



Crystal structure of (4*R*,5*S*,6*R*)-6-azido-5-benzyloxy-3,3,4-trifluoroazepan-1-ium 2,2,2-trifluoroacetate from synchrotron data

Alpesh Ramanlal Patel,^{a,†} Mohan M. Bhadbhade^b and Fei Liu^{a,*}

Received 28 August 2015

Accepted 14 October 2015

^aDepartment of Chemistry and Biomolecular Sciences, Macquarie University, Sydney, NSW 2109, Australia, and ^bMark Wainwright Analytical Centre, The University of New South Wales, Sydney, 2052 NSW, Australia. *Correspondence e-mail: fei.liu@mq.edu.au

Edited by J. Simpson, University of Otago, New Zealand

† Current address: School of Chemistry, The University of New South Wales, Sydney NSW 2052, Australia.

Keywords: crystal structure; azepane; fluorination; fluorine *gauche* effect; hydrogen bonding

CCDC reference: 1431203

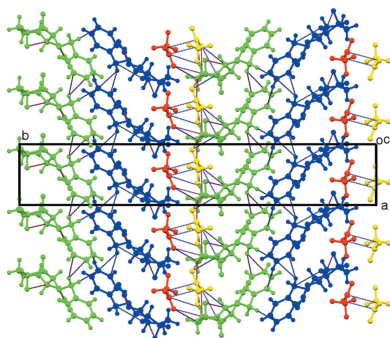
Supporting information: this article has supporting information at journals.iucr.org/e

The structure of the title compound, C₁₅H₁₆F₆N₄O₃, was determined using synchrotron radiation on an extremely small crystal (0.015 × 0.01 × 0.01 mm). Although the diffraction was weak, leading to high residuals and a poor data-to-parameter ratio, the data allowed ready solution and refinement to reveal the entire structure. The solid-state structure is in accordance with the absolute configuration assigned based on that of the known starting material. The compound comprises a highly substituted seven-membered N-heterocyclic cation and a trifluoromethanesulfonate counter-anion. The title compound crystallizes with two independent cations (*A* and *B*) and anions (*C* and *D*) in the asymmetric unit. Two geminal F atoms, a single F atom, a benzyl ether and an azide group are substituted on consecutive C atoms between the NH₂ and CH₂ units of the azepan-1-ium ring system. The seven-membered rings adopt different conformations with the principal differences occurring in the CF₂CHFCH₂ segments of the ring systems. The geminal F atoms on the quaternary C atom exhibit essentially identical bond angles [109 (2) and 106 (2)°] in the two independent molecules. The two molecules associate as a dimeric unit *via* two C—H...F interactions. An extensive series of N—H...O, N—H...F, C—H...O, C—H...N, C—H...F and C—H...π contacts generate a three-dimensional network with cations and anions linked into *ABCD* repeat columns along *a*.

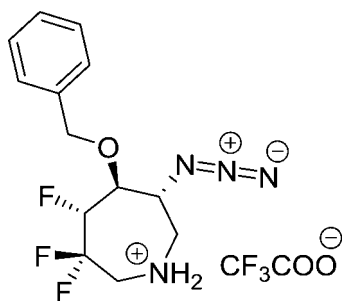
1. Chemical context

Fluorine is virtually absent in naturally occurring bioactive molecules. However, about 20% of pharmaceuticals and 30% of agrochemicals have at least one fluorine atom (Müller *et al.*, 2007; Isanbor & O'Hagan, 2006). Because fluorine is the most electronegative atom, it is small and forms very strong C—F bonds. The replacement of hydrogen by the bioisosteric fluorine in pharmacophores can lead to improved physical, chemical and biological properties (Ritter, 2012; Bégué & Bonnet-Delpon, 2006; Kirk, 2006).

Substituted azepane rings are prevalent in many bioactive natural compounds (Wipf & Spencer, 2005; Núñez-Villanueva *et al.*, 2011). Recently, substituted azepane rings and related compounds (iminocyclitols or iminosugars) have attracted considerable attention from medicinal chemists because of their great potential as glycosidase inhibitors (Stütz, 1999) and antidiabetic (Painter *et al.*, 2004), anticancer (Zitzmann *et al.*, 1999) and antiviral agents (Laver *et al.*, 1999) and are also effective against HIV (Sinnott, 1990). The conformational control of such flexible ring structures is important to their bioactivity.



We have previously reported stereospecific deoxyfluorination reactions of substituted seven-membered N-heterocycles such as azepanes (Patel & Liu, 2013, 2015; Patel *et al.*, 2013, 2014). The fluorine atoms that were added were found to regulate the conformational preferences of the N-heterocycle rings, and these fluorine-directed conformational changes were analysed by NMR techniques in solution in conjunction with computational modelling. Solution conformation analysis of the trifluorinated azepane was found to be difficult, and its direct solid-state structural analysis was also not feasible without having to add various substituents (Patel *et al.*, 2014). Incorporation of benzyloxy and azide substituents in the 5- and 6-positions of the seven-membered ring led to crystal formation. However, the crystals were extremely small ($0.015 \times 0.01 \times 0.01$ mm) and diffraction data were obtained on the title trifluorinated azepane compound, $C_{15}H_{16}F_6N_4O_3$ (**1**), directly using synchrotron radiation.



2. Structural commentary

The compound crystallizes in a chiral space group (monoclinic, $P2_1$) with two sets of cations and anions (molecule *A* and *B*) in

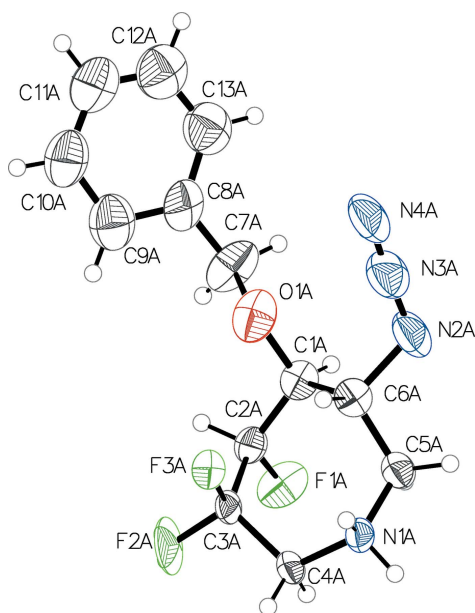


Figure 1

One of the two molecules (*A*) in the asymmetric unit, showing the atom numbering. Displacement ellipsoids are drawn at the 50% probability level.

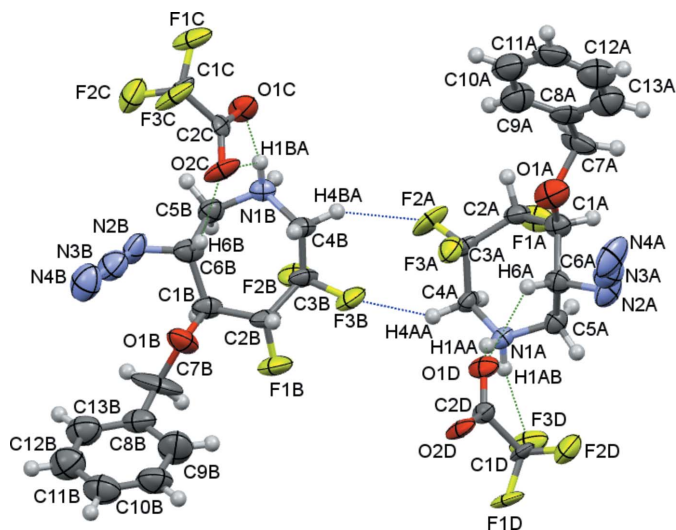


Figure 2

A view of the complete asymmetric unit consisting of two molecules of (**1**) and two trifluoromethanesulfonate anions. In this and subsequent figures, hydrogen bonds are drawn as dashed lines.

the asymmetric unit. Each cation has the same stereochemistry. An ORTEP view of the cation in molecule *A*, Fig. 1, depicts the absolute configuration and atom-labelling scheme. The *B* cation and anion are labelled similarly but with trailing *B* characters after the atom numbers. The absolute configuration was assigned based on that of the known starting material.

An alternative ORTEP view, Fig. 2, shows the asymmetric unit with association between *A* and *B* molecules *via* two C—H...F interactions to form dimers. The asymmetric unit is completed by the two triflate anions *C* and *D*. These are variously linked in an *A* to *C* and *B* to *D* fashion by N—H...O, N—H...F, C—H...O and C—H...F hydrogen bonds, Table 1.

