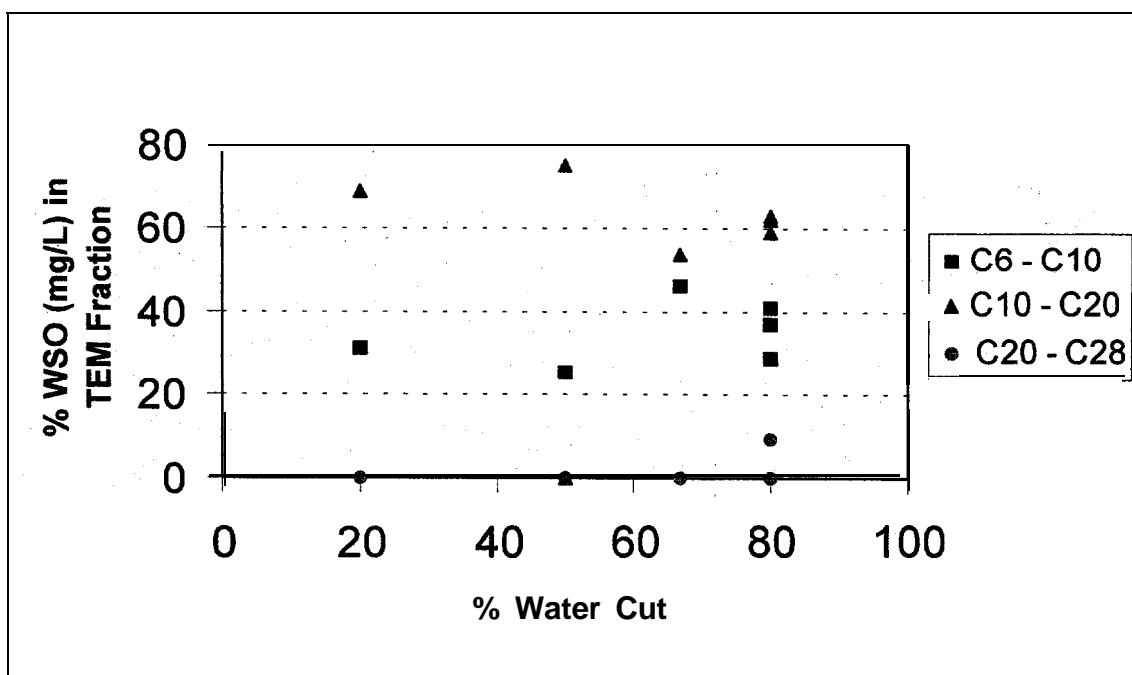
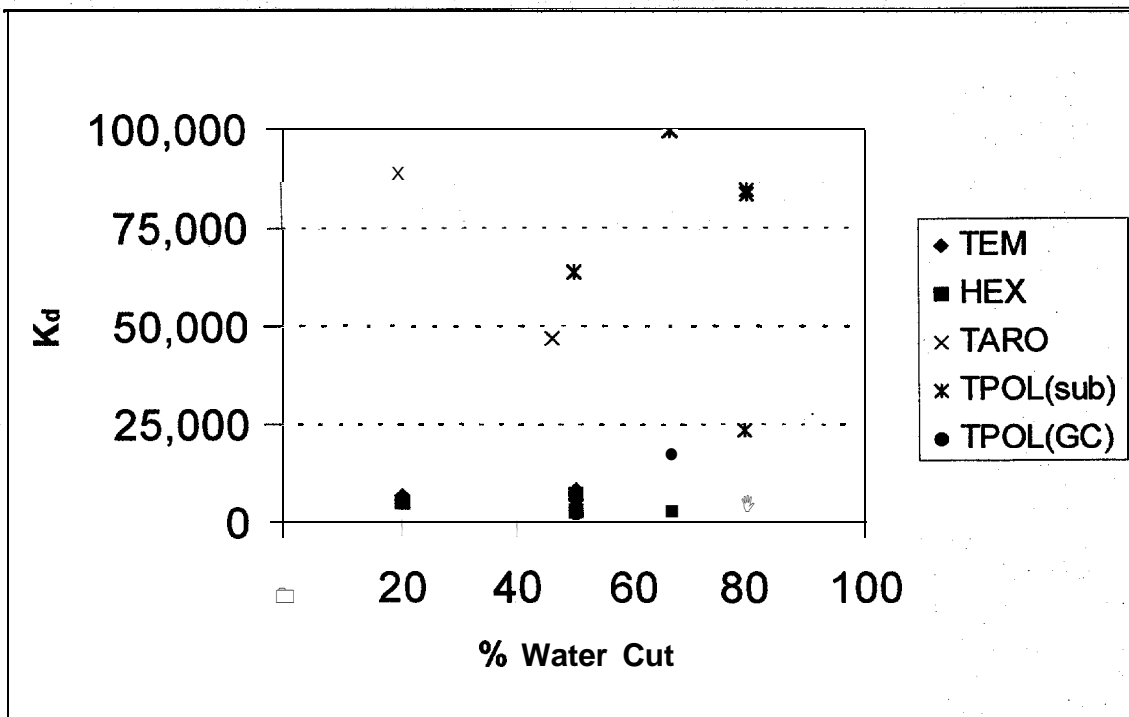


Fig. 8.2. WSO carbon content versus percent water cut.



-Fig. 8.3. WSO  $K_d$  versus percent water cut.

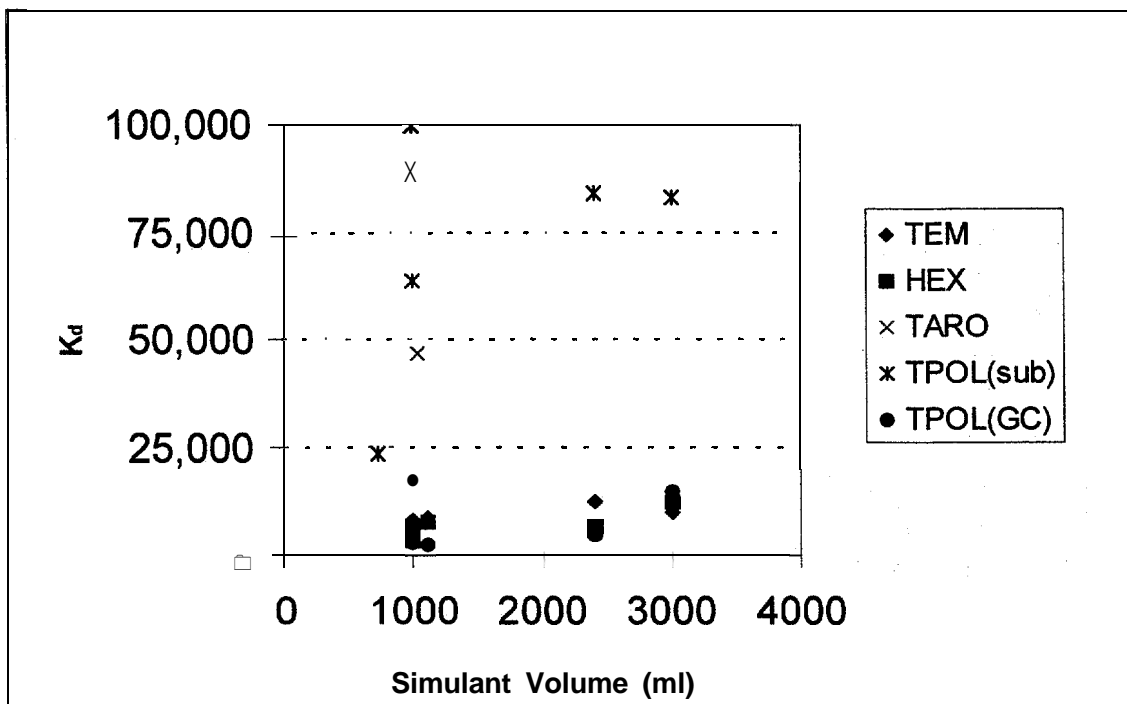


Fig. 8.4. WSO  $K_d$  versus GOM simulant volume.

