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Dissolution Kinetics of Cerussite in Acidic Sodium Chloride Solutions

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Abstract

The dissolution kinetics of cerussite in acidic sodium chloride solutions was investigated with respect to experimental variables such as particle size, stirring speed, sodium chloride concentration, hydrochloric acid concentration, and reaction temperature. Results show that leaching reagent concentration and reaction temperature have significant effects on the extraction of lead, whereas particle size and stirring speed have a relatively moderate effect on the leaching rate. The dissolution process followed the kinetic law of the shrinking core model, and a corresponding mixed control model was found suitable to represent the rate of the controlling step. The apparent activation energy of this process was determined to be 40.46 kJ/mol, and a corresponding dissolution kinetic equation is also presented to describe the dissolution reaction.

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