The two molecules differ significantly in their seven-membered ring conformations, in particular around C2 and C3 with significantly different torsion angles, Fig. 3, where the molecules are involved in making dimeric contacts. Torsion angles within the two rings are shown in Fig. 3.

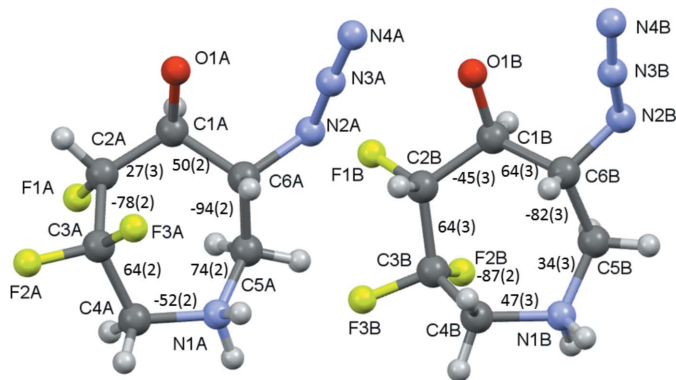


Figure 3

Conformations and torsion angles of the seven-membered rings of molecules *A* and *B*.

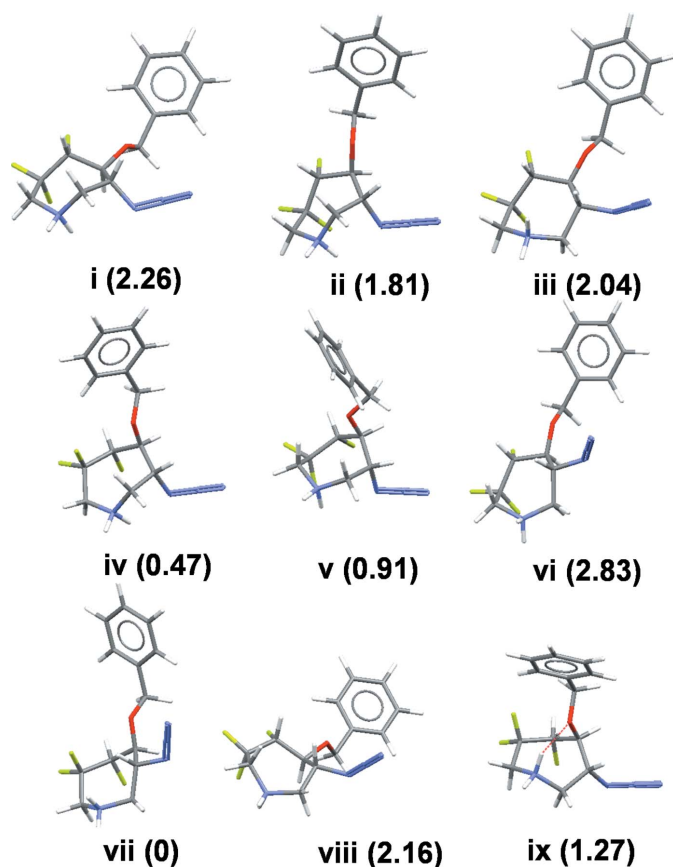


Figure 4
Nine conformations of compound (**1**) found by computational analysis. The number in parenthesis is the relative energy in kcal mol⁻¹.

3. Ring conformation analysis

A computational analysis of ring conformations of compound (**1**) was carried out using protocols reported earlier (Patel *et al.*, 2013, 2014). Conformers were first generated by the stochastic method and minimized in the MMFF94x force field with chloroform as the solvent to produce nine conformational clusters within 3–5 kcal mol⁻¹ in energy that are distinct in their azepane-ring conformations, Fig. 4. Representative conformers were then subjected to DFT geometry optimization [SV(P) basis set at the B3LYP level in COSMO solvent chloroform]. Two of the nine ring geometries (geometries vi and vii, Fig. 4) found by this computational analysis matched to geometries *A* and *B* of compound (**1**) in the unit cell, respectively. Hence the X-ray structure reported here for (**1**) validates our conformational analysis methodology as reported earlier (Patel *et al.*, 2013, 2014).

4. Supramolecular features

In the crystal structure, *C* anions form chains along the *a*-axis direction through F3C⁺⋯O1C⁻ contacts at a distance of 2.78 (2) Å. Each anion further connects to an *A* cation with O1C⁻ accepting three interactions and N1A as a bifurcated donor, leading to the formation of N1A–H1AA⋯O1C⁻, N1A–H1AB⋯O1C⁻ and C4A–H4AA⋯O1C⁻ hydrogen

Table 1

Hydrogen-bond geometry (Å, °).

Cg1 and *Cg2* are the mid-points of the C10A–C11A and C10B–C11B bonds, respectively.

<i>D</i> –H⋯ <i>A</i>	<i>D</i> –H	H⋯ <i>A</i>	<i>D</i> ⋯ <i>A</i>	<i>D</i> –H⋯ <i>A</i>
N1A–H1AA⋯O1D ⁱ	0.91	1.89	2.75 (2)	156
N1A–H1AB⋯O2D ⁱⁱ	0.91	1.79	2.677 (19)	163
N1B–H1BA⋯O1C ⁱⁱⁱ	0.91	1.76	2.66 (2)	170
N1B–H1BB⋯O2C ^{iv}	0.91	1.85	2.75 (2)	167
N1A–H1AA⋯O1C	0.91	2.57	2.88 (3)	101
N1A–H1AB⋯O1C	0.91	2.57	2.88 (3)	101
N1B–H1BB⋯O2D ^v	0.91	2.63	2.94 (3)	101
N1B–H1BA⋯O2D ^v	0.91	2.70	2.94 (3)	96
N1A–H1AB⋯F3D ⁱ	0.91	2.60	3.010 (16)	108
N1B–H1BB⋯F3C ^{iv}	0.91	2.55	2.954 (17)	108
C4A–H4AA⋯O1C	0.99	2.47	3.12 (3)	122
C6A–H6A⋯O1D ⁱ	1.00	2.42	3.14 (3)	129
C4B–H4BA⋯O2D ^v	0.99	2.39	3.08 (3)	126
C6B–H6B⋯O2C ⁱⁱⁱ	1.00	2.37	3.22 (3)	142
C12B–H12B⋯N4A ^{vi}	0.95	2.71	3.42 (4)	133
C4A–H4AA⋯F3B	0.99	2.50	3.29 (3)	137
C4A–H4AA⋯F3C ^v	0.99	2.69	3.27 (3)	118
C5A–H5AA⋯F1A	0.99	2.59	3.18 (2)	118
C5A–H5AB⋯F3D ⁱ	0.99	2.70	3.31 (2)	120
C6A–H6A⋯F3A	1.00	2.35	2.88 (2)	112
C4B–H4BA⋯F2A	0.99	2.55	3.39 (3)	142
C4B–H4BA⋯F3D	0.99	2.85	3.34 (3)	111
C5B–H5BA⋯F2B	0.99	2.51	2.95 (2)	107
C7B–H7BA⋯F1B	0.99	2.59	3.15 (5)	116
C7B–H7BA⋯Cg1 ^v	0.99	2.87	3.73 (4)	146
C7A–H7AA⋯Cg2 ^v	0.99	2.64	3.46 (4)	140

Symmetry codes: (i) *x*, *y*, *z* – 1; (ii) *x* + 1, *y*, *z* – 1; (iii) *x*, *y*, *z* + 1; (iv) *x* + 1, *y*, *z* + 1; (v) *x* + 1, *y*, *z*; (vi) *–x*, *y* + $\frac{1}{2}$, *–z* + 1.

bonds and generating *R*₂¹(4) and *R*₂¹(5) ring motifs, respectively (Bernstein *et al.*, 1995). These contacts generate columns of *A* molecules along *a*. These columns are further supported by weak C7A–H7AA⋯Cg2 contacts (*Cg2* is the mid-point of the C10A–C11A bond of the C8A–C13A phenyl ring), Fig. 5. Similarly, *B* cations are linked to *D* anions with O2D⁻ accepting three interactions and forming N1B–H1BA⋯O2D⁻, N1B–H1BB⋯O2D⁻ and C4B–H4BA⋯O2D⁻ hydrogen bonds. Unlike the *AC* system however, a C4B–H4BA⋯F1D⁻ hydrogen bond completes the *B*⋯*D* cation–anion contacts.