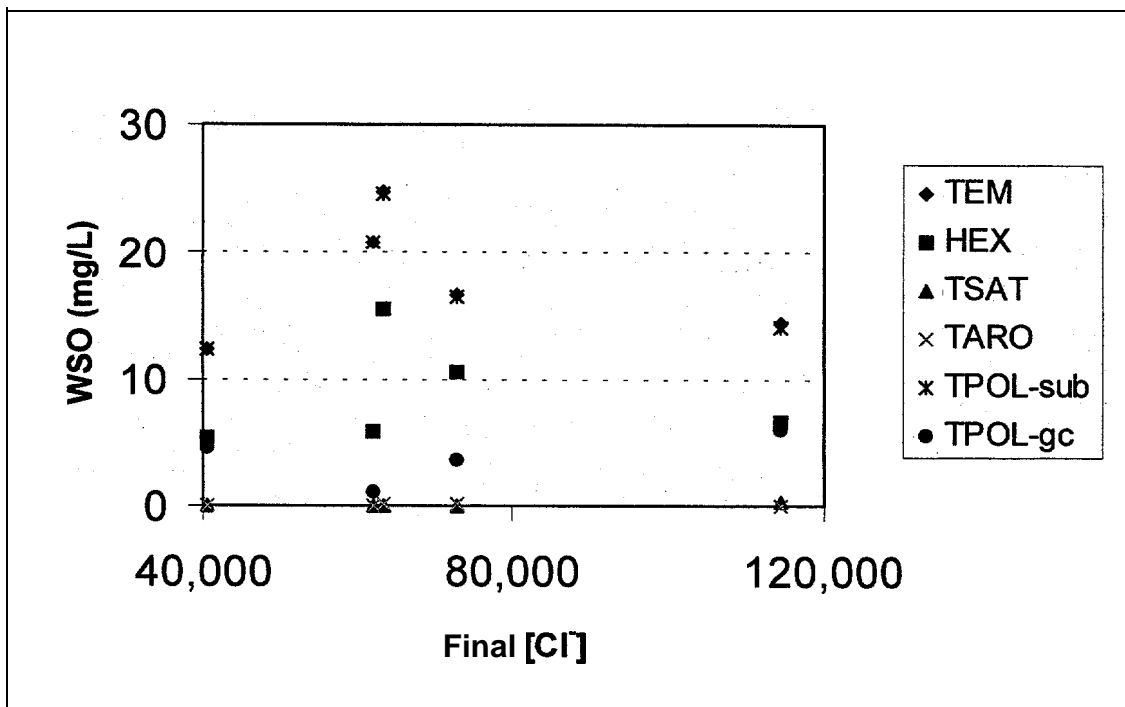


Fig. 8.5. Chemical fraction as a function of total dissolved solids (as Cl⁻).

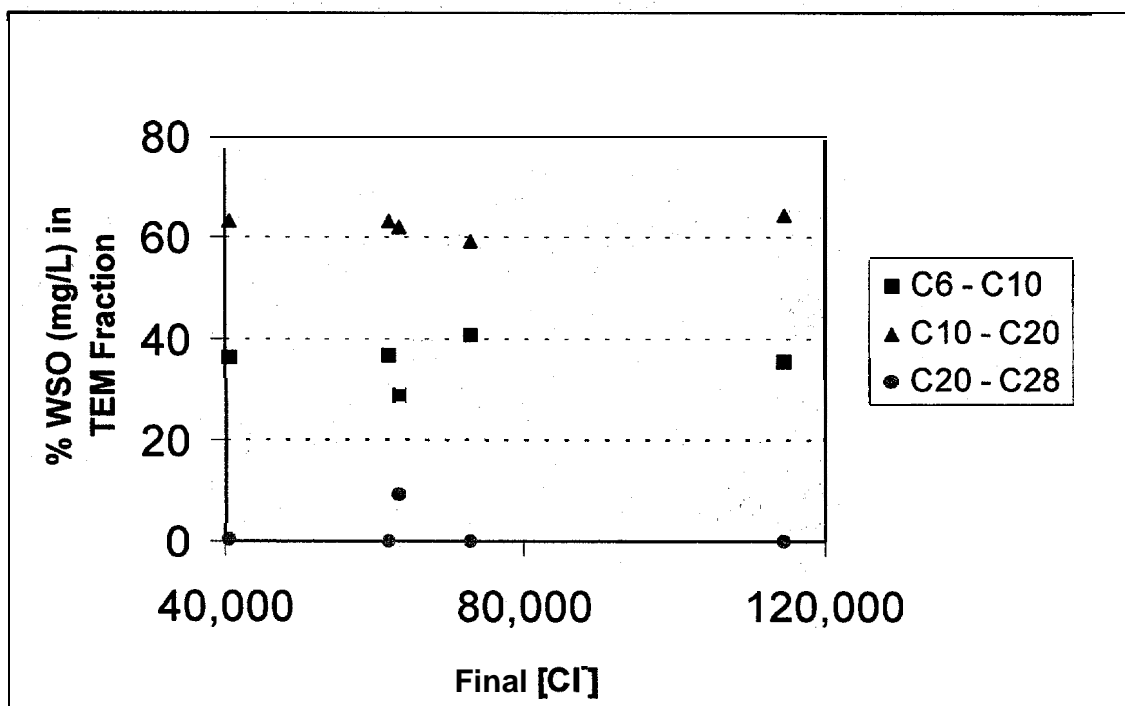


Fig. 8.6. Carbon content in TEM fractions as a function of total dissolved solids (Cl⁻).

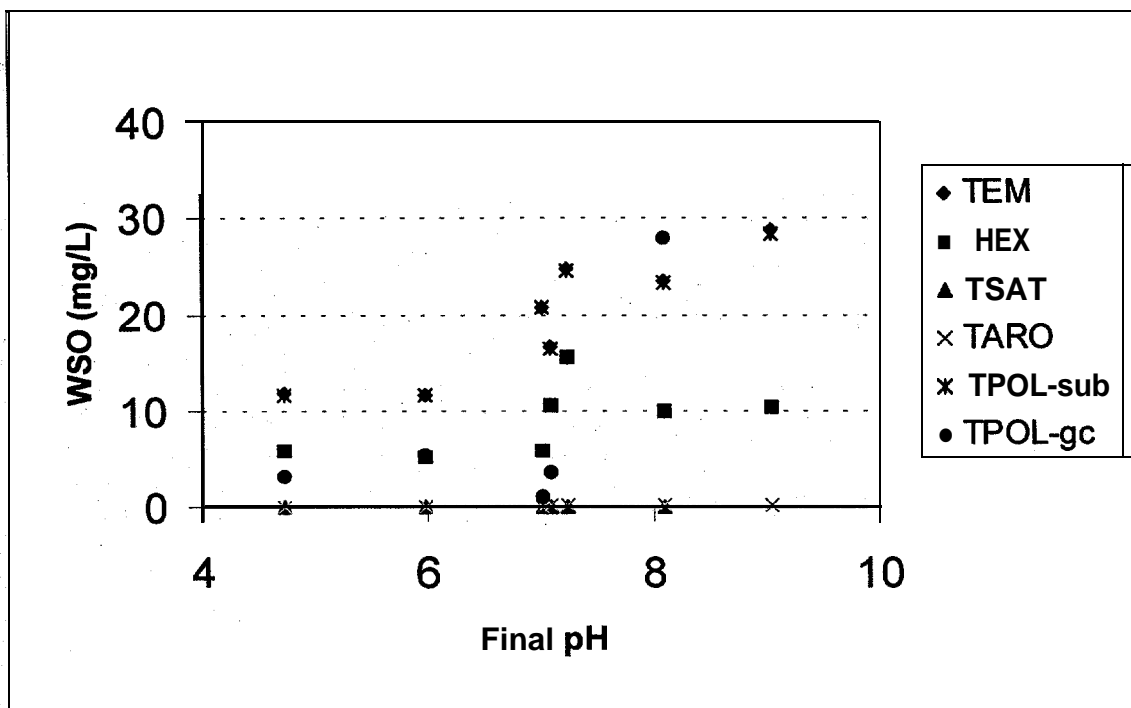


Fig. 8.7. Chemical character as a function of GOM brine pH.