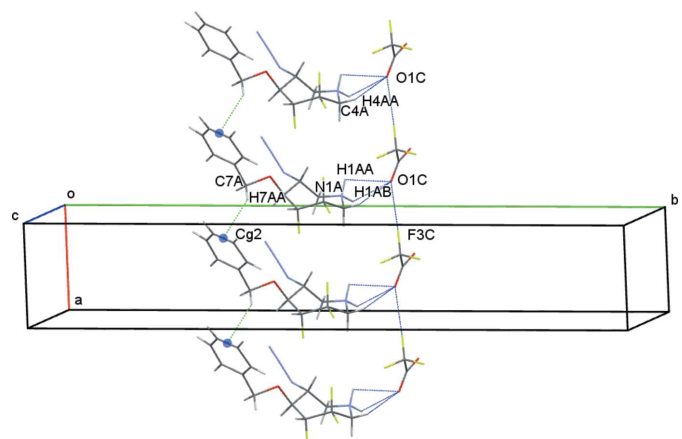


Figure 5
Intermolecular contacts between *A* cations and *C* anions viewed along *c*. Midpoints of the C10A–C11A bonds are shown as coloured spheres.

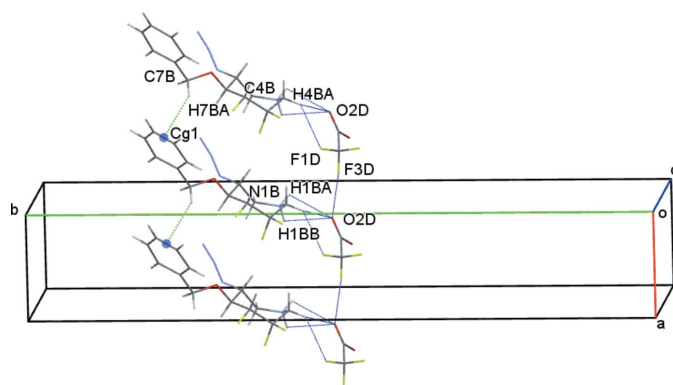


Figure 6
Intermolecular contacts between *B* cations and *D* anions viewed along *c*. Midpoints of the C10B–C11B bonds are shown as coloured spheres.

These generate $R_2^1(4)$ and $R_2^2(5)$ ring motifs respectively. Weak C7B–H7BA...Cg1 contacts (Cg1 is the midpoint of the C10B–C11B bond of the C8B–C13B phenyl ring) link adjacent *B* molecules, also forming columns of *B* cations and *D* anions along the *a*-axis direction, Fig. 6. Contacts between the *A* and *B* cations are limited to very weak C12B–H12B...N4A hydrogen bonds linking adjacent columns of *A* and *B* cations, Fig. 7. This eclectic mixture of contacts generates columns with an *ABCD* repeat unit in the direction of the *a* axis, Fig. 8. Additional N–H...O, C–H...O and C–H...F contacts result in a three-dimensional network of cations and anions stacked along *c*.

5. Database survey

A survey of the Cambridge Structural Database (Version 5.36, with three updates) (Groom & Allen, 2014) reveals the crystal structures of 11 unsubstituted azepanium (hexamethylenimine) cations with a variety of counter-anions, see for example: Verlooy *et al.* (2010); Bakshi *et al.* (1994); Moritani *et al.* (1987); Kashino *et al.* (1981); Cameron & Scheeren (1977). Two of these salts also form co-crystals, Moritani & Kashino (2002); Misaki *et al.* (1989). However the structure of (3*R*,4*R*,5*S*,6*S*)-4,5,6-trihydroxy-3-methyl azepanium chloride

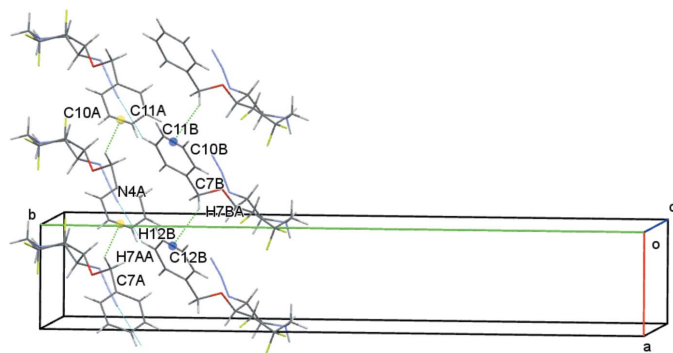


Figure 7
Intermolecular contacts between the *A* and *B* cations viewed along *c*. Mid-points of the C10A–C11A and C10B–C11B bonds are shown as coloured spheres.

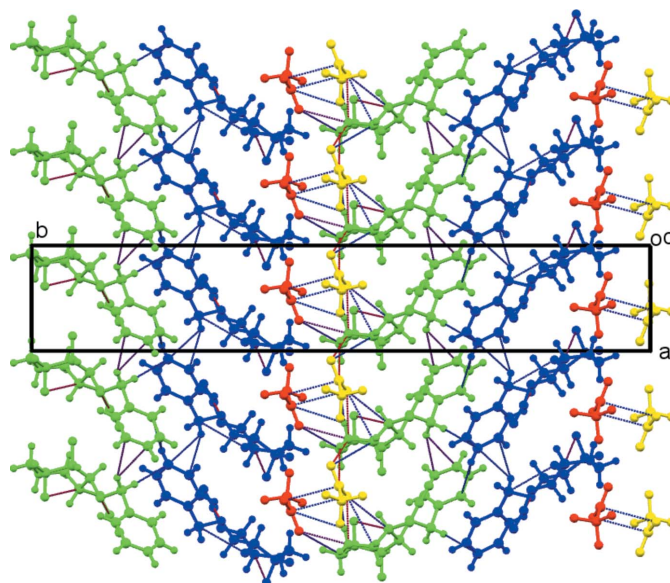


Figure 8
Packing of molecules in the unit cell viewed along *c*. Molecules *A* (green) and *B* (blue), trifluoromethanesulfonate anions *C* (red) and *D* (yellow). Hydrogen-bonding contacts are shown as dashed lines.

is the only one to be reported of a substituted azepanium salt, Li *et al.* (2008), highlighting the novelty of the present report.

6. Synthesis and crystallization

(4*R*,5*S*,6*R*)-6-Azido-5-benzyloxy-3,3,4-trifluoroazepane-1-carboxylic acid-*tert*-butyl ester (10 mg, 25.0 μ mol) was dissolved in trifluoroacetic acid (TFA, 500 μ L) at 298 K. The solution was allowed to stir for 5 min before the TFA was evaporated under an N_2 flow. The reaction flask was kept under high vacuum (0.005 torr, 298 K) for 3 h to remove traces of TFA. A colorless, oily residue was obtained which was recrystallized from dichloromethane to give colorless needles characterized as (**1**) (10.0 mg, 97%). 1H NMR (600 MHz, $CDCl_3$) δ 7.44–7.34 (*m*, 5H), 4.93 (*dd*, $J = 44.19$ ($^1J_{HF}$), 14.7 Hz, 1H), 4.80 (*d*, $J = 11.44$ Hz, 1H), 4.73 (*d*, $J = 11.44$ Hz, 1H), 4.08 (*dd*, $J = 8.71, 8.68$ Hz, 1H), 3.89–3.82 (*m*, 1H), 3.67–3.57 (*m*, 2H), 3.48 (*d*, $J = 14.0$ Hz, 1H), 3.10 (*dd*, $J = 14.0, 9.70$ Hz, 1H); ^{13}C NMR (150 MHz, $CDCl_3$) δ 135.7, 129.0, 128.5, 128.5, 118.4 (*dd*, $^1J_{CF} = 247.66$ Hz, $^2J_{CF} = 28.07$ Hz), 90.2 (*ddd*, $^1J_{CF} = 186.03$ Hz, $^2J_{CF} = 34.98$ Hz, $^3J_{CF} = 27.82$ Hz), 79.6 (*dd*, $^2J_{CF} = 24.93$ Hz, $^3J_{CF} = 7.20$ Hz), 73.9, 60.6, 45.8 (*dd*, $^2J_{CF} = 39.76$ Hz, $^2J_{CF} = 25.66$ Hz), 45.6.