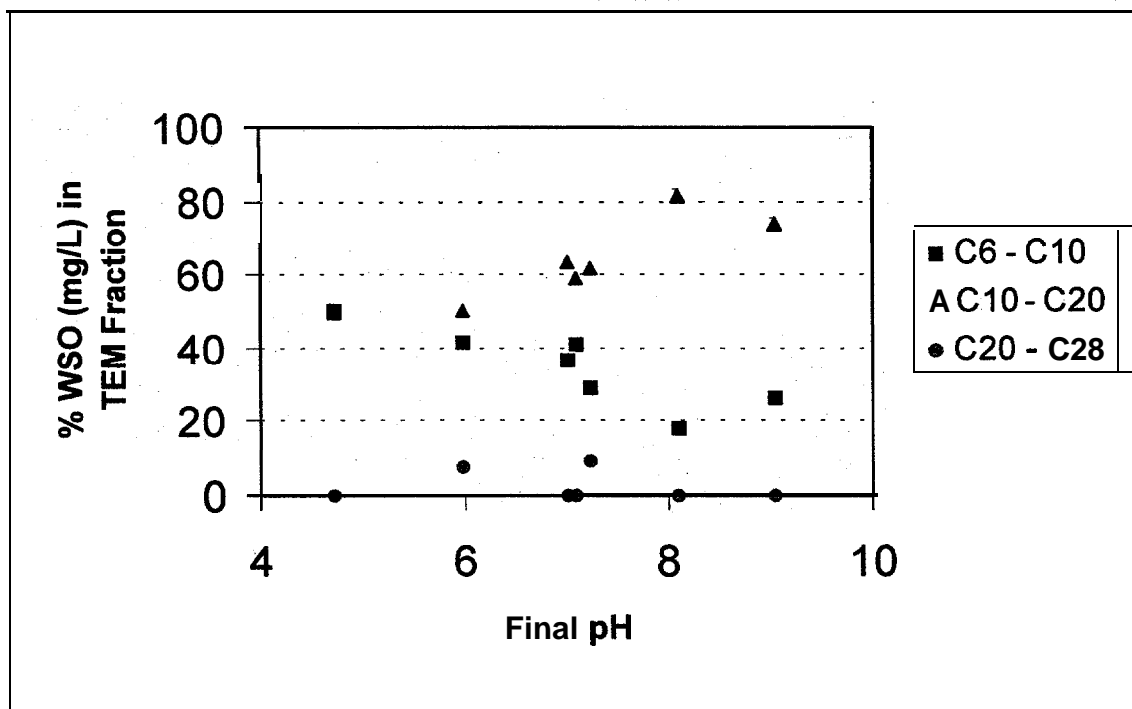


Fig. 8.8. Carbon content in TEM fractions versus GOM brine pH.

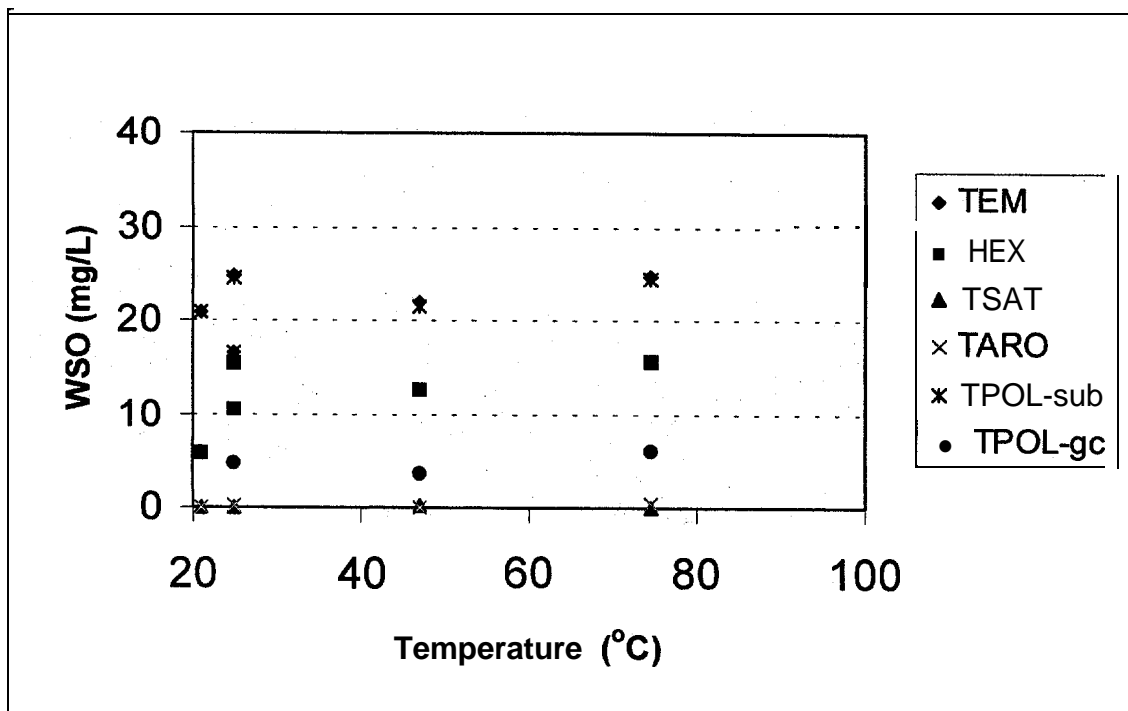


Fig. 8.9. Chemical character as a function of GOM brine temperature.

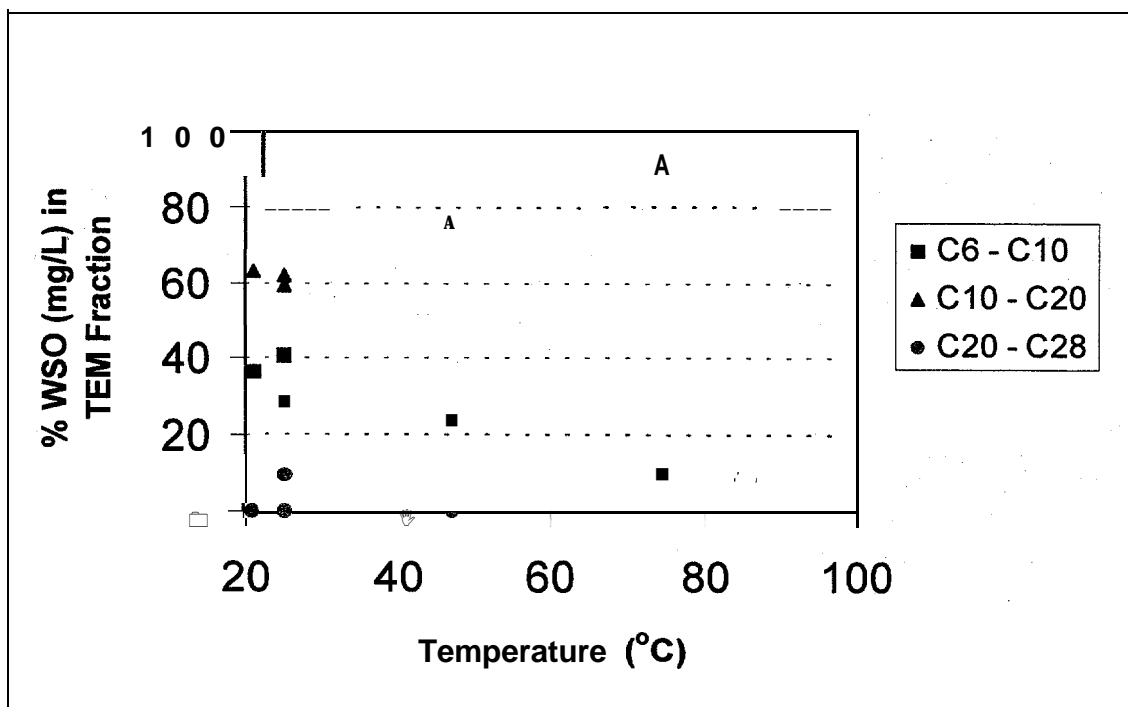


Fig. 8.10. Carbon content in TEM fractions as GOM brine temperature.