7. Refinement

Crystal data, data collection and structure refinement details are summarized in Table 2. All H atoms were refined using a riding model, with N–H = 0.91 Å, C–H = 0.95 Å for aromatic, 1.00 Å for methine and 0.99 Å for methylene, all with $U_{iso}(H) = 1.2U_{eq}(N/C)$. Because of the lower reflections-to-parameter ratio, anisotropic displacement parameters of several atoms in the least-squares refinement had to be

Table 2

Experimental details.

Crystal data	
Chemical formula	$C_{13}H_{16}F_3N_4O^+ \cdot C_2F_3O_2^-$
M_r	414.32
Crystal system, space group	Monoclinic, $P2_1$
Temperature (K)	100
a, b, c (Å)	5.8780 (12), 34.503 (7), 8.8120 (18)
β (°)	92.42 (3)
V (Å ³)	1785.6 (6)
Z	4
Radiation type	Synchrotron, $\lambda = 0.7293$ Å
μ (mm ⁻¹)	0.16
Crystal size (mm)	0.015 × 0.01 × 0.01
Data collection	
Diffractometer	Bruker APEXII CCD
No. of measured, independent and observed [$I > 2\sigma(I)$] reflections	13709, 3642, 2175
R_{int}	0.386
θ_{max} (°)	21.5
$(\sin \theta/\lambda)_{max}$ (Å ⁻¹)	0.502
Refinement	
$R[F^2 > 2\sigma(F^2)]$, $wR(F^2)$, S	0.116, 0.261, 0.97
No. of reflections	3642
No. of parameters	505
No. of restraints	193
H-atom treatment	H-atom parameters constrained
$\Delta\rho_{max}$, $\Delta\rho_{min}$ (e Å ⁻³)	0.56, -0.41

Computer programs: *XDS* (Kabsch, 2010), *SHELXS97* (Sheldrick, 2008), *SHELXL2014* (Sheldrick, 2015), *OLEX2* (Dolomanov *et al.*, 2009) and *Mercury* (Macrae *et al.*, 2008).

restrained using the RIGU command. These were applied to azide groups, atoms in the seven-membered and a few atoms in phenyl rings.

Acknowledgements

The authors thank the Australian Synchrotron facility for the diffraction data. The support from Dr Luke Hunter and Dr Samuel Kutty is greatly appreciated.

References

- Bakshi, P. K., Linden, A., Vincent, B. R., Roe, S. P., Adhikesavalu, D., Cameron, T. S. & Knop, O. (1994). *Can. J. Chem.* **72**, 1273–1293.
 Bégué, J.-P. & Bonnet-Delpon, D. (2006). *J. Fluor. Chem.* **127**, 992–1012.

- Bernstein, J., Davis, R. E., Shimon, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
 Cameron, T. S. & Scheeren, H. W. (1977). *J. Chem. Soc. Chem. Commun.* pp. 939–941.
 Dolomanov, O. V., Bourhis, L. J., Gildea, R. J., Howard, J. A. K. & Puschmann, H. (2009). *J. Appl. Cryst.* **42**, 339–341.
 Groom, C. R. & Allen, F. H. (2014). *Angew. Chem. Int. Ed.* **53**, 662–671.
 Isanbor, C. & O'Hagan, D. (2006). *J. Fluor. Chem.* **127**, 303–319.
 Kabsch, W. (2010). *Acta Cryst.* **D66**, 125–132.
 Kashino, S., Sasahara, N., Kataoka, S.-I. & Haisa, M. (1981). *Bull. Chem. Soc. Jpn.* **54**, 962–966.
 Kirk, K. L. (2006). *J. Fluor. Chem.* **127**, 1013–1029.
 Laver, W. G., Bischofberger, N. & Webster, R. G. (1999). *Sci. Am.* **280**, 78–87.
 Li, H., Liu, T., Zhang, Y., Favre, S., Bello, C., Vogel, P., Butters, T. D., Oikonomakos, N. G., Marrot, J. & Blériot, Y. (2008). *ChemBioChem*, **9**, 253–260.
 Macrae, C. F., Bruno, I. J., Chisholm, J. A., Edgington, P. R., McCabe, P., Pidcock, E., Rodriguez-Monge, L., Taylor, R., van de Streek, J. & Wood, P. A. (2008). *J. Appl. Cryst.* **41**, 466–470.
 Misaki, S., Kashino, S. & Haisa, M. (1989). *Acta Cryst.* **C45**, 917–921.
 Moritani, Y. & Kashino, S. (2002). *Bull. Chem. Soc. Jpn.* **75**, 1235–1239.
 Moritani, Y., Sasahara, N., Kashino, S. & Haisa, M. (1987). *Acta Cryst.* **C43**, 154–158.
 Müller, K., Faeh, C. & Diederich, F. (2007). *Science*, **317**, 1881–1886.
 Núñez-Villanueva, D., Bonache, M. A. N., Infantes, L., García-López, M. T., Martín-Martínez, M. & González-Muñiz, R. (2011). *J. Org. Chem.* **76**, 6592–6603.
 Painter, G. F., Eldridge, P. J. & Falshaw, A. (2004). *Bioorg. Med. Chem.* **12**, 225–232.
 Patel, A. R., Ball, G., Hunter, L. & Liu, F. (2013). *Org. Biomol. Chem.* **11**, 3781–3785.
 Patel, A. R., Hunter, L., Bhadbhade, M. M. & Liu, F. (2014). *Eur. J. Org. Chem.* pp. 2584–2593.
 Patel, A. R. & Liu, F. (2013). *Tetrahedron*, **69**, 744–752.
 Patel, A. R. & Liu, F. (2015). *Aust. J. Chem.* **68**, 50–56.
 Ritter, S. K. (2012). *Chem. Eng. News*, **90**, 10–17.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Sheldrick, G. M. (2015). *Acta Cryst.* **C71**, 3–8.
 Sinnott, M. L. (1990). *Chem. Rev.* **90**, 1171–1202.
 Stütz, A. E. (1999). In *Iminosugars as Glycosidase Inhibitors: Nojirimycin and Beyond*. Weinheim: Wiley-VCH.
 Verlooy, P. L. H., Robeyns, K., Van Meervelt, L., Lebedev, O. I., Van Tendeloo, G., Martens, J. A. & Kirschhock, C. E. A. (2010). *Microporous Mesoporous Mater.* **130**, 14–20.
 Wipf, P. & Spencer, S. R. (2005). *J. Am. Chem. Soc.* **127**, 225–235.
 Zitzmann, N., Mehta, A. S., Carrouée, S., Butters, T. D., Platt, F. M., McCauley, J., Blumberg, B. S., Dwek, R. A. & Block, T. M. (1999). *Proc. Natl Acad. Sci.* **96**, 11878–11882.

supporting information

Acta Cryst. (2015). E71, 1361-1365 [https://doi.org/10.1107/S2056989015019416]

Crystal structure of (4*R*,5*S*,6*R*)-6-azido-5-benzyloxy-3,3,4-trifluoroazepan-1-ium 2,2,2-trifluoroacetate from synchrotron data

Alpesh Ramanlal Patel, Mohan M. Bhadbhade and Fei Liu

Computing details

Data collection: *XDS* (Kabsch, 2010); cell refinement: *XDS* (Kabsch, 2010); data reduction: *XDS* (Kabsch, 2010); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL2014* (Sheldrick, 2015); molecular graphics: *OLEX2* (Dolomanov *et al.*, 2009) and *Mercury* (Macrae *et al.*, 2008); software used to prepare material for publication: *OLEX2* (Dolomanov *et al.*, 2009).

(4*R*,5*S*,6*R*)-6-Azido-5-benzyloxy-3,3,4-trifluoroazepan-1-ium 2,2,2-trifluoroacetate

Crystal data

$C_{13}H_{16}F_3N_4O^+ \cdot C_2F_3O_2^-$

$M_r = 414.32$

Monoclinic, $P2_1$

$a = 5.8780$ (12) Å

$b = 34.503$ (7) Å

$c = 8.8120$ (18) Å

$\beta = 92.42$ (3)°

$V = 1785.6$ (6) Å³

$Z = 4$

$F(000) = 848$

$D_x = 1.541$ Mg m⁻³

Synchrotron radiation, $\lambda = 0.7293$ Å

$\mu = 0.16$ mm⁻¹

$T = 100$ K

Plate, colourless

$0.02 \times 0.01 \times 0.01$ mm

Data collection

Bruker APEXII CCD

diffractometer

ω scans

13709 measured reflections

3642 independent reflections

2175 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.386$

$\theta_{\text{max}} = 21.5^\circ$, $\theta_{\text{min}} = 2.4^\circ$

$h = -5 \rightarrow 5$

$k = -34 \rightarrow 34$

$l = -8 \rightarrow 8$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.116$

$wR(F^2) = 0.261$

$S = 0.97$

3642 reflections

505 parameters

193 restraints

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0001P)^2]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\text{max}} < 0.001$

$\Delta\rho_{\text{max}} = 0.56$ e Å⁻³

$\Delta\rho_{\text{min}} = -0.41$ e Å⁻³

Absolute structure: Flack x determined using

390 quotients $[(F^-)-(F)]/[(F^+)+(F)]$ (Parsons *et al.*, 2013)

Absolute structure parameter: 2.2 (10)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1A	0.606 (3)	0.3889 (6)	0.6446 (15)	0.067 (5)
F1A	1.1433 (19)	0.4338 (5)	0.6460 (13)	0.062 (4)
F2A	0.923 (2)	0.4956 (5)	0.7835 (11)	0.080 (5)
F3A	0.615 (2)	0.4787 (4)	0.6578 (10)	0.050 (3)
N1A	0.835 (3)	0.4880 (5)	0.3749 (14)	0.035 (4)
H1AA	0.6842	0.4943	0.3686	0.042*
H1AB	0.9039	0.5018	0.3020	0.042*
N2A	0.627 (4)	0.3894 (7)	0.3202 (19)	0.066 (6)
N3A	0.444 (5)	0.3736 (8)	0.351 (2)	0.069 (7)
N4A	0.278 (4)	0.3579 (9)	0.373 (2)	0.090 (9)
C1A	0.794 (3)	0.4058 (7)	0.571 (2)	0.042 (5)
H1A	0.9033	0.3848	0.5454	0.051*
C2A	0.913 (4)	0.4343 (7)	0.668 (2)	0.044 (5)
H2A	0.8938	0.4258	0.7751	0.053*
C3A	0.844 (3)	0.4777 (7)	0.6604 (17)	0.034 (4)
C4A	0.928 (3)	0.5013 (7)	0.5242 (17)	0.039 (5)
H4AA	0.8869	0.5289	0.5379	0.046*
H4AB	1.0967	0.4998	0.5247	0.046*
C5A	0.855 (3)	0.4469 (7)	0.337 (2)	0.042 (5)
H5AA	1.0121	0.4381	0.3625	0.050*
H5AB	0.8243	0.4433	0.2270	0.050*
C6A	0.691 (3)	0.4228 (7)	0.423 (2)	0.043 (5)
H6A	0.5523	0.4384	0.4446	0.051*
C7A	0.678 (4)	0.3547 (9)	0.737 (3)	0.070 (8)
H7AA	0.7974	0.3618	0.8142	0.084*
H7AB	0.7359	0.3339	0.6716	0.084*
C8A	0.465 (4)	0.3422 (9)	0.810 (2)	0.062 (6)
C9A	0.373 (5)	0.3642 (11)	0.921 (3)	0.081 (7)
H9A	0.4346	0.3887	0.9491	0.097*
C10A	0.178 (5)	0.3483 (11)	0.994 (3)	0.082 (7)
H10A	0.1049	0.3630	1.0692	0.098*
C11A	0.099 (5)	0.3121 (10)	0.955 (3)	0.079 (7)
H11A	−0.0217	0.3010	1.0091	0.095*
C12A	0.182 (5)	0.2932 (11)	0.848 (3)	0.084 (7)
H12A	0.1102	0.2699	0.8148	0.101*
C13A	0.375 (5)	0.3053 (11)	0.778 (3)	0.080 (7)
H13A	0.4478	0.2887	0.7091	0.096*
O1B	0.613 (3)	0.6932 (5)	0.8311 (16)	0.060 (4)
F1B	0.916 (3)	0.6548 (5)	0.6524 (12)	0.079 (5)

F2B	1.1956 (19)	0.6200 (4)	0.8697 (11)	0.055 (4)
F3B	1.005 (2)	0.5828 (5)	0.7138 (11)	0.067 (4)
N1B	0.961 (3)	0.5980 (6)	1.1146 (16)	0.036 (4)
H1BA	0.8766	0.5850	1.1825	0.044*
H1BB	1.1099	0.5922	1.1364	0.044*
N2B	0.682 (3)	0.6932 (7)	1.1396 (17)	0.060 (6)
N3B	0.485 (4)	0.7051 (7)	1.1158 (18)	0.067 (6)
N4B	0.313 (4)	0.7214 (9)	1.101 (2)	0.081 (8)
C1B	0.792 (4)	0.6723 (7)	0.891 (2)	0.045 (5)
H1B	0.9306	0.6892	0.8960	0.054*
C2B	0.830 (4)	0.6396 (7)	0.7792 (19)	0.046 (5)
H2B	0.6767	0.6285	0.7498	0.055*
C3B	0.983 (4)	0.6055 (7)	0.836 (2)	0.045 (5)
C4B	0.898 (4)	0.5830 (8)	0.9578 (18)	0.044 (5)
H4BA	0.9560	0.5562	0.9494	0.052*
H4BB	0.7300	0.5819	0.9457	0.052*
C5B	0.930 (4)	0.6390 (8)	1.137 (2)	0.048 (5)
H5BA	1.0715	0.6520	1.1079	0.058*
H5BB	0.9160	0.6433	1.2471	0.058*
C6B	0.734 (4)	0.6602 (8)	1.057 (2)	0.050 (5)
H6B	0.5977	0.6428	1.0517	0.060*
C7B	0.666 (5)	0.7294 (11)	0.765 (4)	0.098 (12)
H7BA	0.7893	0.7257	0.6933	0.117*
H7BB	0.7228	0.7473	0.8460	0.117*
C8B	0.471 (4)	0.7466 (9)	0.686 (2)	0.064 (6)
C9B	0.369 (5)	0.7277 (11)	0.557 (3)	0.090 (9)
H9B	0.4288	0.7044	0.5175	0.109*
C10B	0.176 (5)	0.7452 (10)	0.492 (3)	0.080 (8)
H10B	0.1056	0.7328	0.4058	0.097*
C11B	0.078 (5)	0.7791 (10)	0.541 (3)	0.078 (7)
H11B	−0.0543	0.7895	0.4908	0.094*
C12B	0.180 (4)	0.7969 (11)	0.665 (3)	0.082 (8)
H12B	0.1141	0.8200	0.7020	0.098*
C13B	0.375 (5)	0.7826 (10)	0.737 (3)	0.076 (7)
H13B	0.4460	0.7963	0.8204	0.092*
O1C	0.709 (3)	0.5680 (5)	0.3297 (14)	0.052 (5)
O2C	0.420 (2)	0.5922 (6)	0.1855 (13)	0.055 (5)
F1C	0.440 (2)	0.5625 (5)	0.5606 (11)	0.074 (5)
F2C	0.414 (2)	0.6230 (5)	0.5063 (14)	0.077 (5)
F3C	0.1583 (19)	0.5832 (5)	0.4231 (9)	0.071 (5)
C1C	0.382 (3)	0.5872 (8)	0.447 (2)	0.042 (6)
C2C	0.519 (4)	0.5805 (6)	0.306 (2)	0.033 (5)
O1D	0.378 (2)	0.4904 (5)	1.2965 (13)	0.047 (4)
O2D	0.090 (2)	0.5162 (5)	1.1558 (11)	0.045 (4)
F1D	0.366 (2)	0.5249 (4)	0.9285 (10)	0.054 (3)
F2D	0.375 (2)	0.4654 (5)	0.9654 (12)	0.071 (4)
F3D	0.638 (2)	0.5015 (5)	1.0584 (11)	0.066 (4)
C1D	0.420 (3)	0.4990 (8)	1.0293 (18)	0.037 (4)