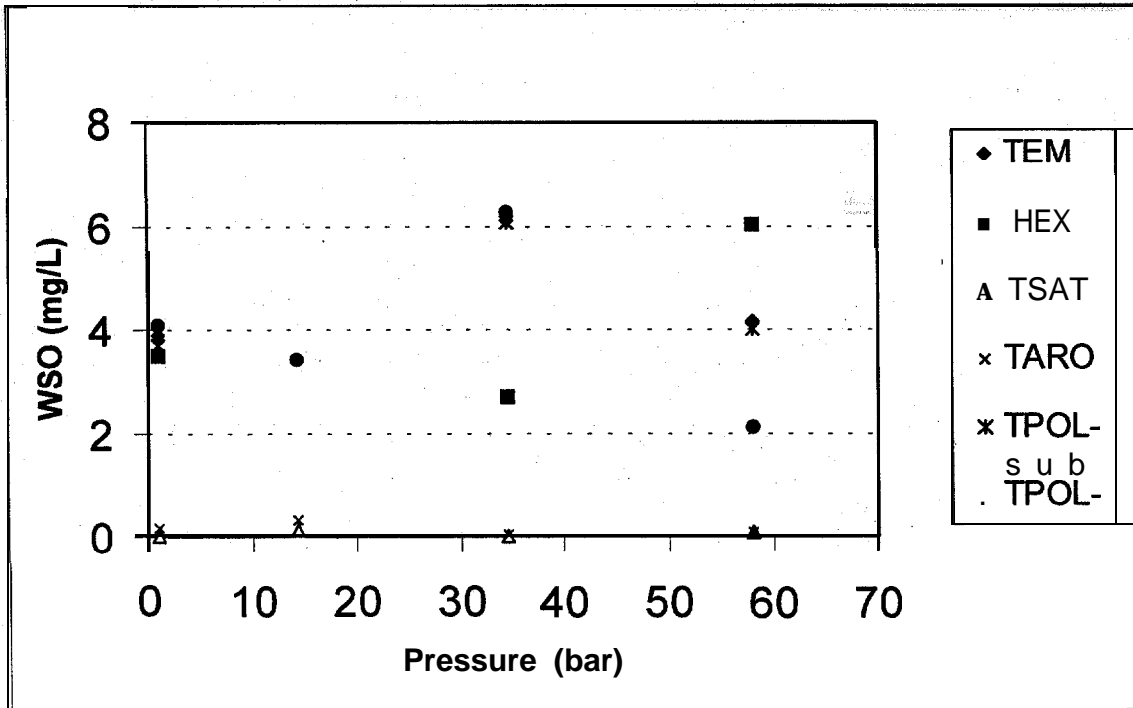


Fig. 8.11 Chemical Character as a function of GOM brine pressure.

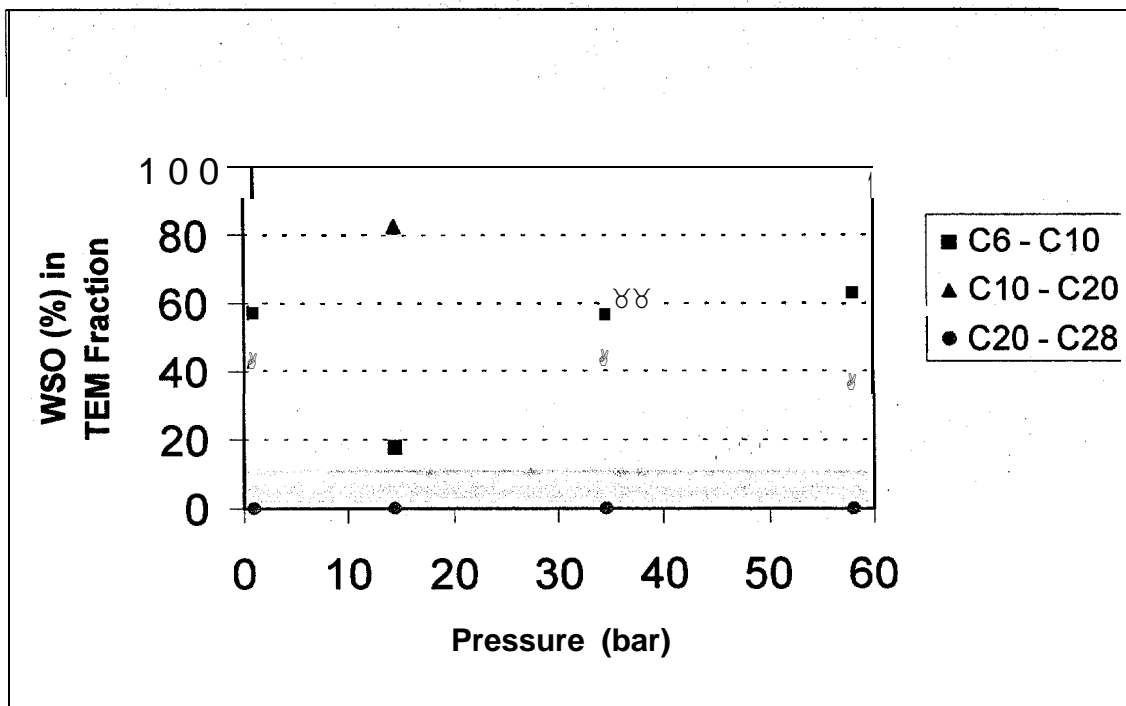


Fig. 8.12 Carbon content in TPOL as a function of GOM brine pressure.

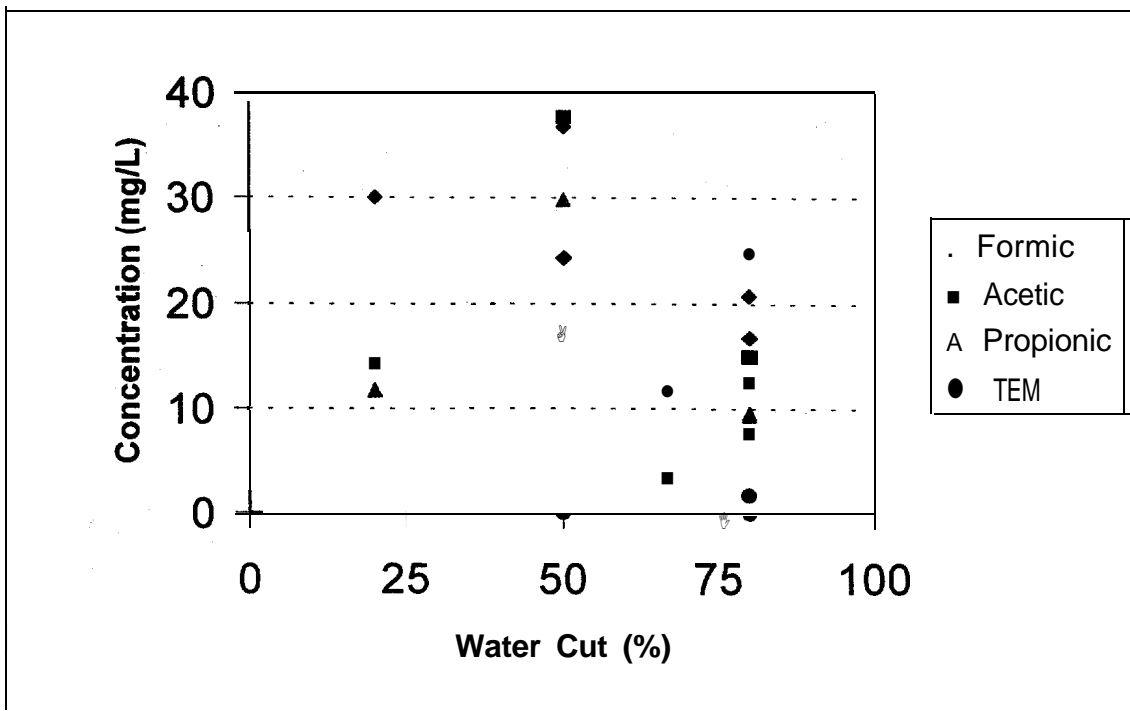


Fig. 8.13. Influence of percent water cut on organic acid content.