C2D 0.281 (4) 0.5034 (7) 1.1753 (19) 0.036 (6)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1A	0.045 (10)	0.100 (17)	0.058 (9)	0.007 (10)	0.001 (7)	0.006 (10)
F1A	0.035 (8)	0.079 (12)	0.070 (7)	0.005 (7)	−0.004 (5)	0.024 (8)
F2A	0.097 (10)	0.120 (15)	0.022 (6)	−0.010 (10)	−0.007 (5)	−0.020 (7)
F3A	0.060 (8)	0.051 (10)	0.042 (6)	−0.004 (7)	0.013 (5)	−0.006 (6)
N1A	0.036 (9)	0.043 (10)	0.026 (6)	−0.004 (7)	0.002 (5)	−0.004 (6)
N2A	0.077 (14)	0.070 (17)	0.050 (10)	−0.010 (12)	−0.007 (9)	−0.024 (11)
N3A	0.078 (15)	0.075 (18)	0.053 (11)	−0.007 (12)	−0.012 (10)	−0.010 (11)
N4A	0.090 (16)	0.11 (2)	0.068 (12)	−0.036 (14)	0.005 (11)	−0.046 (13)
C1A	0.033 (12)	0.050 (11)	0.044 (8)	0.005 (8)	−0.002 (7)	0.004 (7)
C2A	0.058 (13)	0.046 (10)	0.028 (8)	−0.006 (8)	−0.001 (8)	0.007 (7)
C3A	0.035 (12)	0.047 (10)	0.020 (8)	−0.009 (8)	−0.003 (7)	−0.005 (7)
C4A	0.055 (12)	0.036 (11)	0.024 (6)	−0.008 (9)	0.000 (6)	−0.006 (7)
C5A	0.046 (11)	0.043 (10)	0.037 (9)	−0.002 (8)	−0.002 (7)	−0.006 (7)
C6A	0.042 (11)	0.043 (12)	0.043 (8)	0.006 (8)	−0.004 (7)	0.003 (8)
C7A	0.067 (17)	0.07 (2)	0.077 (15)	0.031 (16)	0.012 (12)	0.042 (16)
C8A	0.075 (14)	0.075 (16)	0.034 (10)	0.012 (11)	−0.018 (9)	0.019 (9)
C9A	0.091 (16)	0.091 (17)	0.060 (13)	0.011 (12)	−0.005 (10)	0.002 (11)
C10A	0.094 (16)	0.098 (17)	0.053 (12)	0.011 (12)	0.000 (10)	0.013 (12)
C11A	0.093 (15)	0.088 (17)	0.056 (12)	0.016 (12)	−0.014 (10)	0.032 (11)
C12A	0.086 (16)	0.097 (17)	0.070 (14)	0.002 (12)	−0.012 (10)	0.013 (12)
C13A	0.090 (16)	0.084 (17)	0.066 (13)	0.002 (11)	−0.008 (10)	0.004 (11)
O1B	0.068 (10)	0.048 (13)	0.062 (8)	0.003 (9)	−0.020 (7)	0.015 (9)
F1B	0.128 (12)	0.080 (13)	0.030 (6)	−0.003 (10)	0.007 (7)	0.019 (7)
F2B	0.045 (8)	0.072 (11)	0.049 (6)	−0.009 (7)	0.012 (5)	0.017 (7)
F3B	0.100 (11)	0.077 (12)	0.027 (6)	−0.002 (9)	0.018 (6)	−0.001 (7)
N1B	0.030 (9)	0.043 (10)	0.036 (7)	−0.004 (7)	−0.003 (6)	0.004 (6)
N2B	0.067 (11)	0.072 (17)	0.040 (9)	0.004 (11)	−0.004 (9)	−0.023 (10)
N3B	0.062 (12)	0.085 (19)	0.053 (10)	0.007 (11)	0.008 (10)	−0.006 (11)
N4B	0.067 (12)	0.10 (2)	0.078 (13)	0.017 (12)	0.000 (10)	−0.019 (14)
C1B	0.044 (12)	0.052 (12)	0.038 (8)	−0.014 (9)	−0.013 (7)	0.008 (7)
C2B	0.058 (12)	0.050 (11)	0.028 (8)	−0.022 (8)	−0.005 (7)	0.014 (7)
C3B	0.060 (12)	0.051 (12)	0.024 (8)	−0.017 (9)	−0.012 (7)	0.007 (8)
C4B	0.052 (12)	0.047 (11)	0.031 (7)	−0.005 (9)	−0.003 (7)	0.010 (7)
C5B	0.056 (11)	0.050 (10)	0.037 (9)	−0.004 (9)	−0.008 (8)	−0.001 (8)
C6B	0.051 (11)	0.057 (14)	0.042 (8)	−0.008 (9)	−0.006 (7)	0.008 (8)
C7B	0.09 (2)	0.09 (3)	0.11 (2)	−0.034 (19)	−0.037 (16)	0.08 (2)
C8B	0.065 (13)	0.081 (17)	0.046 (10)	0.005 (11)	0.004 (9)	0.020 (10)
C9B	0.098 (15)	0.098 (18)	0.073 (13)	0.031 (13)	−0.027 (11)	−0.004 (12)
C10B	0.091 (15)	0.092 (18)	0.057 (12)	0.018 (12)	−0.010 (10)	0.019 (11)
C11B	0.072 (15)	0.087 (18)	0.075 (13)	0.008 (12)	−0.007 (10)	0.018 (11)
C12B	0.069 (14)	0.100 (18)	0.076 (13)	0.013 (12)	−0.003 (10)	0.004 (12)
C13B	0.071 (14)	0.093 (18)	0.064 (12)	0.008 (11)	−0.003 (10)	0.008 (11)
O1C	0.036 (9)	0.077 (14)	0.043 (7)	0.014 (9)	−0.007 (6)	−0.002 (8)

O2C	0.035 (8)	0.108 (16)	0.022 (7)	−0.011 (9)	−0.005 (6)	0.003 (8)
F1C	0.063 (8)	0.130 (16)	0.028 (6)	0.010 (9)	0.010 (5)	0.020 (8)
F2C	0.085 (10)	0.091 (15)	0.055 (7)	0.010 (10)	0.010 (6)	−0.027 (9)
F3C	0.044 (9)	0.151 (17)	0.019 (5)	0.010 (8)	0.003 (5)	−0.004 (7)
C1C	0.028 (14)	0.08 (2)	0.020 (10)	−0.002 (12)	−0.002 (9)	−0.012 (13)
C2C	0.047 (15)	0.016 (14)	0.038 (12)	−0.010 (11)	0.004 (11)	−0.002 (10)
O1D	0.035 (8)	0.070 (13)	0.037 (7)	0.006 (8)	0.013 (6)	0.012 (8)
O2D	0.028 (9)	0.087 (14)	0.020 (6)	0.020 (8)	0.009 (5)	0.006 (7)
F1D	0.060 (7)	0.089 (9)	0.013 (5)	0.004 (6)	0.003 (4)	0.016 (5)
F2D	0.099 (10)	0.078 (10)	0.038 (6)	0.001 (7)	0.015 (6)	−0.016 (6)
F3D	0.040 (6)	0.116 (13)	0.043 (6)	0.008 (6)	0.003 (5)	0.019 (7)
C1D	0.039 (9)	0.059 (10)	0.012 (8)	0.006 (7)	0.000 (7)	−0.001 (7)
C2D	0.030 (13)	0.050 (17)	0.027 (11)	0.012 (12)	0.000 (9)	−0.001 (10)

Geometric parameters (Å, °)

O1A—C1A	1.43 (3)	N1B—H1BB	0.9100
O1A—C7A	1.48 (3)	N1B—C4B	1.51 (2)
F1A—C2A	1.37 (2)	N1B—C5B	1.44 (3)
F2A—C3A	1.31 (2)	N2B—N3B	1.24 (3)
F3A—C3A	1.34 (2)	N2B—C6B	1.39 (3)
N1A—H1AA	0.9100	N3B—N4B	1.16 (3)
N1A—H1AB	0.9100	C1B—H1B	1.0000
N1A—C4A	1.48 (2)	C1B—C2B	1.52 (3)
N1A—C5A	1.46 (3)	C1B—C6B	1.57 (3)
N2A—N3A	1.25 (3)	C2B—H2B	1.0000
N2A—C6A	1.50 (3)	C2B—C3B	1.55 (3)
N3A—N4A	1.14 (3)	C3B—C4B	1.43 (3)
C1A—H1A	1.0000	C4B—H4BA	0.9900
C1A—C2A	1.46 (3)	C4B—H4BB	0.9900
C1A—C6A	1.53 (3)	C5B—H5BA	0.9900
C2A—H2A	1.0000	C5B—H5BB	0.9900
C2A—C3A	1.55 (3)	C5B—C6B	1.51 (3)
C3A—C4A	1.55 (3)	C6B—H6B	1.0000
C4A—H4AA	0.9900	C7B—H7BA	0.9900
C4A—H4AB	0.9900	C7B—H7BB	0.9900
C5A—H5AA	0.9900	C7B—C8B	1.45 (4)
C5A—H5AB	0.9900	C8B—C9B	1.42 (4)
C5A—C6A	1.50 (3)	C8B—C13B	1.45 (4)
C6A—H6A	1.0000	C9B—H9B	0.9500
C7A—H7AA	0.9900	C9B—C10B	1.38 (4)
C7A—H7AB	0.9900	C10B—H10B	0.9500
C7A—C8A	1.50 (4)	C10B—C11B	1.38 (4)
C8A—C9A	1.37 (4)	C11B—H11B	0.9500
C8A—C13A	1.40 (4)	C11B—C12B	1.37 (4)
C9A—H9A	0.9500	C12B—H12B	0.9500
C9A—C10A	1.44 (4)	C12B—C13B	1.38 (4)
C10A—H10A	0.9500	C13B—H13B	0.9500

C10A—C11A	1.37 (4)	O1C—C2C	1.21 (2)
C11A—H11A	0.9500	O2C—C2C	1.26 (2)
C11A—C12A	1.26 (4)	F1C—C1C	1.35 (3)
C12A—H12A	0.9500	F2C—C1C	1.35 (3)
C12A—C13A	1.38 (4)	F3C—C1C	1.33 (2)
C13A—H13A	0.9500	C1C—C2C	1.53 (3)
O1B—C1B	1.36 (3)	O1D—C2D	1.27 (2)
O1B—C7B	1.42 (3)	O2D—C2D	1.21 (2)
F1B—C2B	1.35 (2)	F1D—C1D	1.29 (3)
F2B—C3B	1.37 (2)	F2D—C1D	1.31 (3)
F3B—C3B	1.34 (2)	F3D—C1D	1.30 (2)
N1B—H1BA	0.9100	C1D—C2D	1.56 (3)
C1A—O1A—C7A	111.5 (17)	O1B—C1B—H1B	108.8
H1AA—N1A—H1AB	107.1	O1B—C1B—C2B	105.9 (15)
C4A—N1A—H1AA	107.8	O1B—C1B—C6B	107.8 (18)
C4A—N1A—H1AB	107.8	C2B—C1B—H1B	108.8
C5A—N1A—H1AA	107.8	C2B—C1B—C6B	116.6 (19)
C5A—N1A—H1AB	107.8	C6B—C1B—H1B	108.8
C5A—N1A—C4A	118.2 (17)	F1B—C2B—C1B	108.4 (18)
N3A—N2A—C6A	113.7 (18)	F1B—C2B—H2B	107.3
N4A—N3A—N2A	177 (3)	F1B—C2B—C3B	108.9 (18)
O1A—C1A—H1A	108.8	C1B—C2B—H2B	107.3
O1A—C1A—C2A	111.7 (16)	C1B—C2B—C3B	117.1 (15)
O1A—C1A—C6A	104.8 (15)	C3B—C2B—H2B	107.3
C2A—C1A—H1A	108.8	F2B—C3B—C2B	107.8 (18)
C2A—C1A—C6A	113.8 (19)	F2B—C3B—C4B	112.1 (15)
C6A—C1A—H1A	108.8	F3B—C3B—F2B	105.8 (18)
F1A—C2A—C1A	111.1 (18)	F3B—C3B—C2B	105.2 (14)
F1A—C2A—H2A	106.5	F3B—C3B—C4B	110 (2)
F1A—C2A—C3A	105.3 (17)	C4B—C3B—C2B	116 (2)
C1A—C2A—H2A	106.5	N1B—C4B—H4BA	108.5
C1A—C2A—C3A	120.3 (17)	N1B—C4B—H4BB	108.5
C3A—C2A—H2A	106.5	C3B—C4B—N1B	115 (2)
F2A—C3A—F3A	108.7 (15)	C3B—C4B—H4BA	108.5
F2A—C3A—C2A	109.5 (16)	C3B—C4B—H4BB	108.5
F2A—C3A—C4A	106.2 (18)	H4BA—C4B—H4BB	107.5
F3A—C3A—C2A	106.6 (17)	N1B—C5B—H5BA	107.1
F3A—C3A—C4A	109.1 (15)	N1B—C5B—H5BB	107.1
C4A—C3A—C2A	116.6 (17)	N1B—C5B—C6B	120.7 (19)
N1A—C4A—C3A	114.0 (17)	H5BA—C5B—H5BB	106.8
N1A—C4A—H4AA	108.8	C6B—C5B—H5BA	107.1
N1A—C4A—H4AB	108.8	C6B—C5B—H5BB	107.1
C3A—C4A—H4AA	108.8	N2B—C6B—C1B	109 (2)
C3A—C4A—H4AB	108.8	N2B—C6B—C5B	109.3 (18)
H4AA—C4A—H4AB	107.7	N2B—C6B—H6B	108.8
N1A—C5A—H5AA	109.4	C1B—C6B—H6B	108.8
N1A—C5A—H5AB	109.4	C5B—C6B—C1B	111.8 (19)

N1A—C5A—C6A	111.3 (16)	C5B—C6B—H6B	108.8
H5AA—C5A—H5AB	108.0	O1B—C7B—H7BA	109.2
C6A—C5A—H5AA	109.4	O1B—C7B—H7BB	109.2
C6A—C5A—H5AB	109.4	O1B—C7B—C8B	112 (2)
N2A—C6A—C1A	107.4 (19)	H7BA—C7B—H7BB	107.9
N2A—C6A—H6A	110.0	C8B—C7B—H7BA	109.2
C1A—C6A—H6A	110.0	C8B—C7B—H7BB	109.2
C5A—C6A—N2A	105.6 (16)	C9B—C8B—C7B	120 (3)
C5A—C6A—C1A	113.6 (16)	C9B—C8B—C13B	119 (2)
C5A—C6A—H6A	110.0	C13B—C8B—C7B	121 (3)
O1A—C7A—H7AA	111.0	C8B—C9B—H9B	121.9
O1A—C7A—H7AB	111.0	C10B—C9B—C8B	116 (3)
O1A—C7A—C8A	103.9 (18)	C10B—C9B—H9B	121.9
H7AA—C7A—H7AB	109.0	C9B—C10B—H10B	117.2
C8A—C7A—H7AA	111.0	C11B—C10B—C9B	126 (3)
C8A—C7A—H7AB	111.0	C11B—C10B—H10B	117.2
C9A—C8A—C7A	121 (3)	C10B—C11B—H11B	121.4
C9A—C8A—C13A	119 (3)	C12B—C11B—C10B	117 (3)
C13A—C8A—C7A	119 (3)	C12B—C11B—H11B	121.4
C8A—C9A—H9A	121.5	C11B—C12B—H12B	118.8
C8A—C9A—C10A	117 (3)	C11B—C12B—C13B	122 (3)
C10A—C9A—H9A	121.5	C13B—C12B—H12B	118.8
C9A—C10A—H10A	120.0	C8B—C13B—H13B	120.5
C11A—C10A—C9A	120 (3)	C12B—C13B—C8B	119 (3)
C11A—C10A—H10A	120.0	C12B—C13B—H13B	120.5
C10A—C11A—H11A	119.4	F1C—C1C—F2C	105.5 (15)
C12A—C11A—C10A	121 (3)	F1C—C1C—C2C	112.4 (19)
C12A—C11A—H11A	119.4	F2C—C1C—C2C	112.3 (19)
C11A—C12A—H12A	119.0	F3C—C1C—F1C	105.9 (17)
C11A—C12A—C13A	122 (4)	F3C—C1C—F2C	106.2 (19)
C13A—C12A—H12A	119.0	F3C—C1C—C2C	114.0 (15)
C8A—C13A—H13A	120.2	O1C—C2C—O2C	130.8 (17)
C12A—C13A—C8A	120 (3)	O1C—C2C—C1C	115.5 (17)
C12A—C13A—H13A	120.2	O2C—C2C—C1C	113.4 (19)
C1B—O1B—C7B	116.5 (19)	F1D—C1D—F2D	106.0 (14)
H1BA—N1B—H1BB	107.4	F1D—C1D—F3D	107.8 (18)
C4B—N1B—H1BA	108.4	F1D—C1D—C2D	112.2 (18)
C4B—N1B—H1BB	108.4	F2D—C1D—C2D	109.7 (19)
C5B—N1B—H1BA	108.4	F3D—C1D—F2D	108.8 (19)
C5B—N1B—H1BB	108.4	F3D—C1D—C2D	112.2 (15)
C5B—N1B—C4B	115.7 (16)	O1D—C2D—C1D	115.4 (17)
N3B—N2B—C6B	113.9 (19)	O2D—C2D—O1D	128.7 (16)
N4B—N3B—N2B	170 (3)	O2D—C2D—C1D	115.6 (15)
O1A—C1A—C2A—F1A	145.1 (17)	F1B—C2B—C3B—F3B	−51 (2)
O1A—C1A—C2A—C3A	−91 (2)	F1B—C2B—C3B—C4B	−172.2 (18)
O1A—C1A—C6A—N2A	−71 (2)	F2B—C3B—C4B—N1B	37 (3)
O1A—C1A—C6A—C5A	172.1 (19)	F3B—C3B—C4B—N1B	153.7 (17)