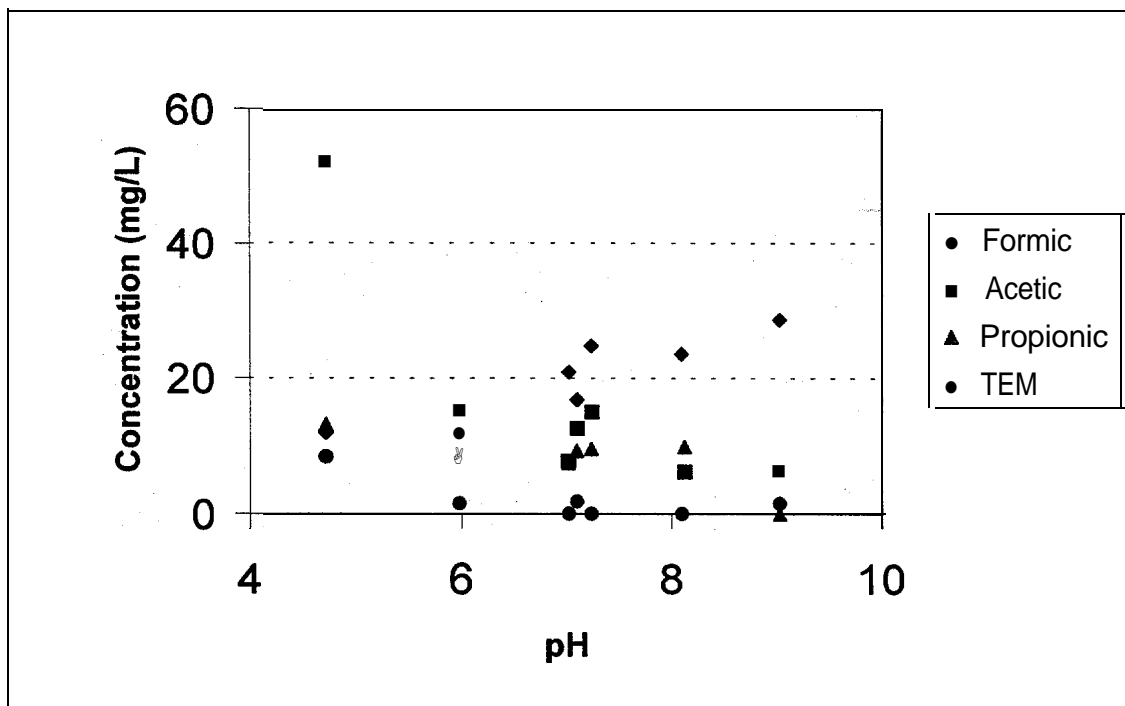


Fig. 8.14. Influence of brine pH on organic acid content.

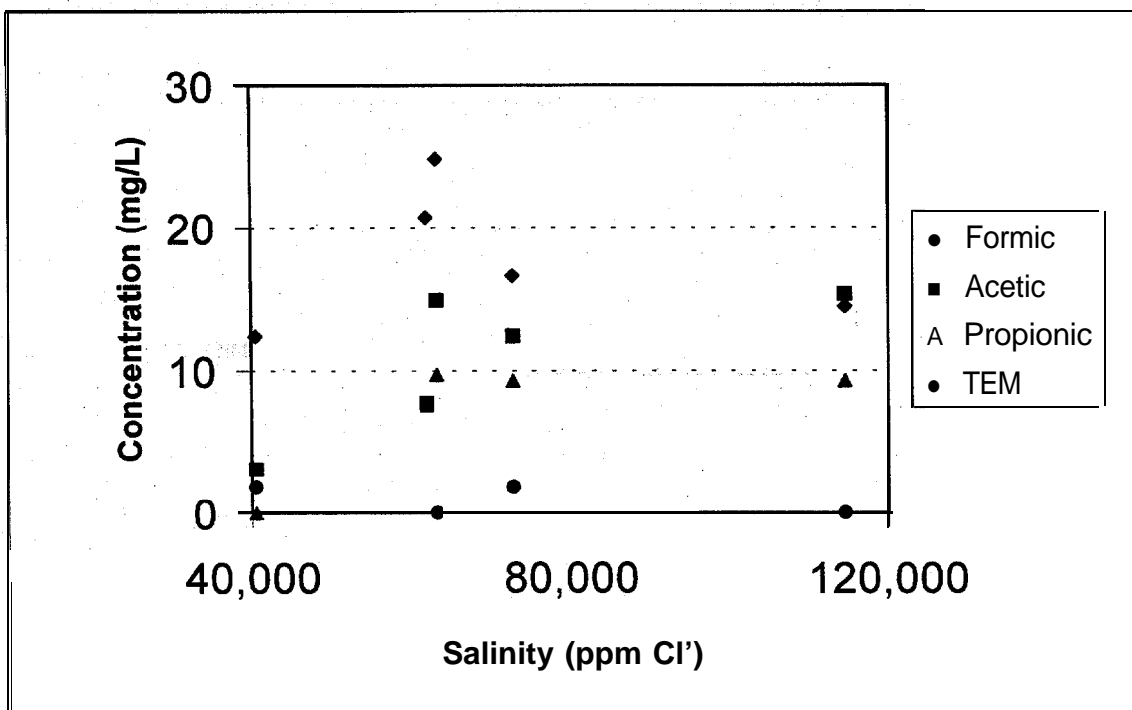


Fig. 8.15. Influence of salinity on organic acid content.