O1A—C7A—C8A—C9A	−70 (3)	N1B—C5B—C6B—N2B	155.3 (19)
O1A—C7A—C8A—C13A	118 (2)	N1B—C5B—C6B—C1B	−84 (3)
F1A—C2A—C3A—F2A	−72.0 (17)	N3B—N2B—C6B—C1B	81 (2)
F1A—C2A—C3A—F3A	170.6 (12)	N3B—N2B—C6B—C5B	−157 (2)
F1A—C2A—C3A—C4A	48.5 (19)	C1B—O1B—C7B—C8B	171 (2)
F2A—C3A—C4A—N1A	−173.5 (17)	C1B—C2B—C3B—F2B	−62 (2)
F3A—C3A—C4A—N1A	−57 (2)	C1B—C2B—C3B—F3B	−174.5 (17)
N1A—C5A—C6A—N2A	148.0 (17)	C1B—C2B—C3B—C4B	64 (3)
N1A—C5A—C6A—C1A	−94 (2)	C2B—C1B—C6B—N2B	−174.0 (19)
N3A—N2A—C6A—C1A	80 (2)	C2B—C1B—C6B—C5B	65 (3)
N3A—N2A—C6A—C5A	−159 (2)	C2B—C3B—C4B—N1B	−88 (2)
C1A—O1A—C7A—C8A	177.4 (19)	C4B—N1B—C5B—C6B	35 (3)
C1A—C2A—C3A—F2A	161.7 (17)	C5B—N1B—C4B—C3B	47 (3)
C1A—C2A—C3A—F3A	44 (2)	C6B—N2B—N3B—N4B	−158 (12)
C1A—C2A—C3A—C4A	−78 (2)	C6B—C1B—C2B—F1B	−169.4 (16)
C2A—C1A—C6A—N2A	166.2 (18)	C6B—C1B—C2B—C3B	−46 (3)
C2A—C1A—C6A—C5A	50 (2)	C7B—O1B—C1B—C2B	−105 (2)
C2A—C3A—C4A—N1A	64 (2)	C7B—O1B—C1B—C6B	129 (2)
C4A—N1A—C5A—C6A	74 (2)	C7B—C8B—C9B—C10B	177 (3)
C5A—N1A—C4A—C3A	−52 (2)	C7B—C8B—C13B—C12B	−176 (3)
C6A—C1A—C2A—F1A	−97 (2)	C8B—C9B—C10B—C11B	0 (5)
C6A—C1A—C2A—C3A	27 (3)	C9B—C8B—C13B—C12B	3 (4)
C7A—O1A—C1A—C2A	−87 (2)	C9B—C10B—C11B—C12B	0 (5)
C7A—O1A—C1A—C6A	148.9 (19)	C10B—C11B—C12B—C13B	1 (4)
C7A—C8A—C9A—C10A	−175 (2)	C11B—C12B—C13B—C8B	−3 (4)
C7A—C8A—C13A—C12A	178 (2)	C13B—C8B—C9B—C10B	−2 (4)
C8A—C9A—C10A—C11A	2 (4)	F1C—C1C—C2C—O1C	−30 (3)
C9A—C8A—C13A—C12A	7 (4)	F1C—C1C—C2C—O2C	155.7 (18)
C9A—C10A—C11A—C12A	−5 (4)	F2C—C1C—C2C—O1C	89 (2)
C10A—C11A—C12A—C13A	9 (5)	F2C—C1C—C2C—O2C	−86 (2)
C11A—C12A—C13A—C8A	−9 (4)	F3C—C1C—C2C—O1C	−150 (2)
C13A—C8A—C9A—C10A	−3 (3)	F3C—C1C—C2C—O2C	35 (3)
O1B—C1B—C2B—F1B	71 (2)	F1D—C1D—C2D—O1D	−154.0 (19)
O1B—C1B—C2B—C3B	−165.6 (17)	F1D—C1D—C2D—O2D	31 (3)
O1B—C1B—C6B—N2B	−55 (2)	F2D—C1D—C2D—O1D	89 (2)
O1B—C1B—C6B—C5B	−176 (2)	F2D—C1D—C2D—O2D	−86 (3)
O1B—C7B—C8B—C9B	−64 (4)	F3D—C1D—C2D—O1D	−32 (3)
O1B—C7B—C8B—C13B	115 (3)	F3D—C1D—C2D—O2D	153 (2)
F1B—C2B—C3B—F2B	61 (2)		

Hydrogen-bond geometry (\AA , $^\circ$)

Cg1 and Cg2 are the mid-points of the C10A—C11A and C10B—C11B bonds, respectively.

$D\cdots H\cdots A$	$D\cdots H$	$H\cdots A$	$D\cdots A$	$D\cdots H\cdots A$
N1A—H1AA \cdots O1D ⁱ	0.91	1.89	2.75 (2)	156
N1A—H1AB \cdots O2D ⁱⁱ	0.91	1.79	2.677 (19)	163
N1B—H1BA \cdots O1C ⁱⁱⁱ	0.91	1.76	2.66 (2)	170
N1B—H1BB \cdots O2C ^{iv}	0.91	1.85	2.75 (2)	167

N1A—H1AA···O1C	0.91	2.57	2.88 (3)	101
N1A—H1AB···O1C	0.91	2.57	2.88 (3)	101
N1B—H1BB···O2D ^v	0.91	2.63	2.94 (3)	101
N1B—H1BA···O2D ^v	0.91	2.70	2.94 (3)	96
N1A—H1AB···F3D ⁱ	0.91	2.60	3.010 (16)	108
N1B—H1BB···F3C ^{iv}	0.91	2.55	2.954 (17)	108
C4A—H4AA···O1C	0.99	2.47	3.12 (3)	122
C6A—H6A···O1D ⁱ	1.00	2.42	3.14 (3)	129
C4B—H4BA···O2D ^v	0.99	2.39	3.08 (3)	126
C6B—H6B···O2C ⁱⁱⁱ	1.00	2.37	3.22 (3)	142
C12B—H12B···N4A ^{vi}	0.95	2.71	3.42 (4)	133
C4A—H4AA···F3B	0.99	2.50	3.29 (3)	137
C4A—H4AA···F3C ^v	0.99	2.69	3.27 (3)	118
C5A—H5AA···F1A	0.99	2.59	3.18 (2)	118
C5A—H5AB···F3D ⁱ	0.99	2.70	3.31 (2)	120
C6A—H6A···F3A	1.00	2.35	2.88 (2)	112
C4B—H4BA···F2A	0.99	2.55	3.39 (3)	142
C4B—H4BA···F3D	0.99	2.85	3.34 (3)	111
C5B—H5BA···F2B	0.99	2.51	2.95 (2)	107
C7B—H7BA···F1B	0.99	2.59	3.15 (5)	116
C7B—H7BA···Cg1 ^v	0.99	2.87	3.73 (4)	146
C7A—H7AA···Cg2 ^v	0.99	2.64	3.46 (4)	140

Symmetry codes: (i) $x, y, z-1$; (ii) $x+1, y, z-1$; (iii) $x, y, z+1$; (iv) $x+1, y, z+1$; (v) $x+1, y, z$; (vi) $-x, y+1/2, -z+1$.