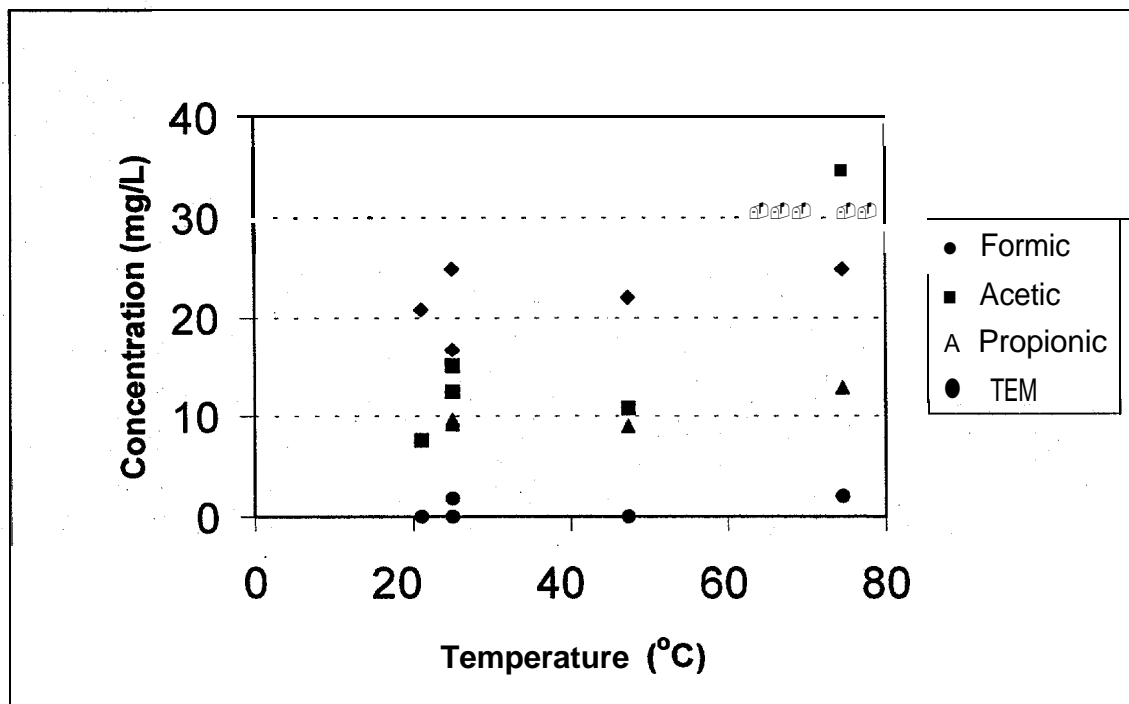
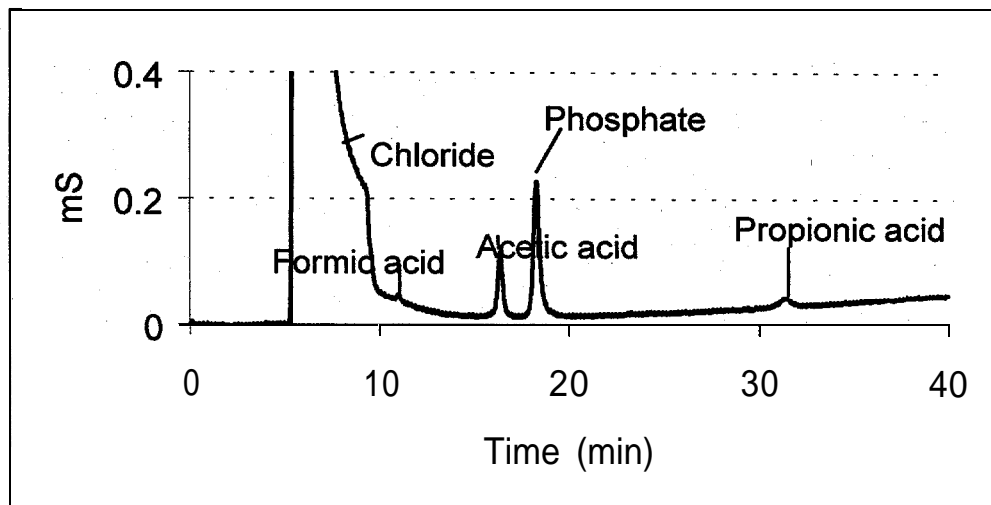


Fig. 8.16. Influence of temperature on organic acid content.



**Fig. 8.17. Ion chromatogram of organic acids in produced water.**

## APPENDIX



## APPENDIX

The tables in this appendix list data acquired for the oil/water contact experiments, analyses of oil samples, certification of GOM brine simulant, and protocol standards. These experiments are denoted by “**PERF-**”, “**Deep-**”, “**GOM-**”, and “**STD-**”, respectively, in each of the tables. The data for **PERF** designations are further delineated according to the experimental parameter that was being studied. Baseline experimental conditions were defined as 80% water cut, 65,000 ppm salinity, **pH 7** brine, **25°C**, and ambient pressure. Experiment numbers for the oil/water contacts are listed in table A. 1. At least two of these experiments were performed to study the variation in WSO content for a given parameter performed outside the baseline values.

**Table A.1 PERF experiments associated with  
each oil/water contact parameter**

Parameter	Oil/water contact
Baseline	<b>PERF-8; PERF- 16; PERF-20</b>
Water cut	<b>PERF-6; PERF-7; PERF-9; PERF-10</b>
PH	<b>PERF-11; PERF-12; PERF-13</b>
Salinity	<b>PERF-14; PERF-15</b>
Temperature	<b>PERP-17; PERF-18</b>
Pressure	<b>PERF-2 1; PERF-22; PERF-23; PERF-24; PERF-25; PERF-26; PERF-27; PERF-28</b>

The details pertaining as to the starting conditions of each experiment are summarized in Table A.2. The equilibrium concentrations of major anion species and the **pH** levels at the completion of an oil/water contact are listed in Table A.3. Corresponding cation data are presented in Table A.4. Summation of WSO content for oil/water contacts, as-received oil, and recovery standards are contained in Tables A.5 through A.7. The WSO results based on gravimetric and **IR** analyses are located in Table A.5; those based on the open LC protocol are presented in Tables A.6 and A.7.





Table A.2 Starting conditions for PERF experiments

Experiment	Simulant batch	GOM pH	Temp. (°C)	Pressure (in. Hg)	Salinity (mg/L)	Water volume (mL)	Oil volume (mL)	Total contact time (days)	Comments
GOM-1	4L-Prep	7.5	---	---	62975	---	---	---	
PERF-3blk <sup>a</sup>	GOM-1	7.5	25.0	24.06	62975	1000	0	---	+ 312.5 ppm Na <sup>+</sup> plus 8640 ppm SO <sub>4</sub> <sup>2-</sup>
PERF-4oil	---	7.5	25.0	---	---	0	2	---	
GOM-2	4L-Prep	6.9	---	---	60793	---	---	---	Start: no iron added to GOM simulant
PERF-6	GOM-2	6.9	25.0	29.06	60793	1000	500	4.83	Start: addition of recovery standards
PERF-7	GOM-2	6.9	25.0	29.06	60793	1000	1000	5.76	Start: gravimetric analysis for TEM and HEX fractions only
PERF-8	GOM-2	6.9	25.0	29.06	60793	1000	250	4.00	
GOM-3	4L-Prep	6.9	---	---	65308	---	---	---	
PERF-9	GOM-3	6.9	25.0	29.06	65308	1000	4000	4.00	
PERF-10	GOM-3	6.9	25.0	29.06	65308	1100	1100	5.75	GOM pH adjustment to 8.7 with 50% NaOH prior to oil contact
GOM-4	4L-Prep	6.8	---	---	62143	---	---	---	
PERF-11	GOM-4	9.03	24.5	29.06	65308	1200	300	2.96	pH continuously adjusted with 50% NaOH
PERF-12	GOM-4	4.73	24.5	29.1	65308	1200	300	5.94	pH continuously adjusted with concentrated HCl
PERF-13	GOM-4	9.5	24.5	29.1	65308	1200	300	5.86	pH continuously adjusted with 50% NaOH
GOM-5	4L-Prep	7.4	---	---	40638	---	---	---	50% NaCl content in simulant
PERF-14	GOM-5	7.4	24.5	29.14	40638	1200	300	4.89	50% NaCl content in simulant
PERF-Std1 <sup>b</sup>	GOM-5	7.4	24.5	29.14	40638	1000	0	---	Start standards recovery with hexanoic acid
GOM-6	2L-Prep	6.82	---	---	---	---	---	---	---
PERF-15	GOM-6	6.82	25.0	29.31	114500	1200	300	5.33	150% NaCl content
PERF-Deep1	---	---	25.0	29.31	---	---	0.995	---	0.8391 g oil in 17.66 mL CH <sub>2</sub> Cl <sub>2</sub>
GOM-7	4L-Prep	---	---	---	62143	---	---	---	Simulant used for Std2 and Std3
G O M - 8	6L-Prep	7.1	---	---	72900	---	---	---	2 L GOM-7 + 4 L GOM-8
PERF-Std2 <sup>b</sup>	GOM-5	---	24.5	29.14	40638	1000	0	---	Standards recovery with hexanoic acid

<sup>a</sup>Process blanks<sup>b</sup>Recovery of surrogate standards

Table A.2. (continued)

Experiment	Simulant batch	GOM PH	Temp. (°C)	Pressure (in. Hg)	Salinity (mg/L)	Water volume (mL)	Oil volume (mL)	Total contact time (days)	Comments
PERF-Std3 <sup>b</sup>	GOM-5	---	24.5	29.14	40638	1000	0	---	Standards recovery with hexanoic acid and tribromophenol
PERF-16	GOM-8	7.1	25.0	29.41	72900	2400	600	6.02	Baseline with double water & oil
PERF-17	GOM-8	7.42	47.2	29.40	72900	1200	300	5.83	Temperature (47.2°C) in water bath
PERF-18	GOM-8	7.42	74.5	29.19	72900	1200	300	3.88	Temperature (74°C) in water bath
PERF-19	GOM-9	5.98	23.5	29.01	60908	1200	300	4.10	pH continuously adjusted with conc. HCl
PERF-20	GOM-9	7.02	21.0	29.07	61911	3000	750	4.92	Baseline with triple water & oil
GOM-9	5L-Prep	7.1	---	---	65820	---	---	---	4 L prepared on 11/08 plus 1 L on 11 /14/00
GOM-10	2L-Prep	6.7	---	---	72900	---	---	---	
GOM-11	2L-Prep	6.8	---	---	67860	---	---	---	
PERF-25	GOM-11		50.0	29.09	67860	200	72.8	4.00	No Hoke vessel data taken
PERF-26	GOM-11		50.0	200.00	67860	200	60.6	4.00	Constant 200 psi
PERF-27	GOM-11		50.0	500.00	67860	200	57.1	4.90	Initial pressure 600 psi; He. added once for 500->600psi; final 500 psi
PERF-28	GOM-11		50.0	850.00	67860	200	46.7	4.04	Initial pressure 1 000 psi; He added for 600->900psi; final 850 psi
PERF-Deep2	---	---	25.0	29.31	---	---	0.991	---	0.8377 g oil in 18.025 mL CH <sub>2</sub> Cl <sub>2</sub>
PERF-Deep3	---	---	25.0	29.31	---	---	0.796	---	0.6727 g mixed oil in 17.12 mL CH <sub>2</sub> Cl <sub>2</sub>
PERF-Deep4	---	---	25.0	29.31	---	---	1.689	---	1.4273 g mixed oil in 30.605 mL CH <sub>2</sub> Cl <sub>2</sub>
PERF-Deep5	---	---	25.0	29.31	---	---	0.995	---	0.8407 g mixed oil in 19.76 mL CH <sub>2</sub> Cl <sub>2</sub>

<sup>a</sup>Process blanks<sup>b</sup>Recovery of surrogate standards

**Table A.3 Equilibrium concentrations of inorganic constituent**

Experiment	Salinity (ppm Cl <sup>-</sup> )	HCO <sub>3</sub> <sup>-</sup> (ppm)	CO <sub>3</sub> <sup>2-</sup> (ppm)	pH
PERFS	62,487	137	0	6.97
PERF-7	64,734	163	0	7.02
PERF-8	63,289	164	0	7.24
<b>PERF-9</b>	63,788	101	0	6.99
PERF-10	67,856	77	0	6.25
PERF-11	59,232	139	6	9.03
PERF-12	72,793	27	0	4.73
PERF-13	71,603	32	0	8.09
PERF-14	40,140	123	0	7.6
PERF-15	114,480	89	0	7.14
PERF-16	65,791	124	0	7.09
PERF-17	66,606	74	0	7.14
PERF-18	66,333	54	0	7.26
PERF-19	66,333	21	0	5.98
PERFPO	61,911	128	0	7.02
PERF-25	65,721	NA	NA	NA
PERF-26	71,772	NA	NA	NA
PERF-27	79,647	NA	NA	NA
<b>PERF-28</b>	<b>67,050<sup>c</sup></b>	<b>NA</b>	<b>NA</b>	<b>N/i</b>

NA = Not analyzed.

**Table A.4. Equilibrium concentrations of ICP metals**

Experiment	RCRA metals by ICP (ppm)			Non-RCRA metals by ICP (ppm)				
	Se	Cd	Sb	Ba	Ca	Mg	Na	Sr
PERFB	U <sup>a</sup>	0.079 ± 0.005	co.75 ± 0.14	42.3	3200	850	33,400	140
PERF-7	U	0.080 ± 0.006	co.75 ± 0.29	44	3300	860	33,800	140
PERFS	U	0.080 ± 0.009	co.75 ± 0.20	43	3180	840	32,900	138
PERF-9	U	0.080 ± <b>0.010</b>	co.75 ± <b>0.15</b>	<b>45</b>	<b>3240</b>	860	33,500	144
PERF-10	U	0.080 ± 0.010	< <b>0.75</b> ± 0.15	40	3300	880	<b>34,200</b>	148
PERF-11	0.9 ± 0.4	0.085 ± <b>0.015</b>	co.75 ± 0.15	44	3040	850	33,300	146
PERF-12	<b>1.0 ± 0.6</b>	0.086 ± 0.022	< <b>0.75</b> ± 0.27	41	3150	840	33,300	146
PERF-13	0.7 ± 0.3	0.076 ± 0.016	< <b>1.5</b> ± 0.85	37.5	3080	840	32,270	138
PERF-14	U	<b>0.077</b> ± 0.015	U	<b>94</b>	<b>3290</b>	860	17,250	145
PERF-15	0.7 ± 0.7	0.074 ± 0.020	< <b>1.5</b> ± <b>0.77</b>	<b>90</b>	3170	823	49,418	137
PERF-Deep1	3.5 ± 1.0	U	U	26	924	138	7,000	250
PERF-16	U	0.073 ± 0.007	< <b>1.5</b> ± <b>0.53</b>	53	<b>3260</b>	853	32,800	140
PERF-17	U	0.075 ± 0.019	< <b>1.5</b> ± <b>0.45</b>	44	3280	864	33,230	142
PERF-18	U	0.066 ± 0.004	< <b>1.5</b> ± 0.92	41	3320	877	33,600	144
PERF-19	U	0.073 ± 0.001	< <b>1.5</b> ± 0.76	43	3240	834	32,430	138
<b>PERF-20</b>	U	0.071 ± 0.005	< <b>1.5</b> ± 0.32	53	3255	844	32,930	140

<sup>a</sup>U = undetected above ICP detection limit for this element.

**Table A.5 Equilibrium concentrations of organic matter by gravimetric and IR analyses**

Experiment	WSO (ppm), gravimetric		Oil (g/kg), gravimetric <sup>a</sup>		VSO (ppm), IR in produced water
	TEM fraction	HEX fraction	TEM fraction	HEX fraction	
PERFB	27.0	27.0	0.064	0.048	
PERF-7	35	26.3	0.042	0.025	
PERF-8	18.0	47.8	0.085	0.176	
PERF-9	19.2	33.3	0.0057	0.0074	
PERF-10	14.4	7.71	0.0171	0.0071	
PERF-11	28.9	27.97	0.137	0.078	
PERF-12	28.9	28.0	0.068	0.028	
PERF-13	14.4	9.00	0.094	0.057	
PERF-14	18.6	10.8	0.088	0.034	49.8
PERF-15	16	7.14	0.076	0.024	12.2
PERF-Deep1	---	---	673	547	---
PERF-16	23.4	26	0.11	0.09	10.5
PERF-17	35.2	33.0	0.17	0.13	17.6
PERF-18	46.25	42.9	0.22	0.17	23.1
PERF-19	16.1	10.9	0.08	0.04	8.7
PERF-20	16.8	14.4	0.08	0.006	10.5
PERF-25	32.2	7.4	0.11	0.02	NA
PERF-26	83.0	37.7	0.33	0.12	NA
PERF-27	39.2	26.1	0.16	0.09	NA
PERF-28	29.3	35.9	0.15	0.18	NA
PERF-Deep2	---	---	561	518	NA
PERF-Deep3	---	---	434	---	NA
PERF-Deep4	---	---	731	537	NA
PERF-Deep5	---	---	636	---	NA

<sup>a</sup> For produced water samples, an equivalent oil value was calculated by multiplying the quantity of WSO in the water by the volume of water and then dividing this number by the volume of oil used in the contact.

Table A.6 Equilibrium concentrations of WSO in open LC fractions TEM, HEX, and TSAT

Experiment Name	WSO (ppm) in TEM fraction				WSO (ppm) in HEX fraction				WSO (mg/L) in TSAT fraction			
	Total	C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>	Total	C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>	Total	C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>
PERFS	11.6	5.35	6.29	0	10.5	2.56	7.98	0.0	0.062	0.062	0.0	0.0
PERF-7	36.8	12.0	24.81	0.030	21.9	6.52	15.35	0.0	0.042	0.04	0.0	0.0
PERF-8	24.8	7.1	15.35	2.293	15.5	0.86	14.68	0.0	0.034	0.03	0.0	0.0
PERF-9	30.0	9.3	20.63	0.000	13.8	3.29	10.51	0.0	<b>0.051</b>	0.05	0.0	0.0
<b>PERF-10</b>	24.4	6.1	18.28	0.000	9.6	2.13	7.45	0.0	0.058	0.06	0.0	0.0
PERF-11	28.6	7.4	21.21	0.004	10.5	0.70	9.76	0.0	Lost	Lost	Lost	Lost
PERF-12	11.9	6.0	5.87	0.014	5.8	0.24	5.57	0.0	0.039	0.04	0.0	0.0
PERF-13	23.5	4.3	19.19	<b>0.038</b>	10.0	0.50	9.46	0.0	0.010	0.00	0.0	0.0
PERF-14	12.4	4.5	7.86	0.064	5.4	0.48	4.86	0.01	0.066	<b>0.03</b>	0.02	0.008
PERF-15	14.5	5.1	9.34	0.007	6.6	1.21	5.39	0.00	0.253	0.05	0.21	0.000
PERF-Deep1	251,028	139,915	109,871	1241	73,478	9,514	58,658	5,306	66,590	11,991	50,192	4,408
PERF-Std3	0.0	0.0	0.00	0.000	0.0	0.00	0.00	0.00	0.000	0.00	0.00	0.000
PERF-16	16.6	6.8	9.86	0.000	10.6	0.34	10.28	0.02	0.015	0.02	0.00	0.000
PERF-17	21.9	5.2	16.63	0.022	12.6	0.49	12.08	0.01	0.216	0.08	0.14	0.000
PERF-18	24.8	2.2	22.57	0.046	15.7	0.44	15.27	0.03	0.000	<b>0.00</b>	0.00	0.000
PERF-19	11.7	4.9	5.89	0.920	5.2	-0.18	5.35	0.01	0.005	<b>0.00</b>	0.00	0.000
PERFQO	20.8	7.6	13.12	0.006	5.9	<b>-0.04</b>	5.89	0.00	0.000	<b>0.00</b>	0.00	0.000
PERF-25	3.9	2.2	1.7	0.000	3.48	0.59	2.89	0.00	0.016	<b>0.00</b>	0.01	0.00
PERF-26	46.4	8.4	37.9	<b>0.107</b>	24.36	1.13	23.22	0.0125	0.103	<b>0.05</b>	0.04	0.01
PERF-27	6.1	3.4	2.7	<b>0.000</b>	2.69	0.42	2.27	0.00	0.015	<b>0.01</b>	0.00	0.00
PERF-28	4.7	2.6	2.1	0.006	6.01	1.54	4.47	0.0000	0.071	0.02	0.01	0.04
PERF-Deep2	220,140	120,782	80,496	18,863	88,178	5,331	70,619	12,228	83,880	8,751	71,202	3,927
PERF-Deep3	142,268	79,016	52,579	10,673	<b>NA<sup>a</sup></b>	NA	NA	NA	NA	NA	NA	NA
PERF-Deep4	192,785	98,182	82,188	12,415	50,321	2,528	33,407	14,386	69,514	3,828	63,158	2,529
PERF-Deep5	201,116	97,367	88,838	14,912	NA	NA	NA	NA	NA	NA	NA	NA

<sup>a</sup> NA = not analyzed.

**Table A.7 Equilibrium concentrations of WSO in open LC fractions TARO, TPOL(subtract), and TPOL(GC)**

Experiment	WSO (ppm) in TARO fraction				WSO (ppm) in TPOL fraction (subtract)				WSO (ppm) in TPOL fraction (GC)			
	Total	C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>	Total	C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>	Total	C <sub>6</sub> -C <sub>10</sub>	C <sub>10</sub> -C <sub>20</sub>	C <sub>20</sub> -C <sub>28</sub>
PERF-6	0.03	0.03	0.00	0.00	11.6	5.30	6.30	0	NA	NA	NA	NA
PERF-7	0.46	0.40	0.05	0.00	36.3	11.52	24.75	0.030	5.4	1.67	3.73	0.016
PERF-8	0.15	0.12	0.03	0.00	24.6	6.96	15.32	2.293	NA	NA	NA	NA
PERF-9	0.24	0.21	0.03	0.00	29.7	9.09	20.60	0.000	NA	NA	NA	NA
PERF-10	0.15	0.14	0.01	0.00	24.2	5.91	18.27	0.000	6.2	1.46	4.78	0.002
PERF-11	0.28	0.27	0.01	0.00	28.4	7.17	21.20	0.004	Lost	Lost	Lost	Lost
PERF-12	0.07	0.05	0.02	0.00	11.7	5.71	5.85	0.014	3.1	1.18	1.92	0.000
PERF-13	0.11	0.09	0.01	0.00	23.4	4.18	19.17	0.038	27.9	3.90	11.78	12.3
PERF-14	0.03	0.01	0.02	0.00	12.3	4.46	7.82	0.057	4.6	1.35	3.03	0.25
PERF-15	0.05	0.04	0.01	0.00	14.2	5.05	9.12	0.007	6.1	2.70	3.37	0.003
PERF-Deep1	8,438	1,098	5,726	1,614	175,999	126,826	53,954	-4,781	14,847	13,555	445	847
PERF-Std3	0.00	0.00	0.00	0.00	0.0	0.00	0.00	0.000	0.0	0.02	0.00	0.000
PERF-16	0.17	0.09	0.05	0.00	16.5	6.68	9.81	0.000	3.6	0.00	3.57	0.004
PERF-17	0.08	0.03	0.05	0.00	21.6	5.11	16.44	0.022	6.7	0.00	6.41	0.314
PERF-18	0.33	0.03	0.23	0.07	24.5	2.16	22.34	-0.028	12.6	0.00	12.46	0.145
PERF-19	0.04	0.01	0.03	0.00	11.7	4.87	5.86	0.920	5.3	0.02	5.29	0.005
PERF-20	-0.01	-0.02	0.02	0.00	20.8	7.65	13.11	0.006	1.1	0.00	1.11	0.005
PERF-25	0.370	0.20	0.17	0.00	5.0	2.16	1.92	0.90	4.07	2.65	1.42	0.004
PERF-26	0.093	0.00	0.05	0.05	46.2	8.32	37.8	0.05	3.43	2.59	0.84	0.004
PERF-27	0.024	0.00	0.01	0.01	6.05	3.41	2.64	0.00	6.25	5.44	0.78	0.027
PERF-28	0.213	0.12	0.07	0.01661	5.09	3.34	1.78	0.00	2.08	0.09	1.44	0.5529
PERF-Deep2	19,700	3,963	13,235	2,502	116,560	106,700	-61,326	2,708	19,645	18,972	673	0
PERF-Deep3	NA <sup>a</sup>	NA	NA	NA	NA	NA	NA	NA	NA	NA	NA	NA
PERF-Deep4	35,628	6,619	24,758	4,251	87,643	91,826	-14,377	-4,499	13,297	11,429	648	1,220
PERF-Deep5	NA	NA	NA	NA	NA	NA	NA	NA	NA	NA	NA	NA

<sup>a</sup> NA = not analyzed.

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